

Article

Donor-Type Nickel–Dithiolene Complexes Fused with Bulky Cycloalkane Substituents and Their Application in Molecular Conductors

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Abstract: The effects of substituents on the arrangement of metal–dithiolene complexes based on π -conjugated systems, which are extensively used to synthesize various functional materials, have not been studied adequately. New donor-type nickel–dithiolene complexes fused with bulky cycloalkane substituents [Ni(C_n-dddt)₂] (C₅-dddt = 4a,5,6,6a-pentahydro-1,4-benzodithiin-2,3-dithiolate; C₆-dddt = 4a,5,6,7,8,8a-hexahydro-1,4-benzodithiin-2,3-dithiolate; C₇-dddt = 4a,5,6,7,8,9,9a-heptahydro-1,4-benzodithiin-2,3-dithiolate; and C₈-dddt = 4a,5,6,7,8,9,10,10a-octahydro-1,4-benzodithiin-2,3-dithiolate) were synthesized in this study. All the complexes were crystallized in cis-[Ni(cis-C_n-dddt)₂] conformations with cis-oriented (*R,S*) conformations around the cycloalkylene groups in the neutral state. Unique molecular arrangements with a three-dimensional network, a one-dimensional column, and a helical molecular arrangement were formed in the crystals owing to the flexible cycloalkane moieties. New 2:1 cation radical crystals of [Ni(C₅-dddt)₂]₂(X) (X = ClO₄[−] or PF₆[−]), obtained by electrochemical crystallization, exhibited semiconducting behaviors ($\rho_{rt} = 0.8 \Omega \text{ cm}$, $E_a = 0.09 \text{ eV}$ for the ClO₄[−] crystal; $4.0 \Omega \text{ cm}$, 0.13 eV for the PF₆[−] crystal) under ambient pressure due to spin-singlet states between the dimers of the donor, which were in accordance with the conducting behaviors under hydrostatic pressure ($\rho_{rt} = 0.2 \Omega \text{ cm}$, $E_a = 0.07 \text{ eV}$ for the ClO₄[−] crystal; $1.0 \Omega \text{ cm}$, 0.12 eV for the PF₆[−] crystal at 2.0 GPa).

Keywords: nickel–dithiolene complex; cycloalkane substituent; molecular conductor



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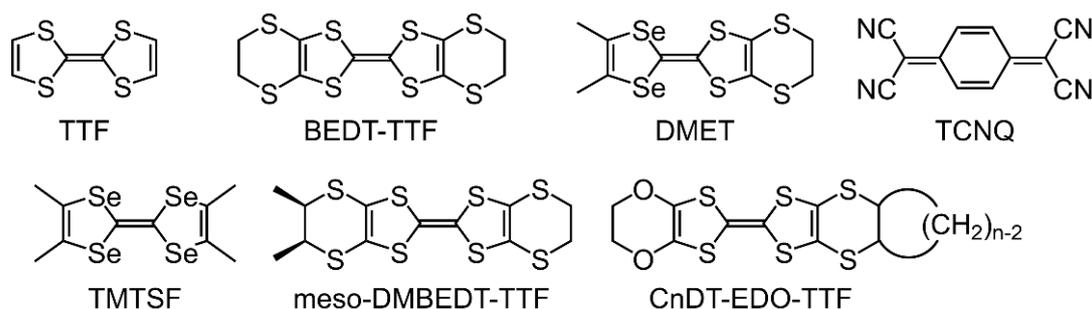


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1. Introduction

Organic π -conjugated molecules, including tetrathiafulvalene (TTF) skeletons, are among the most common building blocks for the realization of various functional materials [1]. The TTF molecule has an electron donor ability, which can help realize various charge transfers and produce radical cation crystals. The TTF-TCNQ (TCNQ: 7,7,8,8-tetracyanoquinodimethane) is a prototype of charge-transfer compounds, where the highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO) bands of the open-shell donors and acceptors, respectively, contribute to the conduction [2,3]. It is the first organic conductor to exhibit a metallic conductivity within a wide temperature range with a minimum of 59 K, where a sharp metal-to-insulator transition is observed [4]. Moreover, TTF has attracted significant research attention with the discovery of superconducting salts [5–8] based on TTF derivatives such as tetramethyltetraselenafulvalene (TMTSF) [9], bis(ethylenedithio)tetrathiafulvalene (BEDT-TTF) [10,11], and dimethyl(ethylenedithio)diselenadithiafulvalene (DMET) [12,13]. The conducting properties of molecular-based materials are significantly influenced by their molecular arrangements in

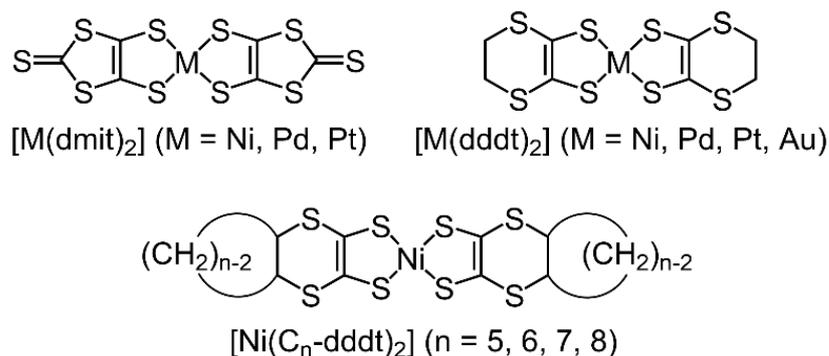
crystal form. The cation radicals of BEDT-TTF and bis(ethylenedithio)tetraselenafulvalene (BETS) can crystallize in various donor packing motifs, including α -, β -, θ -, κ -, and λ -type arrangements by combination with inorganic, organic, and organometallic anions [14–18]. Alternative strategies for tuning the molecular arrangements in organic conductors were reported based on organic synthesis techniques. A potential method involves the direct introduction of various substituents into the TTF or BEDT-TTF skeletons [19–42]. Kimura et al. reported that meso-2-(5,6-dihydro-1,3-dithiolo[4,5-b][1,4]dithiin-2-ylidene)-5,6-dihydro-5,6-dimethyl-1,3-dithiolo[4,5-b][1,4]dithiin (meso-DMBEDT-TTF) forms a superconducting crystal β -(meso-DMBEDT-TTF)₂PF₆ with a transition temperature at 4.3 K under a hydrostatic pressure of 4.0 kbar due to a checkerboard-type charge order state on the donor molecules [43,44]. Another alternative method involves the introduction of bulky cycloalkane moieties into the donor molecules to determine the stacking of the donor molecules; thus, the ratio of the on-site Coulomb repulsion energy U to the bandwidth W can be controlled by changing W for the organic Mott insulators based on the enhanced steric hindrance in the radical cation crystals of C_nDT-EDO-TTF ($n = 5, 6, 7, 8$; *cis*-1,2-cycloalkylene-1,2-dithio)ethylenedioxytetrafulvalene] and its derivatives (Scheme 1) [45,46].



Scheme 1. Organic components of molecular conductors.

Metal bis-dithiolene complexes can exhibit various oxidation states [47] and have smaller HOMO–LUMO gaps than those of the organic donors [48,49], which is advantageous for the components of molecular conductors. Electron-acceptor type [M(dmit)₂] (M = Pd, Pt; dmit = 1,3-dithiole-2-thion-4,5-dithiole) complexes with small gaps between the HOMO and LUMO form dimer structures. These structures induce crossing HOMO–LUMO band inversions through metal–metal interactions in the solid state, which provide various conducting, magnetic, and other physical properties by combining with closed-shell cations [47,50–52]. The electron-donor-type metal bis-dithiolene complexes [M(dddt)₂] (M = Ni, Pd, Pt, and Au; dddt = 5,6-dihydro-1,4-dithiine-2,3-dithiolate) are candidate materials for the development of molecular conductors due to their orbital symmetries, which are similar to those of BEDT-TTF [53–56]. A cation radical crystal, [Ni(dddt)₂]₃(AuBr₂)₂, is the first metal based on the donor-type metal–dithiolene complexes exhibiting metallic behavior down to at least 1.3 K [57]. Recently, Kato et al. reported a single-component molecular Dirac electron system based on Pd(dddt)₂ that exhibits temperature-independent resistivity under high hydrostatic pressures [58–63]. The conducting and magnetic properties exhibited by the molecular crystals based on the dithiolene complexes are significantly dependent on the molecular arrangements of the metal–dithiolene complexes and organic molecular conductors. In the case of molecular conductors based on metal–dithiolene complexes, the molecular arrangements of the complexes are dependent on the combination of counter cations and anions in the crystals [47,48]. However, there are few reports on the substituent effects on the molecular arrangement, which influences the physical properties, despite the advantages of the components of molecular conductors [64,65]. Thus, we designed novel donor-type metal bis-dithiolene complexes [Ni(C_n-dddt)₂] formed by a [Ni(dddt)₂] skeleton fused with bulky cycloalkane rings and investigated the substituent effects on the arrangement of the complexes and their physical properties (Scheme 2).

Donor organic molecules fused with bulky substituents such as cycloalkane and dioxane moieties can form unique molecular arrangements in the crystals [66], which exhibit superconducting and metallic behaviors [45,46,67–69].



Scheme 2. Metal–dithiolene complexes as components of molecular conductors.

In this paper, we report the donor abilities of novel donor complexes $[\text{Ni}(\text{C}_n\text{-dddt})_2]$ with two cycloalkane rings ($\text{C}_5\text{-dddt} = 4\text{a},5,6,6\text{a-pentahydro-1,4-benzodithiin-2,3-dithiolate}$; $\text{C}_6\text{-dddt} = 4\text{a},5,6,7,8,8\text{a-hexahydro-1,4-benzodithiin-2,3-dithiolate}$; $\text{C}_7\text{-dddt} = 4\text{a},5,6,7,8,9,9\text{a-heptahydro-1,4-benzodithiin-2,3-dithiolate}$, and $\text{C}_8\text{-dddt} = 4\text{a},5,6,7,8,9,10,10\text{a-octahydro-1,4-benzodithiin-2,3-dithiolate}$) based on electrochemical measurements and electronic absorption spectra. The packing motifs of the neutral $[\text{Ni}(\text{C}_n\text{-dddt})_2]$ ($n = 5, 6, 7, 8$) were discussed based on X-ray crystallographic analysis results. Furthermore, electronic features of new cation radical crystals $[\text{Ni}(\text{C}_5\text{-dddt})_2]_2[\text{X}]$ ($\text{X} = \text{ClO}_4^-$ and PF_6^-) were revealed using electrical resistivity, magnetic measurements, and band calculations based on the crystal structures.

2. Materials and Methods

2.1. Materials

All the solvents were of analytical grade and were used without further purification. All the reactions were performed in a nitrogen atmosphere. Oligo(1,3-dithiole-2,4,5-trithione) ($(\text{C}_3\text{S}_5)_x$) [70] and $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$ [56] were prepared using a previously reported method.

2.2. Synthesis

2.2.1. Synthesis of 4,5-Cis(cyclopentanedithio)-1,3-dithiole-2-thione (**L1**), 4,5-Cis(cyclohexylenedithio)-1,3-dithiole-2-thione (**L2**), 4,5-Cis(cycloheptalenedithio)-1,3-dithiole-2-thione (**L3**), and 4,5-Cis(cyclooctalenedithio)-1,3-dithiole-2-thione (**L4**)

In a nitrogen atmosphere, cyclopentene (25 mmol) was dissolved in 1,2-dichloroethane (25 mL), and oligo(1,3-dithiole-2,4,5-trithione) (10 mmol) was added. The resulting suspension was heated to 135 °C for 14 h with stirring, and the reaction mixture was filtered to remove insoluble precipitates. The evaporation of the filtration in vacuo, followed by purification of the residue by column chromatography (SiO_2 , carbon disulfide (CS_2)-hexane 1:3), yielded 67.6% 4,5-cis(cyclopentanedithio)-1,3-dithiole-2-thione (**L1**) as yellow crystals after recrystallization from hot ethanol (EtOH). Moreover, 4,5-cis(cyclohexylenedithio)-1,3-dithiole-2-thione (**L2**), 4,5-cis(cycloheptalenedithio)-1,3-dithiole-2-thione (**L3**), and 4,5-cis(cyclooctalenedithio)-1,3-dithiole-2-thione (**L4**) were obtained using a similar procedure with corresponding cycloalkenes as starting materials, yielding 32.7% **L2**, 77.6% **L3**, and 69.2% **L4**.

2.2.2. Synthesis of 4,5-Cis(cyclopentanedithio)-1,3-dithiole-2-one (**L1'**) and 4,5-Cis(cyclohexylenedithio)-1,3-dithiole-2-one (**L2'**), 4,5-Cis(cycloheptanedithio)-1,3-dithiole-2-one (**L3'**), and 4,5-Cis(cyclooctanedithio)-1,3-dithiole-2-one (**L4'**)

Mercuric acetate (1.5 mmol) was added to a solution of **L1'** (1 mmol) in tetrahydrofuran-acetic acid (2:1, 45 mL), and the resulting white suspension was stirred for 30 min at 20 °C. The resulting white precipitate was removed by filtration using Celite and washed thoroughly using chloroform. The combined organic phases were washed using water (2 × 100 mL), saturated sodium hydrogen carbonate (2 × 100 mL), and water (1 × 100 mL), and then they were dried using magnesium sulfate. The removal of the solvent in vacuo yielded 98% 4,5-cis(cyclopentanedithio)-1,3-dithiole-2-one (**L1'**) as large pale yellow needle crystals. Moreover, 4,5-cis(cyclohexylenedithio)-1,3-dithiole-2-one (**L2'**), 4,5-cis(cycloheptanedithio)-1,3-dithiole-2-one (**L3'**), and 4,5-cis(cyclooctanedithio)-1,3-dithiole-2-one (**L4'**) were obtained using a similar procedure; thus yielding 97% **L2'**, 98% **L3'**, and 97% **L4'**.

2.2.3. Synthesis of Monoanionic Complexes $(\text{Bu}_4\text{N})[\text{Ni}(\text{C}_n\text{-dddtd})_2]$ ($n = 5, 1; n = 6, 2; n = 7, 3;$ and $n = 8, 4$)

Sodium (3.0 mmol) and **Lm'** ($m = 1-4$) (1.2 mmol) were dissolved in EtOH (15 mL). The resulting clear orange solution was dropped into an EtOH solution (15 mL) of $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (0.6 mmol) for 30 min with stirring. The resulting solution had a dark purple color. It was stirred in the air for a further 30 min, and the color changed to dark green as the reaction proceeded. After filtration to remove insoluble precipitates, one equivalent tetrabutylammonium bromide was added to the solution, and a dark green solid precipitated immediately. The green solids of monoanionic complexes **1-4** were collected by filtration and recrystallized using acetone. The yields for **1, 2, 3,** and **4** were 79.6%, 63.0%, 46.4%, and 38.8%, respectively.

2.2.4. Synthesis of Neutral Complexes $[\text{Ni}(\text{C}_n\text{-dddtd})_2]$ ($n = 5, 5; n = 6, 6; n = 7, 7;$ and $n = 8, 8$)

An acetonitrile (CH_3CN) (50 mL) solution of iodine (1.3 mmol) was added to a CH_3CN solution (100 mL) of monoanionic complexes **1-4** (0.23 mmol). Black precipitates of **5-8** were obtained immediately. The precipitates of the neutral complexes were collected by filtration, washed using CH_3CN and methanol, and dried under reduced pressure. Black block-shaped single crystals of the neutral complexes **5-8**, which were suitable for X-ray crystallographic analyses, were obtained by recrystallization from their carbon disulfide solutions. The yields for **5, 6, 7,** and **8** were 69.2%, 31.5%, 62.9%, and 36.8%, respectively. The elemental analysis results for these complexes are summarized in Table 1.

Table 1. Results of the elemental analyses of **1-8**.

Complex Formula	Observed (Calculation Result) (%)			Complex Formula	Observed (Calculation Result) (%)			
	C	H	N		C	H	Ni	S
1 $\text{C}_{30}\text{H}_{52}\text{NNiS}_8$	48.32 (48.56)	6.91 (7.06)	1.89 (1.89)	5 $\text{C}_{14}\text{H}_{16}\text{NiS}_8$	33.59 (33.66)	3.14 (3.23)	11.58 (11.75)	51.33 (51.36)
2 $\text{C}_{32}\text{H}_{56}\text{NNiS}_8$	49.65 (49.91)	7.26 (7.33)	1.85 (1.82)	6 $\text{C}_{16}\text{H}_{20}\text{NiS}_8$	36.21 (36.43)	3.70 (3.82)	11.08 (11.13)	48.60 (48.63)
3 $\text{C}_{34}\text{H}_{60}\text{NNiS}_8$	50.88 (51.17)	7.50 (7.58)	1.67 (1.76)	7 $\text{C}_{18}\text{H}_{24}\text{NiS}_8$	38.87 (38.91)	4.26 (4.35)	10.44 (10.56)	46.40 (46.17)
4 $\text{C}_{36}\text{H}_{64}\text{NNiS}_8$	51.90 (52.34)	7.71 (7.81)	1.70 (1.70)	8 $\text{C}_{20}\text{H}_{28}\text{NiS}_8$	41.00 (41.16)	4.81 (4.84)	9.81 (10.06)	43.86 (43.95)

2.2.5. Synthesis of Radical Cation Crystals $[\text{Ni}(\text{C}_5\text{-dddtd})_2][\text{ClO}_4]$ (**9**) and $[\text{Ni}(\text{C}_5\text{-dddtd})_2][\text{PF}_6]$ (**10**)

Dichloromethane (CH_2Cl_2) solutions (10 mL) of **5** (0.03 mmol) containing $(\text{Bu}_4\text{N})(\text{X})$ ($\text{X} = \text{ClO}_4$ or PF_6) (0.09 mmol) as electrolytes were electrolyzed under constant-current conditions (3 μA) for three days in an H-shaped cell equipped with platinum wire electrodes in an argon atmosphere at 20 °C. Black crystals of **9** and **10** were formed on the anode surface. The chemical formulas of **9** and **10** were confirmed by X-ray crystal structural analyses.

2.3. Physical Measurements

2.3.1. Elemental Analyses

The analyses of the carbon, hydrogen, and nitrogen elements were performed using a PERKIN-ELMER 240 C elemental analyzer for **1**, **3**, **7**; and a Yanako MT-6 CHN CORDER for **2**, **4**, **5**, **6**, and **8**. The elemental analyses of the nickel and sulfur atoms of the neutral complexes **5–8** were performed using a SHIMADZU ICPS-1000IV by the Chemical Analysis Team of AD&SD in RIKEN.

2.3.2. Electrochemical Measurements

Cyclic voltammograms of monoanionic complexes **1–4** were obtained using acetonitrile solutions ($1 \times 10^{-3} \text{ mol L}^{-1}$) with $(\text{Bu}_4\text{N})(\text{ClO}_4)$ (0.1 M) as a supporting electrolyte and recorded using a BAS ALS/[H] CH Instruments Model 610 electrochemical analyzer combined with an Ag/Ag⁺ reference electrode, platinum counter electrode, and glassy carbon working electrode at room temperature. The scan rate of the measurements was 100 mV/s.

2.3.3. Electronic Absorption Spectra

The electronic absorption spectra (300–1600 nm) of **1–8** in CH_2Cl_2 , CH_3CN , and benzene solutions ($5 \times 10^{-5} \text{ mol L}^{-1}$) at room temperature were measured by the Chemical Analysis Team of AD&SD in RIKEN using a Shimadzu UV-3100 system.

2.3.4. Electrical Resistivity

The temperature dependence of the electrical resistivity was measured using the standard four-probe method at ambient pressure for **9** and **10**. Gold wires (15 μm in diameter) were attached to the crystal with a carbon paste. Resistivity measurements under the hydrostatic pressure of the salt were performed within the range from 0.3–1.8 GPa using a clamp-type piston-cylinder high-pressure cell [71,72].

2.3.5. Magnetic Susceptibility

The magnetic susceptibilities of **9** (15 mg) and **10** (20 mg) were measured in a magnetic field of 1 T using a superconducting quantum interference device (SQUID) magnetometer (Quantum Design, MPMS-XL7) within the temperature range of 2–300 K.

2.3.6. Band Calculation

The intermolecular overlap integrals (S) between the frontier orbitals were calculated based on the extended Hückel molecular orbital (MO) method. The semi-empirical parameters for the Slater-type atomic orbitals were obtained from the literature [73–75].

2.3.7. Crystal Structure Determinations of **L2–L4**, Neutral Complexes **5–8**, and Radical Cation Crystals **9** and **10**

Crystallographic data for the single crystals of **L2–L4**, **5–8**, and **9–10** were obtained using a Rigaku MicroMax-007 diffractometer with a multilayer mirror monochromated Mo-K α ($\lambda = 0.71073 \text{ \AA}$) and charge-coupled device detector at 93 K. Structural refinements were carried out using the full-matrix least-squares method on F^2 [76]. The structures were determined and refined using SHELXL-2018 in the Olex2 software package [77,78]. The parameters were refined using anisotropic temperature factors, with the exception

of the hydrogen atoms, which were parametrically refined using the riding model with a fixed C-H bond distance of 0.95 Å. The crystallographic data for L2–L4, 5–8, and 9–10 are summarized in Tables 2–4, respectively.

Table 2. Crystallographic data for neutral complexes L2–L4.

	L2	L3	L4
Chemical formula	C _{7.2} H ₈ S ₄	C ₈ H _{9.60} S ₄	C _{7.34} H _{9.34} S _{3.34}
Crystal size/mm ³	0.40 × 0.10 × 0.04	0.30 × 0.06 × 0.03	0.40 × 0.03 × 0.02
Formula weight	222.78	234.00	204.45
Crystal system	Monoclinic	Monoclinic	Monoclinic
Space group	P2 ₁ /c	P2 ₁ /c	P2/c
a/Å	12.7079(3)	5.3224(3)	11.2835(7)
b/Å	9.8356(2)	21.7855(10)	6.8878(3)
c/Å	18.7846(4)	10.5687(5)	17.1727(13)
α/°	90	90	90
β/°	105.640(2)	95.793(5)	106.533(7)
γ/°	90	90	90
V/Å ³	2260.95(9)	1219.19(11)	1279.46(14)
Z	10	5	6
T/K	93	93	93
μ(Mo Kα)/mm ⁻¹	0.980	0.913	0.874
D _{calc} /g cm ⁻³	1.636	1.594	1.592
F(000)	1152	608	640
λ (Mo-Kα)/Å	0.71073	0.71073	0.71073
Measured 2θ range/°	4.75–60.98	3.938–60.722	3.766–60.41
No. of reflections collected	22,206	10,849	11,351
Independent reflections	16,862	7823	6705
Observed reflections with I > 2.00 σ(I)	5456	2882	2973
R _{int}	0.0191	0.0291	0.0255
R(F ²) (I > 2.00σ(I)) ^a	0.0211	0.0293	0.0267
wR(F ²) (all data) ^b	0.0537	0.0740	0.0654
Goodness of fit (GOF)	1.064	1.056	1.041
CCDC number	2,103,657	2,103,658	2,103,659

^a $R(F^2) = \Sigma (F_o^2 - F_c^2) / \Sigma F_o^2$; ^b $wR(F^2) = [\Sigma w(F_o^2 - F_c^2)^2 / \Sigma w(F_o^2)^2]^{1/2}$ and $w^{-1} = [\Sigma^2 (F^2) + (0.0278P)^2 + 0.7843P]$ for L2; $w^{-1} = [\Sigma^2 (F^2) + (0.0450P)^2 + 0.3773P]$ for L3; and $w^{-1} = [\Sigma^2 (F^2) + (0.0379 P)^2 + 0.4043P]$ for L4 (where $P = (F_o^2 + 2F_c^2)/3$).

Table 3. Crystallographic data for neutral complexes 5–8.

	5	6	7	8
Chemical formula	C ₁₄ H ₁₆ NiS ₈	C ₁₆ H ₂₀ NiS ₈	C _{7.2} H _{9.6} Ni _{0.4} S _{3.2}	C _{6.67} H _{9.32} Ni _{0.33} S _{2.66}
Crystal size/mm ³	0.14 × 0.11 × 0.06	0.16 × 0.09 × 0.06	0.07 × 0.04 × 0.01	0.18 × 0.07 × 0.07
Formula weight	499.46	527.51	222.22	194.41
Crystal system	Monoclinic	Monoclinic	Monoclinic	Triclinic
Space group	P2 ₁ /n	P2 ₁	P2 ₁ /n	P-1
a/Å	6.2830(2)	8.0303(3)	12.5305(6)	6.8973(10)
b/Å	9.8825(3)	8.8084(4)	6.7500(3)	8.6190(13)
c/Å	15.2978(4)	14.8128(5)	13.6312(6)	11.5730(15)
α/°	90	90	90	69.507(13)
β/°	100.685(3)	105.033(4)	108.057(5)	85.237(11)
γ/°	90	90	90	71.298(13)
V/Å ³	933.40(5)	1011.91(7)	1096.15(9)	610.09(16)
Z	2	2	5	3
T/K	93	93	93	93
μ(Mo Kα)/mm ⁻¹	1.928	1.783	1.651	1.486
D _{calc} /g cm ⁻³	1.777	1.731	1.683	1.587

Table 3. Cont.

	5	6	7	8
$F(000)$	512	544	576	304
λ (Mo-K α)/Å	0.71073	0.71073	0.71073	0.71073
Measured 2θ range/°	4.938–60.77	5.224–60.742	5.376–61.08	3.75–60.606
No. of reflections collected	10,661	9330	9850	9364
Independent reflections	7606	6979	6434	3699
Observed reflections with $I > 2.00\sigma(I)$	2215	4857	2436	2568
R_{int}	0.0182	0.0187	0.0226	0.0267
$R(F^2)$ ($I > 2.00\sigma(I)$) ^a	0.0237	0.0499	0.0252	0.0827
$wR(F^2)$ (all data) ^b	0.0595	0.1305	0.0544	0.1977
GOF	1.056	1.215	1.051	1.185
Flack parameter		0.328(7)		
CCDC number	2,103,660	2,103,661	2,103,662	2,103,663

^a $R(F^2) = \Sigma (F_o^2 - F_c^2) / \Sigma F_o^2$; ^b $wR(F^2) = [\Sigma w(F_o^2 - F_c^2)^2 / \Sigma w(F_o^2)^2]^{1/2}$ and $w^{-1} = [\Sigma^2 (F^2) + (0.0341P)^2 + 0.4506P]$ for 5; $w^{-1} = [\Sigma^2 (F^2) + (0.0110P)^2 + 9.3099P]$ for 6; $w^{-1} = [\Sigma^2 (F^2) + (0.0275P)^2 + 0.2769P]$ for 7; and $w^{-1} = [\Sigma^2 (F^2) + (0.0195P)^2 + 10.6333P]$ for 8 (where $P = (F_o^2 + 2F_c^2)/3$).

Table 4. Crystallographic data for the radical cation crystals 9 and 10.

	9	10
Chemical formula	C ₇ H ₈ Cl _{0.25} Ni _{0.50} OS ₄	C ₇ H ₈ F _{1.5} Ni _{0.5} P _{0.25} S ₄
Crystal size/mm ³	0.12 × 0.12 × 0.08	0.14 × 0.07 × 0.01
Formula weight	274.59	285.97
Crystal system	Monoclinic	Monoclinic
Space group	<i>P</i> 2/ <i>c</i>	<i>P</i> 2/ <i>c</i>
$a/\text{Å}$	12.6888(3)	12.6865(4)
$b/\text{Å}$	6.25150(10)	6.3020(2)
$c/\text{Å}$	24.1229(6)	24.3145(8)
$\alpha/^\circ$	90	90
$\beta/^\circ$	93.462(2)	92.900(3)
$\gamma/^\circ$	90	90
$V/\text{Å}^3$	1910.03(7)	1941.46(11)
Z	8	8
T/K	93	93
$\mu(\text{Mo K}\alpha)/\text{mm}^{-1}$	1.969	1.928
$D_{\text{calc}}/\text{g cm}^{-3}$	1.910	1.957
$F(000)$	1122	1162
λ (Mo-K α)/Å	0.71073	0.71073
Measured 2θ range/°	4.782–60.788	4.518–60.73
No. of reflections collected	16,993	17,389
Independent reflections	12,771	9840
Observed reflections with $I > 2.00\sigma(I)$	4756	4142
R_{int}	0.0135	0.0338
$R(F^2)$ ($I > 2.00\sigma(I)$) ^a	0.0302	0.0335
$wR(F^2)$ (all data) ^b	0.0659	0.0737
GOF	1.068	1.024
CCDC number	2,103,664	2,103,665

^a $R(F^2) = \Sigma (F_o^2 - F_c^2) / \Sigma F_o^2$; ^b $wR(F^2) = [\Sigma w(F_o^2 - F_c^2)^2 / \Sigma w(F_o^2)^2]^{1/2}$ and $w^{-1} = [\Sigma^2 (F^2) + (0.0186P)^2 + 4.1733P]$ for 9; and $w^{-1} = [\Sigma^2 (F^2) + (0.0403P)^2 + 1.0162P]$ for 10 (where $P = (F_o^2 + 2F_c^2)/3$).

3. Results and Discussion

3.1. Synthesis of Complexes 1–8

3.1.1. Syntheses and Molecular Structures of Precursors L1–L4

One of the most effective methods for synthesizing various TTF derivatives and their precursors involves [2 + 4] cycloadditions of $(C_3S_5)_x$ (as 4π) and a molecule that involves one of its external double bonds (as 2π) [79]. This reaction is stereospecific, similar to the [2 + 4] cycloaddition of tetrathiooxalate with alkanes, thus providing a cis-oriented compound [80]. Precursors L1–L4 were synthesized using a modified version of a method outlined in a previous study, by the [2 + 4] cycloadditions of $(C_3S_5)_x$ and cycloheptene, cyclohexene, cycloheptene, and cyclooctene [81]. The cycloalkylene residues were cis-oriented with respect to the newly formed 1,4-dithiene rings. Single crystals of L1–L4 were obtained by recrystallization from their dichloromethane solutions, which were suitable for X-ray crystallographic analyses, with the exception of L1. Figure 1 presents the molecular structures of two crystallographically independent molecules A and B in the crystal of L2. The fused cyclohexane moieties in molecules A and B formed cis-oriented (*R,S*) conformations at the two carbons C(4)–C(5) and C(13)–C(14), respectively. Stable chair-type configurations of the cyclohexylene groups were observed for molecules A and B. A similar cis-oriented conformation was observed in the crystals of 4,5-(1,4-dioxanediyl-2,3-dithio)-1,3-dithiole-2-thione and 4,5-(cis-cyclohexylenedithio)-1,3-dithiole-2-one [82,83]. The bond distances of C=C and C=S were 1.349 Å and 1.648 Å for molecule A, and 1.352 Å and 1.648 Å for molecule B. The angles between the planes formed by S(1)–C(2)–C(3)–S(2) and S(3)–C(2)–C(3)–S(4) in molecule A and S(6)–C(11)–C(12)–S(7) and S(8)–C(11)–C(12)–S(9) in molecule B were 3.86° and 3.08°, respectively. The bond distances and angles indicate that the C_3S_5 moieties of molecules A and B formed similar planar structures. The torsion angles around the fused cyclohexyl group in molecule A were -71.20° , -6.76° , and 56.18° for C(2)–S(1)–C(4)–C(5), C(3)–S(2)–C(5)–C(4), and S(1)–C(4)–C(5)–S(2), respectively. The corresponding torsion angles in molecule B were -62.56° , -26.37° , and 63.91° for C(11)–S(6)–C(13)–C(14), C(12)–S(7)–C(14)–C(13), and S(6)–C(13)–C(14)–S(7), respectively. The comparison of the torsion angles between molecules A and B revealed that the fused cyclohexylene group in molecule A deviated more from the C_3S_5 plane at the S(1) and S(2) positions than that of molecule B.

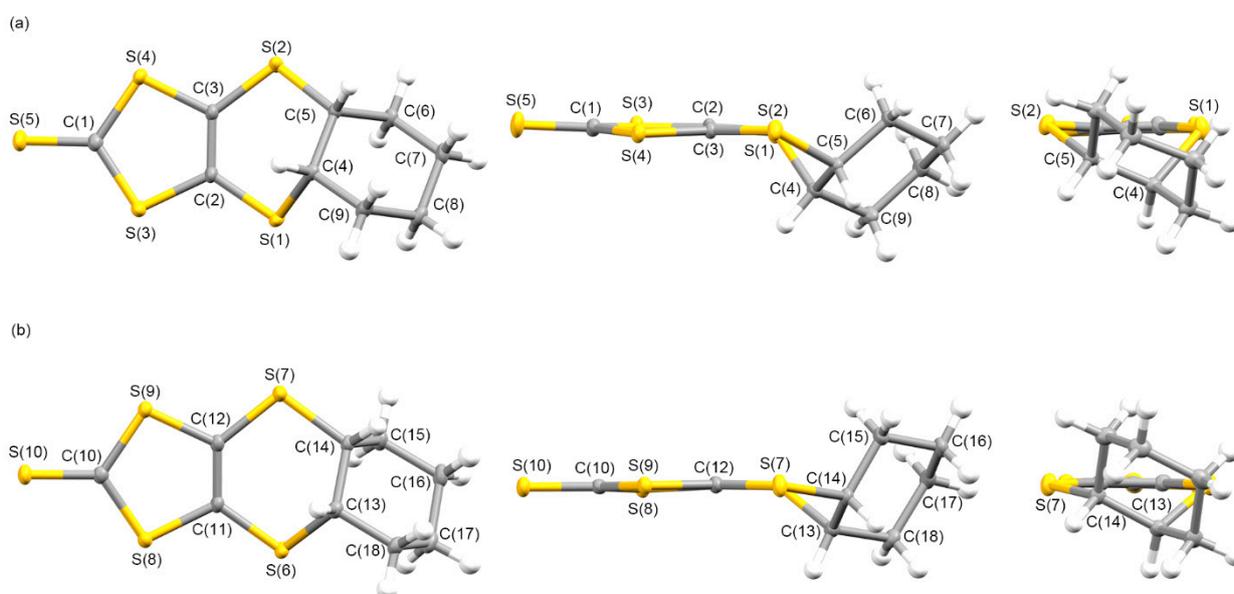


Figure 1. Molecular structures of two crystallographically independent molecules, (a) A and (b) B, in the crystal of L2 along the vertical, parallel, and end-on directions on the C_3S_5 molecular plane with atom labels.

Figure 2 presents the molecular structures of L3 and L4. The fused cycloheptane

and cyclooctane moieties formed cis-oriented (*R,S*) conformations at the two carbons C(4) and C(5), respectively, in addition to molecule **L2**. Twist-chair- and boat-chair-type configurations of the cycloheptylene and cyclooctylene groups were formed in **L3** and **L4**, which were observed in the crystals of C_nDT -EDO-TTF ($n = 7, 8$) [45]. The C_3S_5 moieties of **L3** and **L4** were planar structures, as observed in the crystal of **L2**; where the bond lengths C(1)=S(5) and C(2)=C(3) and the angles between the planes formed by S(1)-C(2)-C(3)-S(2) and S(3)-C(2)-C(3)-S(4) were 1.656 Å, 1.355 Å, and 2.89° for **L3**; and 1.647 Å, 1.344 Å, and 2.75° for **L4**, respectively. The torsion angles around the fused cycloalkylene groups of C(2)-S(1)-C(4)-C(5), C(3)-S(2)-C(5)-C(4), and S(1)-C(4)-C(5)-S(2) were 57.68°, 42.60°, and -72.47° for **L3**; and 58.00°, 41.11°, and -72.18° for **L4**, respectively.

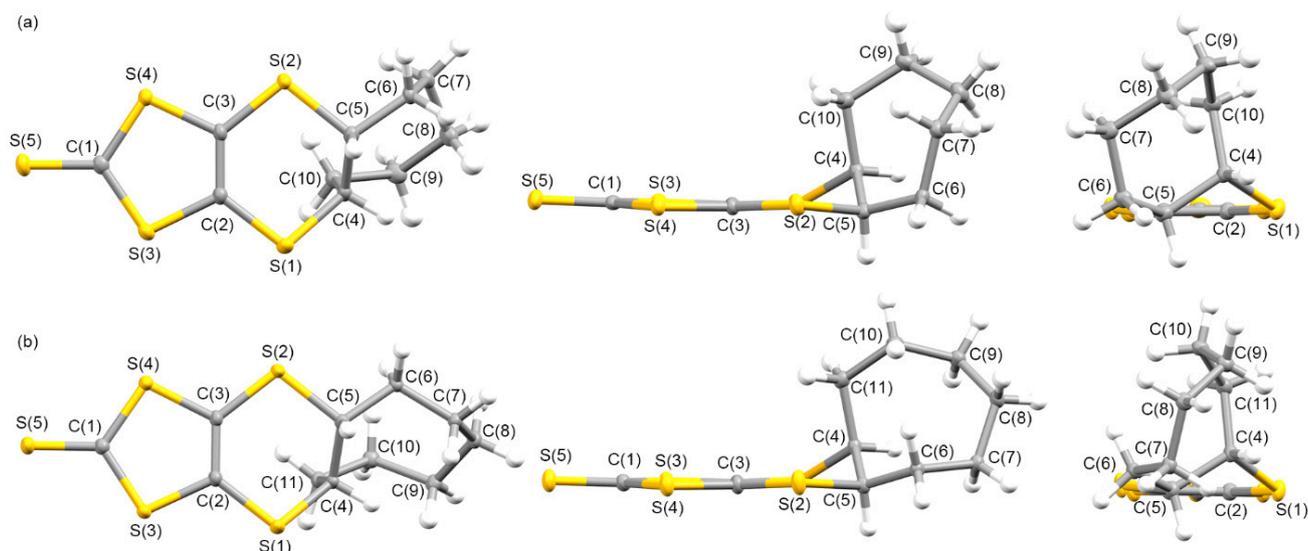
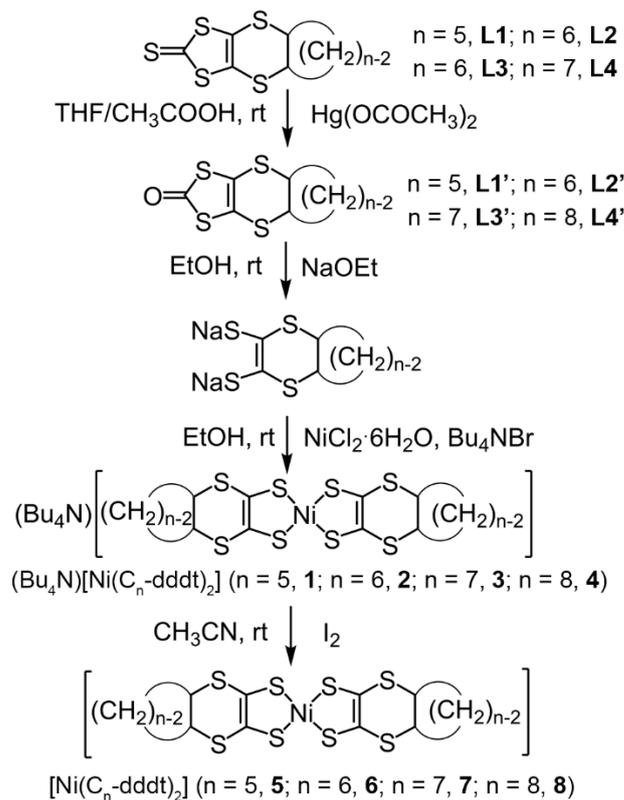


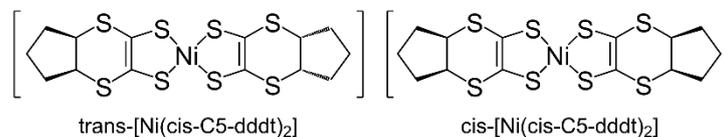
Figure 2. Molecular structures of (a) **L3** and (b) **L4** along the vertical, parallel, and end-on directions on the C_3S_5 molecular plane with atom labels.

3.1.2. Molecular and Crystal Structures of Neutral Complexes 5–8

We successfully obtained single crystals of the neutral complexes that were suitable for X-ray crystallographic analysis by recrystallization from CS_2 solutions of complexes 5–8. The synthetic procedure for the neutral complexes is shown in Scheme 3. The precursors of **L1–L4** formed cis-oriented (*R,S*) conformations around the cycloalkylene groups due to the [2 + 4] cycloadditions. Using the synthetic procedures shown in Scheme 3, Monoanionic complexes 1–4 were obtained as mixtures, including diastereomeric isomers $(Bu_4N)\{\text{trans}[\text{Ni}(\text{cis-}C_n\text{-dddt})_2]\}$ and $(Bu_4N)\{\text{cis}[\text{Ni}(\text{cis-}C_n\text{-dddt})_2]\}$, due to the conformations between the two cycloalkylene groups in the molecules. The oxidation of Monoanionic complexes 1–4 resulted in the formation of mixtures of diastereomeric isomers for neutral complexes 5–8 (Scheme 4). From the recrystallization processes, single crystals of neutral complexes 5–8 were formed by the trans-isomers.



Scheme 3. Synthetic procedure of neutral [Ni(Cn-dddt)₂] complexes.



Scheme 4. Diastereomeric isomers of the trans- and cis-[Ni(cis-C5-dddt)₂] complexes.

The selected bond lengths and angles of the neutral complexes in crystals **5–8** are summarized in Table 5. In molecules of **5–8**, square planar co-ordinations were constructed around the Ni atoms, similar to previously reported square planar Ni-dithiolene complexes such as [Ni(dddtdt)₂] in a neutral state [84–91].

Table 5. Selected bond lengths and angles of neutral complexes **5–8**.

	5	6	7	8
N(1)-S(1)	2.1291(4)	2.134(2)	2.1350(4)	2.1341(18)
Ni(1)-S(2)	2.1340(4)	2.134(2)	2.1343(4)	2.1404(17)
Ni(1)-S(5)		2.137(2)		
Ni(1)-S(6)		2.138(2)		
S(1)-C(1)	1.7147(15)	1.705(8)	1.7143(15)	1.719(7)
S(2)-C(2)	1.7153(15)	1.716(9)	1.7084(16)	1.708(7)
S(3)-C(1)	1.7406(15)	1.736(9)	1.7393(16)	1.748(7)
S(4)-C(2)	1.7406(15)	1.733(9)	1.7418(15)	1.750(7)
S(5)-C(9)		1.715(9)		
S(6)-C(10)		1.699(9)		
S(7)-C(9)		1.739(8)		
S(8)-C(10)		1.755(9)		

Table 5. Cont.

	5	6	7	8
S(3)-C(3)	1.8239(15)	1.819(9)	1.8486(15)	1.829(7)
S(4)-C(4)	1.8126(15)	1.823(9)	1.8291(16)	1.820(7)
C(1)-C(2)	1.386(2)	1.413(12)	1.389(2)	1.380(10)
C(3)-C(4)	1.524(2)	1.549(12)	1.525(2)	1.508(10)
S(7)-C(11)		1.818(9)		
S(8)-C(12)		1.832(9)		
C(9)-C(10)		1.400(12)		
C(11)-C(12)		1.523(12)		
S(1)-Ni(1)-S(2)	91.652(14)	91.79(10)	91.959(15)	91.68(7)
S(5)-Ni(1)-S(6)		91.50(9)		
S(1)-Ni(1)-S(2') ^{a)}	88.347(14)		88.040(15)	88.32(7)
S(1)-Ni(1)-S(5)		88.79(9)		
S(1)-Ni(1)-S(6)		87.97(9)		
C(1)-C(2)-S(2)	119.03(11)	118.3(7)	119.86(12)	119.5(5)
C(1)-C(2)-S(4)	127.43(12)	127.8(7)	122.23(12)	126.2(5)
C(2)-C(1)-S(1)	119.29(11)	119.4(7)	118.98(12)	119.5(5)
C(2)-C(1)-S(3)	126.84(12)	125.6(7)	126.44(12)	126.6(5)
C(9)-C(10)-S(6)		120.3(7)		
C(9)-C(10)-S(8)		125.6(7)		
C(10)-C(9)-S(5)		117.8(6)		
C(10)-C(9)-S(7)		128.3(7)		

^{a)} The S(2') atoms were positioned by the following symmetry operations of S(2) atoms: (2-x, 1-y, 1-z) for **5**, (1-x, 1-y, -z) for **7**, and (-x, 1-y, 1-z) for **8**.

Figure 3 presents the molecular structure of **5**. One half of the trans-[Ni(cis-C5-dddt)₂] molecule was crystallographically independent with an inversion center on the nickel atom in the crystals of **5**. The torsion angles around the fused cyclopentylene group of the molecule were 45.93°, 42.06°, and -62.81° for C(1)-S(3)-C(3)-C(4), C(2)-S(4)-C(4)-C(3), and S(3)-C(3)-C(4)-S(4), respectively. The approximate geometry of the [Ni(dddt)₂] moiety in crystal **5** was highly similar to those of an anionic [Ni(dddt)₂]⁻ molecule [92]. The cyclopentylene ring of **5** adopted a folded conformation from the [Ni(dddt)₂] moiety with the trans-orientation, which is different from the conformation observed in (Ph₄P){trans-[Ni(cis-C5-dddt)₂]} (an envelope conformation) [93].

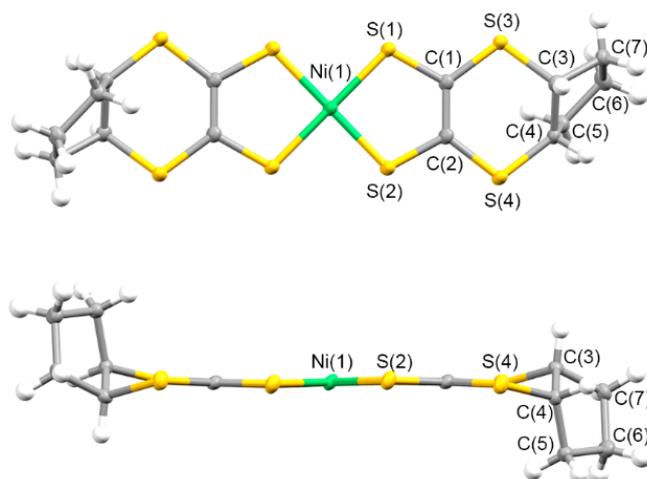


Figure 3. Molecular structure of **5** along the vertical and parallel directions on the molecular plane, with atom labels.

The packing diagrams of **5** are shown in Figure 4. Due to the folded geometry, the

columnar structures of the complex were restricted. The shortest Ni–Ni distance was 6.283 Å, which was the length of the *a* axis. The intermolecular S⋯S contacts shorter than the sum of the van der Waals radius (3.70 Å) were observed between the S(1)–S(1) atoms (3.574 Å) and S(1)–S(4′) atoms (3.604 Å), which formed a three-dimensional network structure in crystal 5.

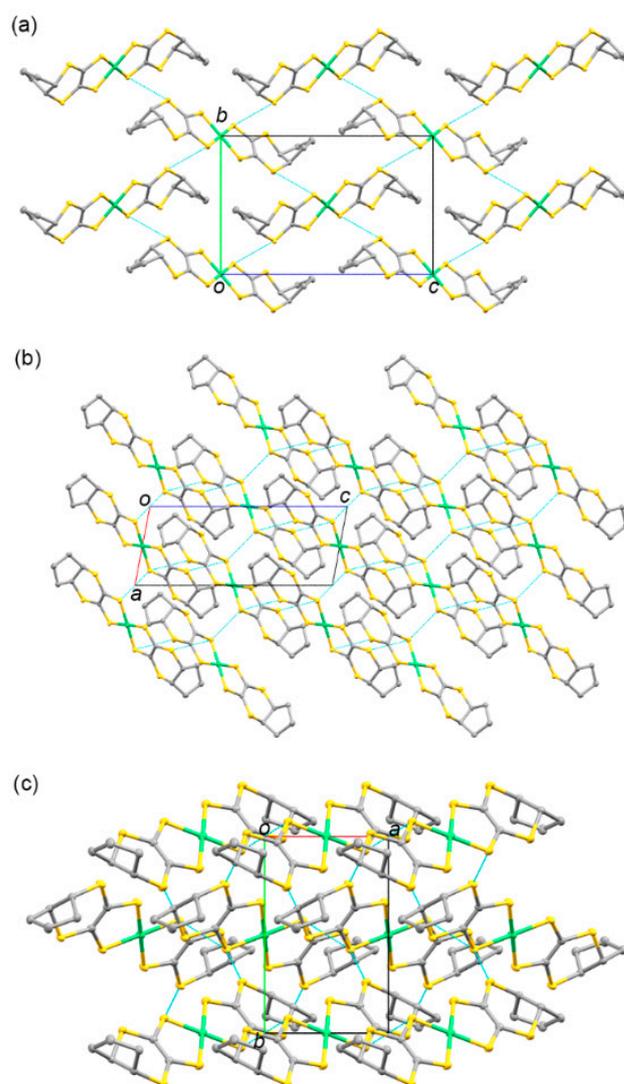


Figure 4. Crystal structures of 5 viewed along the (a) *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity. The blue dashed lines indicate the intermolecular S⋯S contacts shorter than the sum of the van der Waals radius (3.70 Å).

The molecular structure of compound 6 is shown in Figure 5. One trans-[Ni(cis-C6-dddt)₂] molecule was present in the asymmetric unit of the crystal. The torsion angles around the fused cyclohexylene group of the molecule were 43.79°, −68.30°, and −62.81° for C(1)–S(3)–C(3)–C(4), C(2)–S(4)–C(4)–C(3), and S(3)–C(3)–C(4)–S(4), respectively. The corresponding torsion angles in the opposite site of the molecule were −45.87°, −47.04°, and 66.65° for C(9)–S(7)–C(11)–C(12), C(10)–S(8)–C(12)–C(11), and S(7)–C(11)–C(12)–S(8), respectively. The stable chair-type configurations of the cyclohexylene groups were observed. However, the folded conformations of the cyclohexylene groups from the planar [Ni(dddt)₂] core were different from those of L4, which exhibited larger flexures around the terminal S–C bonds (Figure 1).

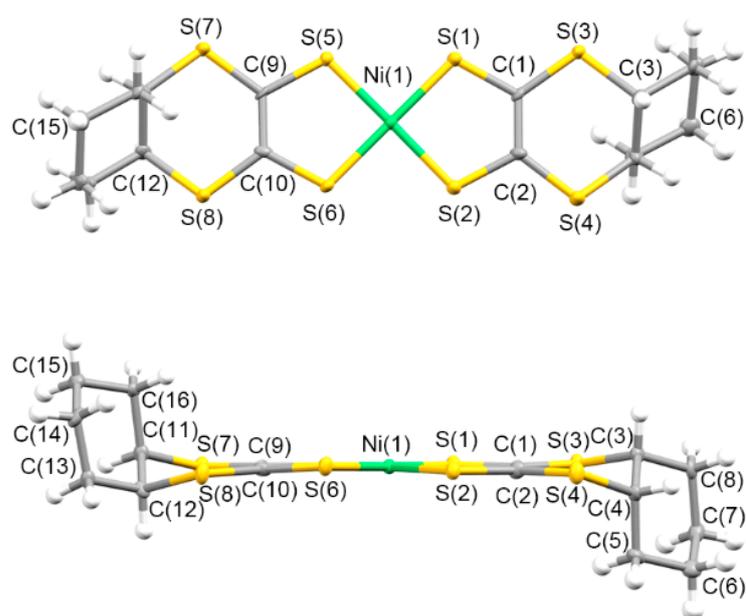


Figure 5. Molecular structure of **6** along the vertical and parallel directions on the molecular plane with atom labels.

A similar molecular geometry of $\text{trans-[Ni(cis-C6-dddt)}_2\text{)]}$ was reported by Bai et al. [94]. However, the packing motif of crystal **6** was significantly different from that of the reported crystal constructed by the centrosymmetric molecular arrangement of $\text{trans-[Ni(cis-C6-dddt)}_2\text{)]}$ with an inversion center on the nickel atom. In the case of crystal **6**, a non-centrosymmetric molecular arrangement with a helix axis along the b axis formed by the $\text{trans-[Ni(cis-C6-dddt)}_2\text{)]}$ molecules was observed (Figure 6). The cyclohexylene rings prevented the formation of stacks of the complexes and intermolecular interactions through the $\text{S}\cdots\text{S}$ contacts. The shortest intermolecular $\text{S}\cdots\text{S}$ distance was 3.749 Å between S(1) and S(8). The shortest Ni–Ni distance was 8.030 Å, which was the length of the a axis.

One half of the $\text{trans-[Ni(cis-C7-dddt)}_2\text{)]}$ molecule was present in the asymmetric unit in Crystal **7** (Figure 7). The twist chair-type configurations of the cycloheptylene groups in **L3** were maintained in the crystals after the coordination reaction of **L3'** and the nickel source (Scheme 3). The twist chair-type configurations of the cycloheptylene groups around the terminal S–C bonds in the molecule exhibited larger flexures (Figure 2), with torsion angles around the fused cycloheptylene group of 10.13° , 68.04° , and -55.70° for C(1)–S(3)–C(3)–C(4), C(2)–S(4)–C(4)–C(3), and S(3)–C(3)–C(4)–S(4), respectively. Figure 8 presents the crystal structure of **7**. The cycloheptylene moieties were in contact along the c axis. There was no intermolecular $\text{S}\cdots\text{S}$ contact, given that the $[\text{Ni(dddt)}_2]$ cores were completely separated by the steric hindrance of the cycloheptane moieties.

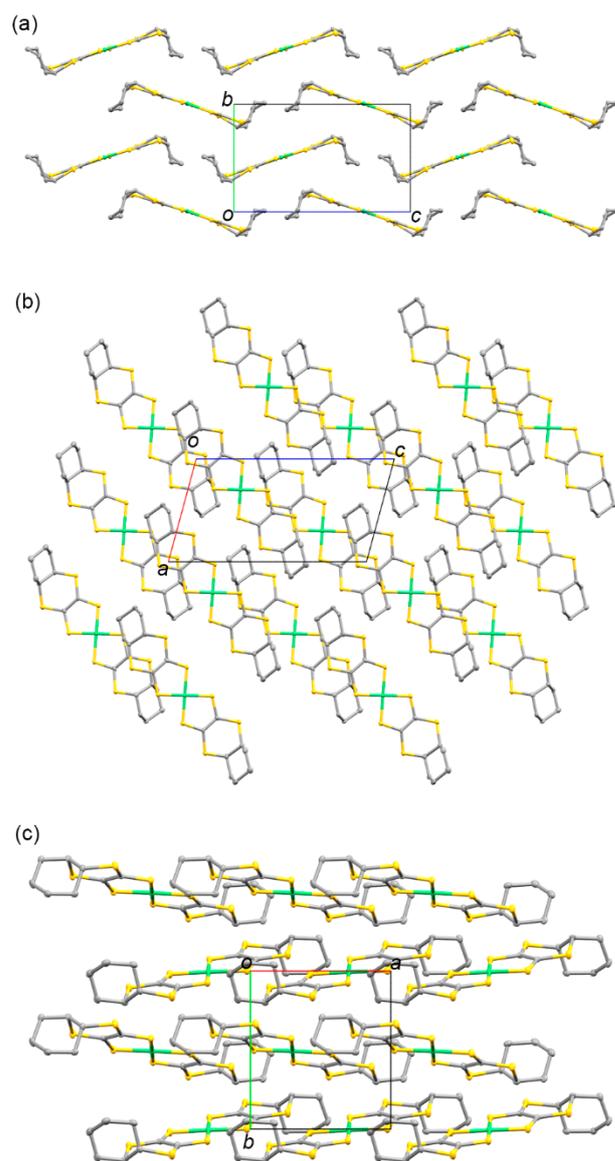


Figure 6. Crystal structure of 6 along the (a) *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity.

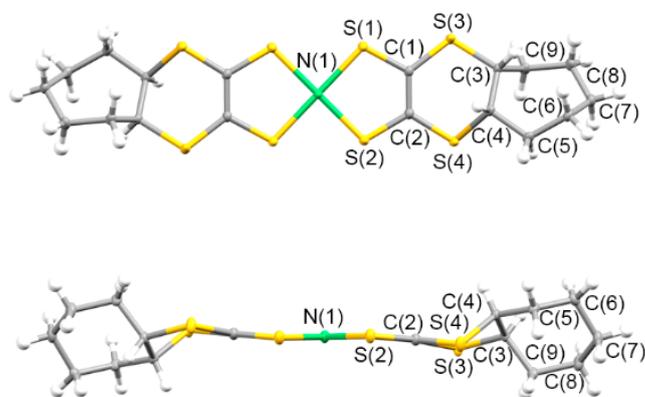


Figure 7. Molecular structure of 7 along the vertical and parallel directions on the molecular plane with atom labels.

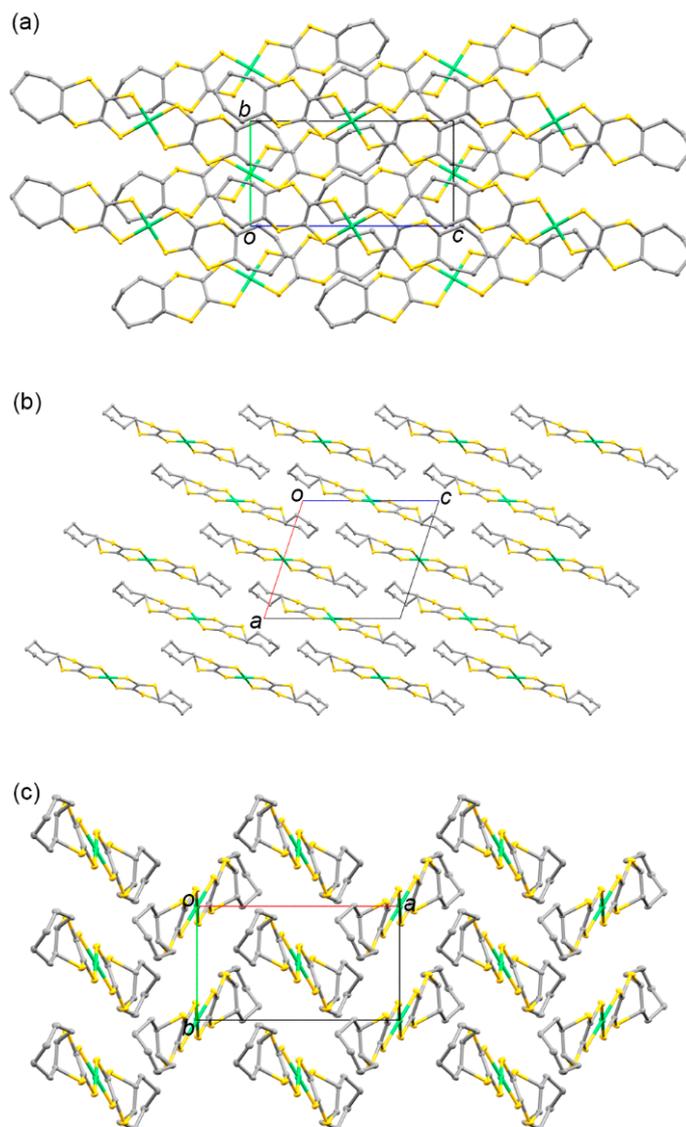


Figure 8. Crystal structure of 7 along the (a) *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity.

Figures 9 and 10 present the molecular and crystal structures of 8. The centrosymmetric crystal was formed by one-half of the $\text{trans-[Ni(cis-C8-dddt)}_2\text{)]}$ molecule in the asymmetric unit. The boat-chair-type configurations of the cyclooctylene groups formed in **L4** remained in molecule 6 with torsion angles of 54.74° , 46.88° , and -70.84° for C(1)-S(1)-C(3)-C(4), C(2)-S(4)-C(4)-C(3), and S(3)-C(3)-C(4)-S(4), respectively. One-dimensional columns along the *a* axis were constructed by ring-over-ring overlaps between the molecules sliding to the molecular longitudinal axis. There were no intercolumnar interactions through the S-S contacts. The shortest distance was 3.775 \AA between S(2) and S(3') in the column. The columns interacted through the side-by-side intermolecular S...S contacts between S(3) and S(3') with a distance of 3.444 \AA .

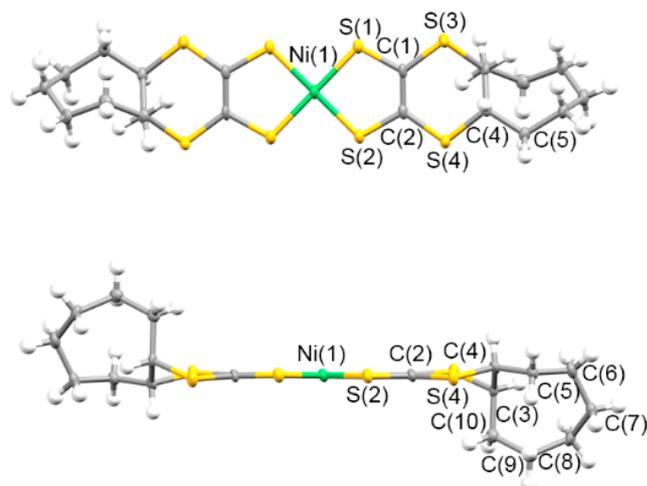


Figure 9. Molecular structure of **8** along the vertical and parallel directions on the molecular plane with atom labels.

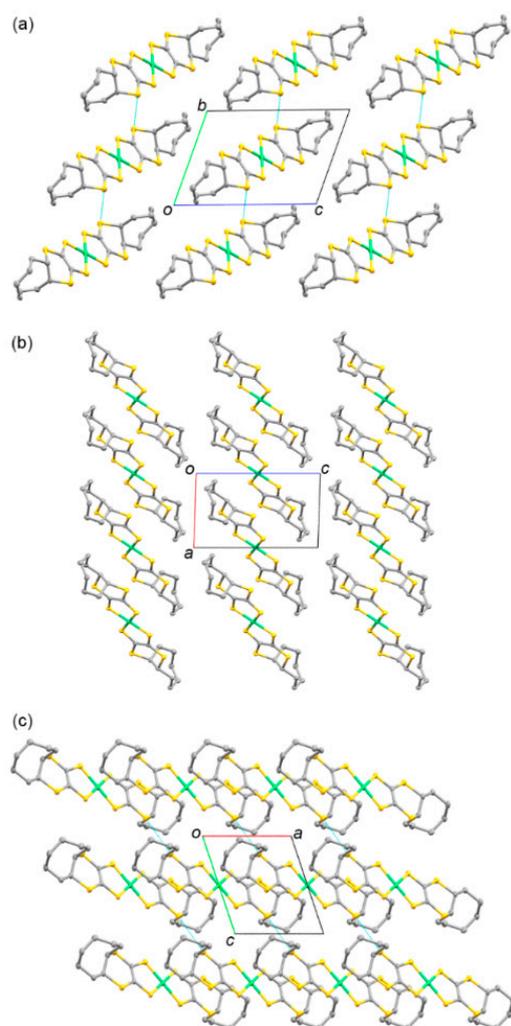


Figure 10. Crystal structure of **8** along the (a) *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity. The blue dashed lines indicate the intermolecular S...S contacts shorter than the sum of the van der Waals radius (3.70 Å).

3.2. Electrochemical Properties of 1–4

Cyclic voltammograms of monoanionic complexes 1–4 and $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$ in MeCN were obtained (Figure 1). The voltammograms of $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$ exhibited two quasi-reversible couples and one irreversible redox couple in the recorded region (Figure 1). The quasi-reversible redox peaks were assigned to $[\text{Ni}(\text{dddt})_2]^0/[\text{Ni}(\text{dddt})_2]^+$ ($E_{1/2}^1 = -1.24$ V) and $[\text{Ni}(\text{dddt})_2]^-/[\text{Ni}(\text{dddt})_2]^0$ ($E_{1/2}^2 = -1.02$ V) (Figure 11a). An irreversible oxidation peak was observed at 0.28 V, which indicates a reduction process between the dianionic $[\text{Ni}(\text{dddt})_2]^{2-}$ and monoanionic $[\text{Ni}(\text{dddt})_2]^{2-}$ states. The electrochemical features of 1–4 were similar to those of $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$. The redox potentials $E_{1/2}^1$ ($[\text{Ni}(\text{C}_n\text{-dddt})_2]^0/[\text{Ni}(\text{C}_n\text{-dddt})_2]^+$) and $E_{1/2}^2$ ($[\text{Ni}(\text{C}_n\text{-dddt})_2]^-/[\text{Ni}(\text{C}_n\text{-dddt})_2]^0$) of 1–4 are summarized in Table 6. Potentials E_1 and E_2 were correlated to the donor abilities of the complexes. The ΔE ($= E_{1/2}^2 - E_{1/2}^1$) values could be considered a measure of the on-site Coulomb energies [47,64]. Based on the $E_{1/2}^1$, $E_{1/2}^2$, and ΔE values, the complexes with fused cycloalkanes maintained the donor ability and on-site Coulomb energy of $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$.

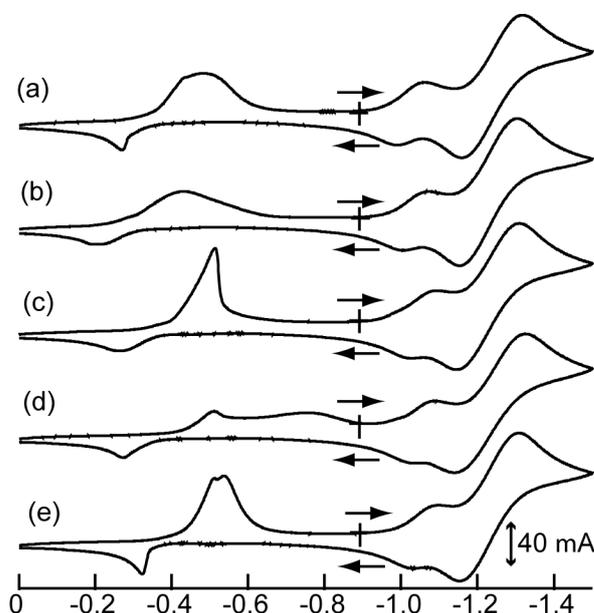


Figure 11. Cyclic voltammograms of (a) $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$, (b) 1, (c) 2, (d) 3, and (e) 4. Conditions: MeCN (1×10^{-3} M), supporting electrolyte: $(\text{Bu}_4\text{N})(\text{ClO}_4)$ (0.1 M), scan rate: 100 mV/s, reference electrode: Ag/Ag⁺, counter electrode: Pt wire, working electrode: glassy carbon at room temperature.

Table 6. Redox potentials $E_{1/2}^1/\text{V}$ (0, −1), $E_{1/2}^2/\text{V}$ (−1, −2), and ΔE ($= E_{1/2}^2 - E_{1/2}^1$)/V of $(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$ and Complexes 1–4.

	$E_{1/2}^1/\text{V}$	$E_{1/2}^2/\text{V}$	$\Delta E/\text{V}$
$(\text{Bu}_4\text{N})[\text{Ni}(\text{dddt})_2]$	−1.24	−1.02	0.22
1	−1.23	−1.03	0.20
2	−1.23	−1.05	0.18
3	−1.23	−1.05	0.18
4	−1.23	−1.06	0.17

Conditions: MeCN (1×10^{-3} M), supporting electrolyte: $(\text{Bu}_4\text{N})(\text{ClO}_4)$ (0.1 M), scan rate: 100 mV/s, reference electrode: Ag/Ag⁺, counter electrode: Pt wire, working electrode: glassy carbon at room temperature.

3.3. Electronic Absorption Spectra of 1–8

The electronic absorption spectra of monoanionic complexes 1–4 and neutral complexes 5–8 in CH_2Cl_2 were measured. Monoanionic complexes 1–4 exhibited a strong and broad absorption band at approximately 1036–1201 nm in the near-infrared region, which

can be attributed to a π - π^* transition between their HOMOs and LUMOs (Table 7) [95,96]. With respect to neutral complexes 5–8, the HOMO–LUMO absorption bands were shifted to the higher energy region (982–1040 nm), which corresponds to the energy gap between HOMO-1 and the HOMOs of the monoanionic complexes due to the weak shoulder absorptions observed at approximately 1000 nm. For example, monoanionic complex 4 exhibited two absorptions at 1201 nm and 1040 nm, which can be attributed to the HOMO–LUMO and HOMO-1 to HOMO gaps, respectively (Figure 12). The HOMO–LUMO gaps of the neutral complexes estimated by the spectra were within the range 1.19–1.26 eV. The spectroscopic behaviors of the complexes were similar to those of other bis(ethylene-1,2-dithiolato) complexes and $M(S_2C_2R_2)_2$ [85–89,93,94,97,98].

Table 7. Wavelength at absorption maxima (λ_{max}) and absorption coefficients (ϵ) for the π - π^* transitions of 1–8 in CH_2Cl_2 .

	λ_{max}/nm	Energy gap/eV	$\epsilon/dm^3 mol^{-1}$
1	1036	1.19	12,000
2	1198	1.03	10,000
3	1144	1.08	13,000
4	1201	1.03	13,000
5	982	1.26	32,000
6	1035	1.20	32,000
7	1031	1.20	39,000
8	1040	1.19	49,000

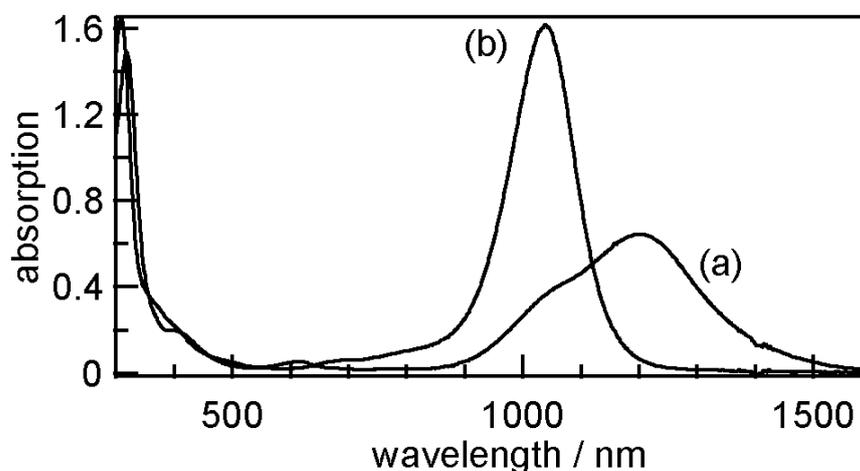


Figure 12. Electronic absorption spectra of (a) 4 (4.82×10^{-5} M) and (b) 8 (3.40×10^{-5} M) in CH_2Cl_2 solutions.

3.4. Structures and Physical Properties of Radical Cation Crystals 9 and 10

3.4.1. Molecular and Crystal Structures of 9 and 10

The electrochemical oxidation of CH_2Cl_2 solutions, including neutral complex 5 and the corresponding supporting electrolytes $(Bu_4N)X$ ($X = ClO_4^-$ and PF_6^-), resulted in the crystallization of 2:1 $[Ni(cis-C5-dddt)_2]_2(X)$ salts. The donor/acceptor ratios of the 2:1 crystals were different from those of the previously reported radical cation crystals based on the $[Ni(dddt)_2]$ complex without bulky substituents, constructing the 3:2 crystals of $[Ni(dddt)_2]_3(A)_2$ ($A = AuBr_2^-, BF_4^-, ClO_4^-, HSO_4^-,$ and $HSeO_4^-$ [55,56,99–103]). The crystals of 9 and 10 included one crystallographically independent donor and half of the ClO_4^- or PF_6^- anions, respectively. Figure 13 and (Supplementary Materials) Figure S1 present the molecular structures of the donors in 9 and 10. In the case of the radical cation crystals, the *cis*- $[Ni(cis-C5-dddt)_2]$ conformation was formed preferentially, although the corresponding neutral complex exhibited the *trans*- $[Ni(cis-C5-dddt)_2]$ conformation

(Scheme 4). The donors exhibited a square planar coordination around the Ni atoms. Moreover, the bond distances of Ni-S = 2.1438(5)–2.1481(5) Å, S-C = 1.703(2)–1.7077(19) Å, and C=C = 1.406(3)–1.408(3) Å for **9**; and Ni-S = 2.1435(6)–2.1498(6) Å, S-C = 1.702(2)–1.705(2) Å, and C=C = 1.404(3)–1.414(3) Å for **10** in the NiS₄C₄ framework were slightly larger than those of the neutral trans-[Ni(cis-C5-dddt)₂]. Figure 14 and Figure S2 present the crystal structures of **9** and **10**. In the crystals, similar molecular arrangements of the donors and anions were observed with a two-fold axis on one of the Cl-O and P-F bonds of the anions along the *b* axis, respectively. In **9**, the perchlorate anions were disordered at the two sites by rotation around the two-fold axis. The cis-[Ni(cis-C5-dddt)₂] conformation would induce dimerized structures of the donors in **9** and **10**, which differ from the molecular arrangement in the cases of the 3:2 crystals of [Ni(dddt)₂]₃(A)₂ significantly. The dimerized structures of the donors were formed by the S⋯S contacts (3.665 Å and 3.664 Å for **9** and **10**, respectively), and the dimers were arranged into columnar structures through the S⋯S contacts (3.682 Å and 3.556 Å for **9** and **10**, respectively) along the *a*-*c* axis directions in **9** and **10**. Intermolecular side-by-side interactions through the S⋯S contacts (3.507–3.637 Å and 3.500–3.629 Å for **9** and **10**, respectively) between the columns were observed along the *b* axis, thus providing two-dimensional donor sheets parallel to the *bc* planes. Although the donor molecule is fused with the bulky cyclopentane moieties, the two-dimensional donor sheet arrangements through the S⋯S contacts preferentially are constructed in the radical cation crystals, which are also observed in the organic radical cation crystals based on the BEDT-TTF and BETS molecules exhibiting metallic and superconducting behaviors [14–18].

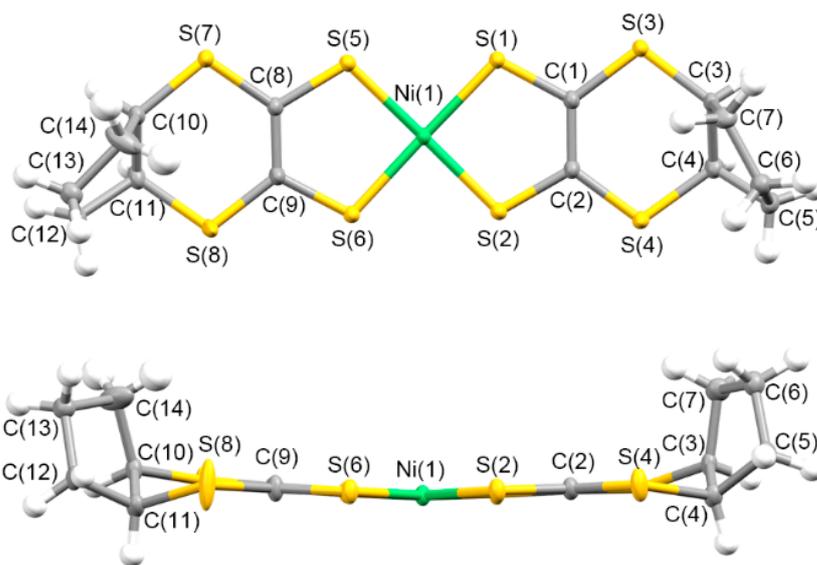


Figure 13. Molecular structure of **9** along the vertical and parallel directions on the molecular plane with atom labels.

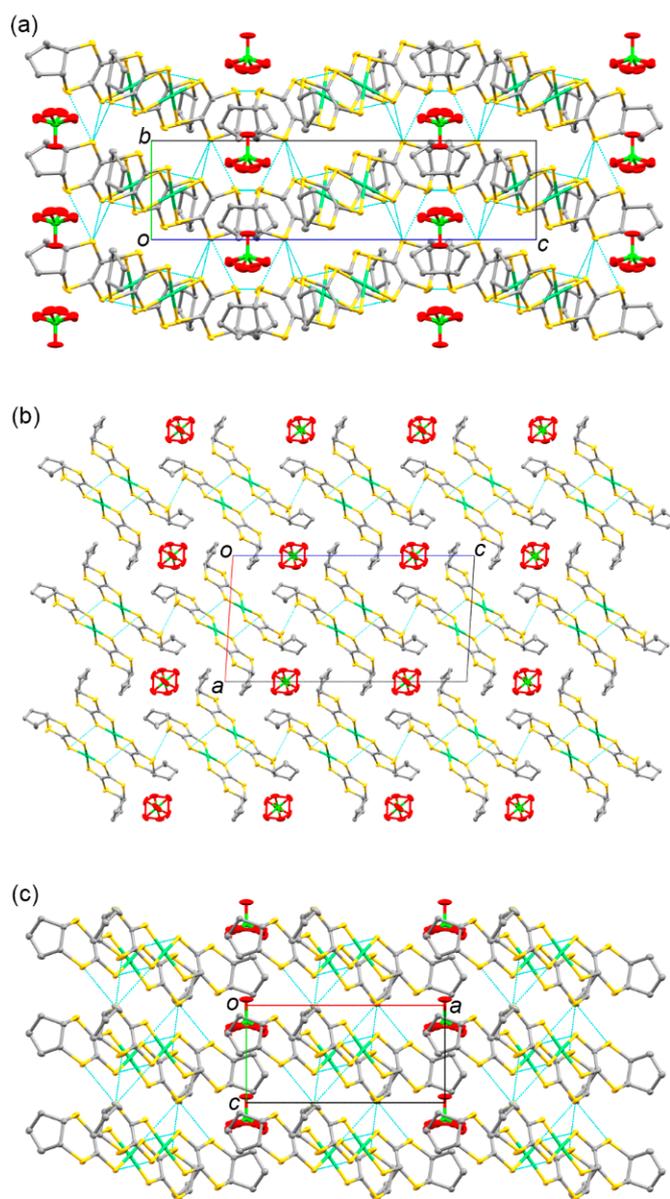


Figure 14. Crystal structure of **9** along (a) the *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity. The dashed lines indicate sulfur–sulfur distances shorter than the sum of the van der Waals radius ($<3.7 \text{ \AA}$).

3.4.2. Molecular Orbitals and Energy Band Calculations of **9** and **10**

To reveal the electronic features of the salts, molecular orbitals (MO) and energy band structures were determined using the extended Hückel and tight-binding methods. Figure 15 presents the determined HOMOs and LUMOs of the donor molecules of **9** and **10** in their neutral states. The highly symmetric HOMOs and LUMOs with the nickel *d* orbitals of the donors were in good agreement with the molecular orbitals of the $[M(\text{ddd}t)_2]$ complex calculated based on the first-principles density-functional theory [56–62].

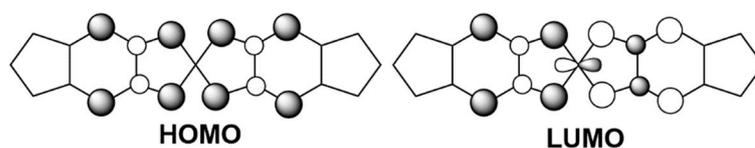


Figure 15. The HOMO and LUMO of $[\text{Ni}(\text{C}_5\text{-ddd}t)_2]$.

The calculated overlap integrals (S) of the crystals between the HOMOs are summarized in Table 8 with a schematic diagram of their donor arrangements. The symbols a1–b4 indicate directions of intermolecular interactions between the donors. Figure S3 presents the corresponding molecular arrangement of the donor molecules viewed along the end-on direction in the crystal. The significant dimerization of the crystals was observed, given that the face-to-face interactions (a2) were significantly stronger than the other interactions (a1) in the columns. Dominant side-by-side interactions (b1 and b4) connected the dimers, thus forming two-dimensional sheets in the crystals. Figure 16 presents the calculated band structure of **9**, which is similar to that of crystal **10**. Both the upper bands crossed the Fermi level, thus forming the small two-dimensional Fermi surfaces of the crystals.

Table 8. Overlap integrals (S) between the HOMOs and a schematic diagram of the donor arrangements for Crystals **9** and **10**.

$S (\times 10^3)$	9	10
a1	−2.7	−1.9
a2	−10	−10
b1	−6.9	−3.8
b2	−0.14	−0.7
b3	−0.14	−0.7
b4	1.2	−0.3

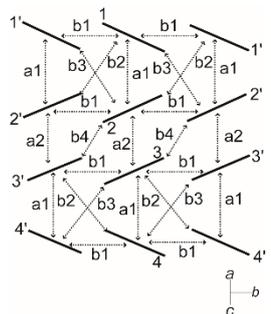
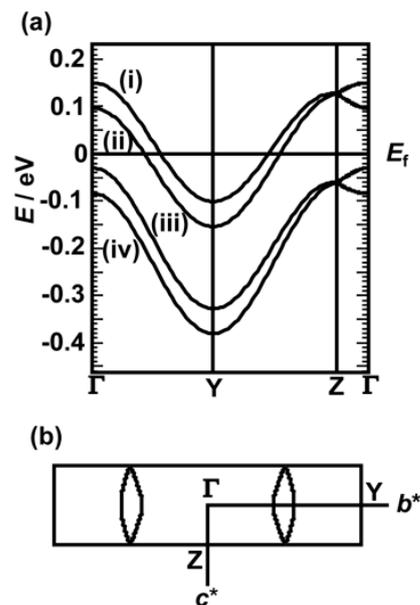



Figure 16. (a) Energy band structure and (b) Fermi surface of **9**. The characters b^* and c^* denote reciprocal lattice axes.

3.4.3. Electrical Conductivities and Magnetic Properties of **9** and **10**

The temperature dependence of the resistivities was measured for single crystals of **9** and **10** K using a standard four-probe direct current (DC) method (Figure 17). The salts exhibited semiconducting behaviors (**9**: $\rho_{rt} = 0.8 \Omega \text{ cm}$ and $E_a = 0.09 \text{ eV}$; **10**: $\rho_{rt} = 4.0 \Omega \text{ cm}$ and $E_a = 0.13 \text{ eV}$). Although the energy band calculations for **9** and **10** suggested a metallic nature, crystals **9** and **10** exhibited semiconducting behaviors under ambient pressure. The results of the resistivity measurements and the energy band calculations indicate that the crystals form a Mott insulating state or a spin-singlet state under ambient pres-

sure [104]. In order to reveal the electronic structures of the crystals, temperature dependence of the resistivities under hydrostatic pressures and magnetic susceptibilities of **9** and **10** were measured. The resistivity measurements under various hydrostatic pressures (0.3–1.8 GPa) exhibited slightly smaller activation energies than those under ambient pressure (**9**: $\rho_{rt} = 0.2 \Omega \text{ cm}$ and $E_a = 0.07 \text{ eV}$; **10**: $\rho_{rt} = 1.0 \Omega \text{ cm}$ and $E_a = 0.12 \text{ eV}$ at 1.8 GPa). The temperature dependencies of the magnetic susceptibilities of **9** and **10** were in accordance with the Curie–Weiss law. However, significantly small χ values at 300 K indicate diamagnetic features ($\chi = 2.9 \times 10^{-5} \text{ cm}^3 \text{ mol}^{-1}$ for **9** and $\chi = 2.3 \times 10^{-5} \text{ cm}^3 \text{ mol}^{-1}$ for **10**). These results suggest that the crystals formed spin-singlet states between the dimers through the S··S contacts to the b1 direction (Table 8), which is in accordance with their conducting behaviors.

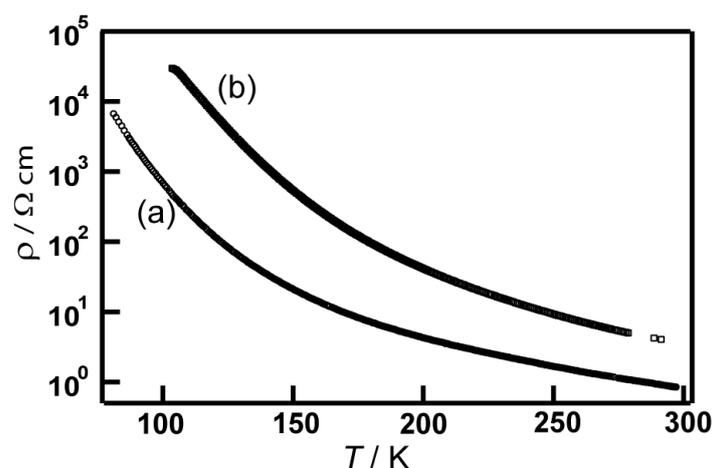


Figure 17. Temperature dependence of electrical resistivities of (a) **9** and (b) **10**.

4. Conclusions

New donor-type nickel–dithiolene complexes based on the $[\text{Ni}(\text{dddt})_2]$ skeleton fused with bulky cycloalkane substituents were synthesized as new components of molecular conductors. All the neutral crystals were crystallized in $\text{cis-}[\text{Ni}(\text{cis-C}_n\text{-dddt})_2]$ conformations with cis- oriented (*R,S*) conformations around the cycloalkylene groups due to the [2 + 4] cycloadditions. The flexible cycloalkane moieties in the molecules can induce unique molecular arrangements in their neutral crystals. The three-dimensional network and one-dimensional columnar structures were constructed through intermolecular S··S contacts in crystals **5** and **8**. No intermolecular S··S contacts were observed in **6** and **7**. However, $[\text{Ni}(\text{C}_6\text{-dddt})_2]$ formed a helical molecular arrangement in the crystal. New cation radical crystals $[\text{Ni}(\text{C}_3\text{-dddt})_2](\text{ClO}_4^- \text{ or } \text{PF}_6^-)$ with dimerized structures were successfully obtained by electrochemical crystallization, which formed different electronic structures of the $[\text{Ni}(\text{dddt})_2]_3(\text{X})_2$ crystals [55,57,99–103]. The radical cation crystals were insulators due to the spin-singlet states between the dimers through the strong intermolecular S··S contacts. However, the donors with cycloalkane substituents can adjust the proportions of the inter- and intra-dimer interactions through the S··S contacts in their radical cation crystals by modification of the cycloalkane moieties, and allow for the realization of various cation radical crystals with unique donor arrangements by introducing linear, tetrahedral, and octahedral anions. Moreover, neutral crystals are potential candidates for new single-component molecular conductors [58–63].

Supplementary Materials: The following are available online at <https://www.mdpi.com/article/10.3390/cryst11101154/s1>, Figure S1: Molecular structure of **10** along vertical and parallel directions on the molecular plane with atom labels. Figure S2: Crystal structure of **10** along the (a) *a*, (b) *b*, and (c) *c* axes. Hydrogen atoms are omitted for clarity. The dashed lines indicate sulfur–sulfur distances shorter than the sum of the van der Waals radius (<3.70 Å), CIFs: CCDC numbers 2103657–2103665 for **5–8**, **9**, and **10**.

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