Supporting Information

Copper-Silver Bimetallic Nanowire Arrays for Electrochemical Reduction of Carbon Dioxide

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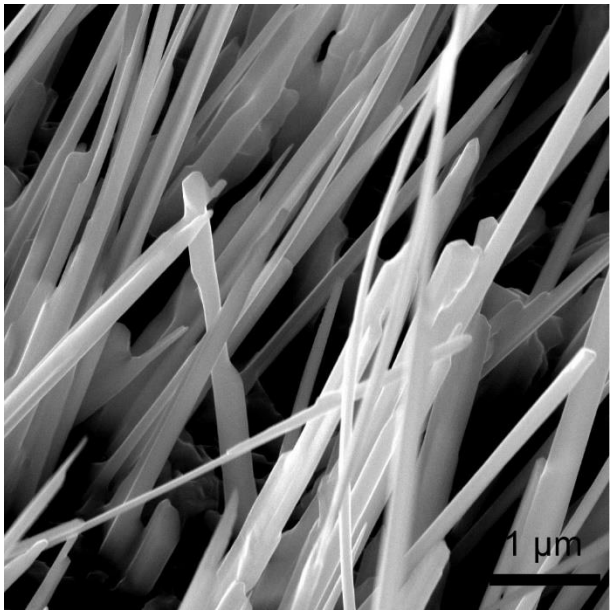

Figure S1. High-magnification SEM images of as-synthesized CuO/Cu_2O NWs.

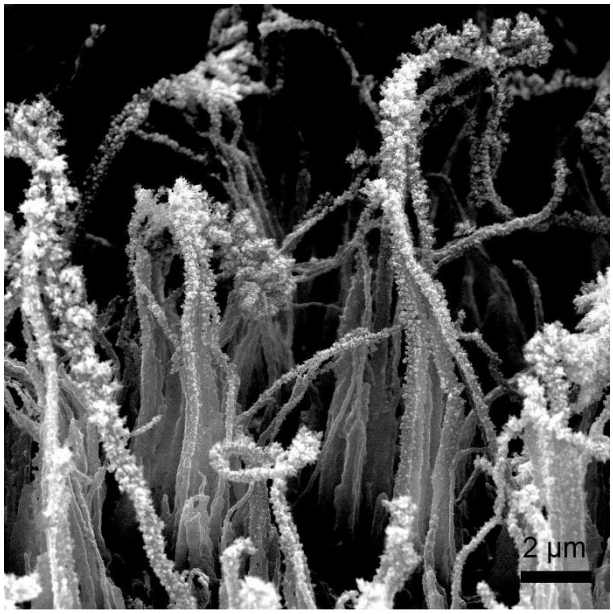

Figure S2. Low-magnification SEM images of as-synthesized Cu-Ag NWs.

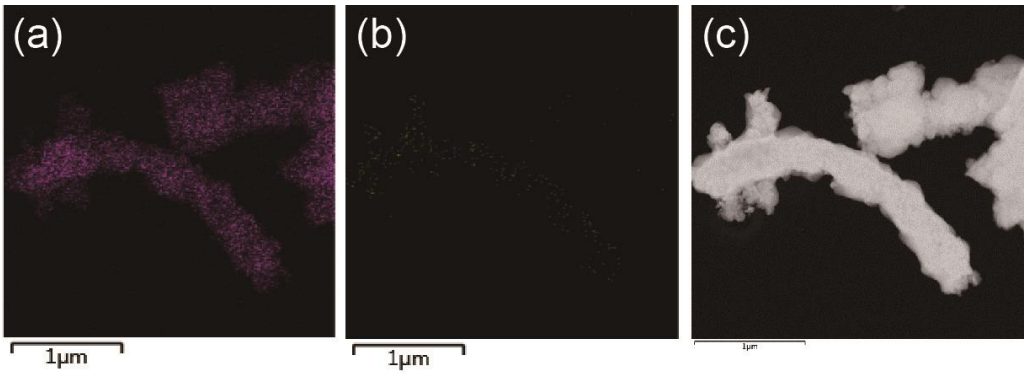

Figure S3. EDXS mapping figure of the Cu-Ag NWs (a) Cu distribution, (b) Ag distribution, and (c) TEM image.

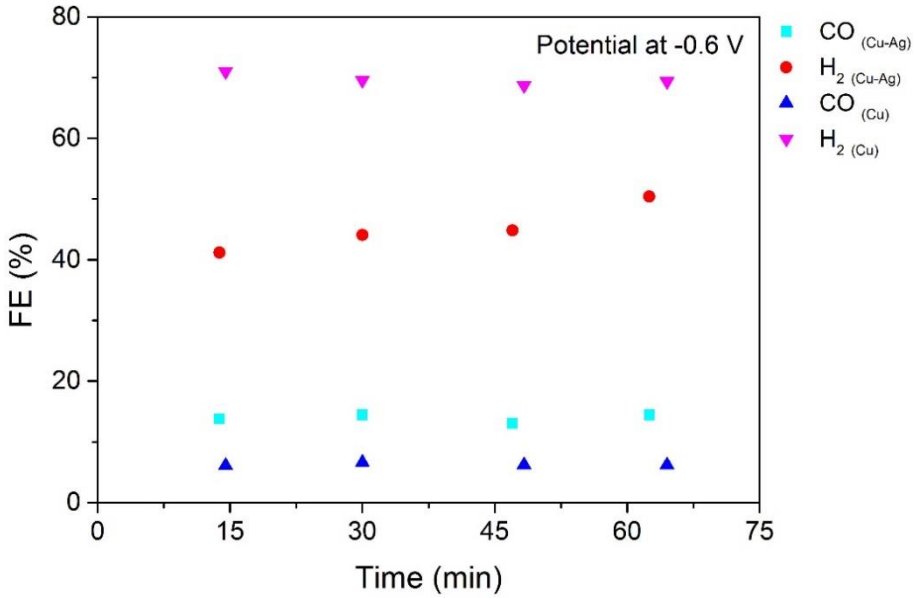

Figure S4. Time-dependent FE of major gaseous products for the Cu NWs and the Cu-Ag NWs at -0.6 V (vs. RHE).

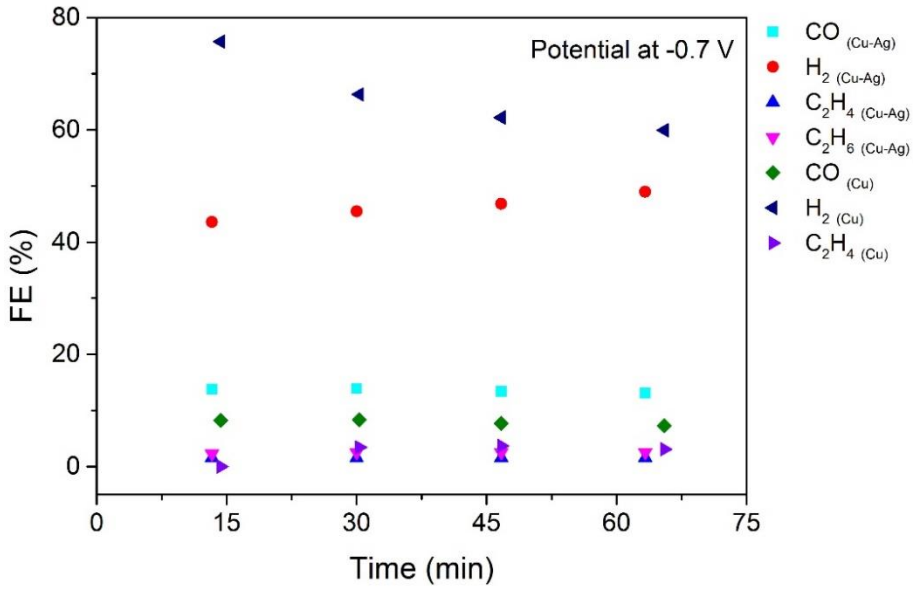

Figure S5. Time-dependent FE of major gaseous products for the Cu NWs and the Cu-Ag NWs at -0.7 V (vs. RHE).

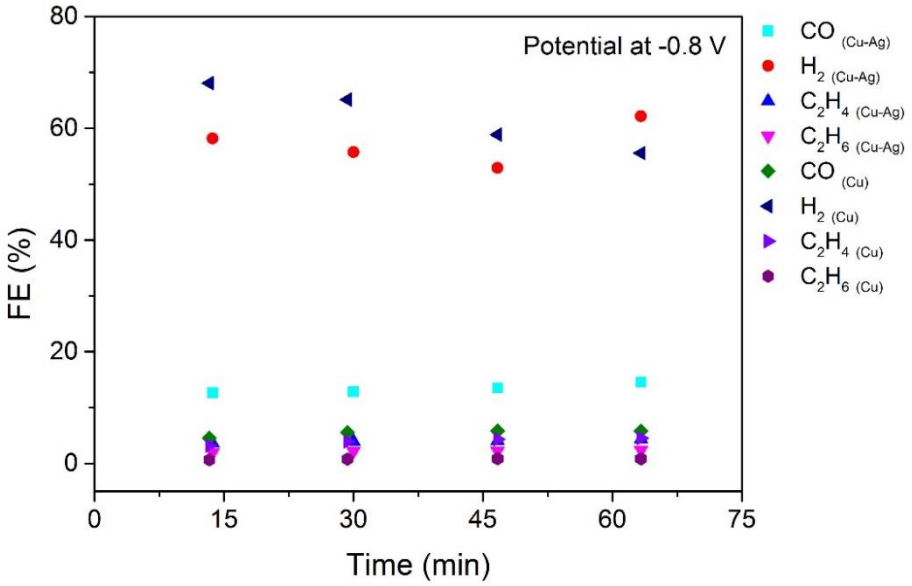

Figure S6. Time-dependent FE of major gaseous products for the Cu NWs and the Cu-Ag NWs at -0.8 V (vs. RHE).

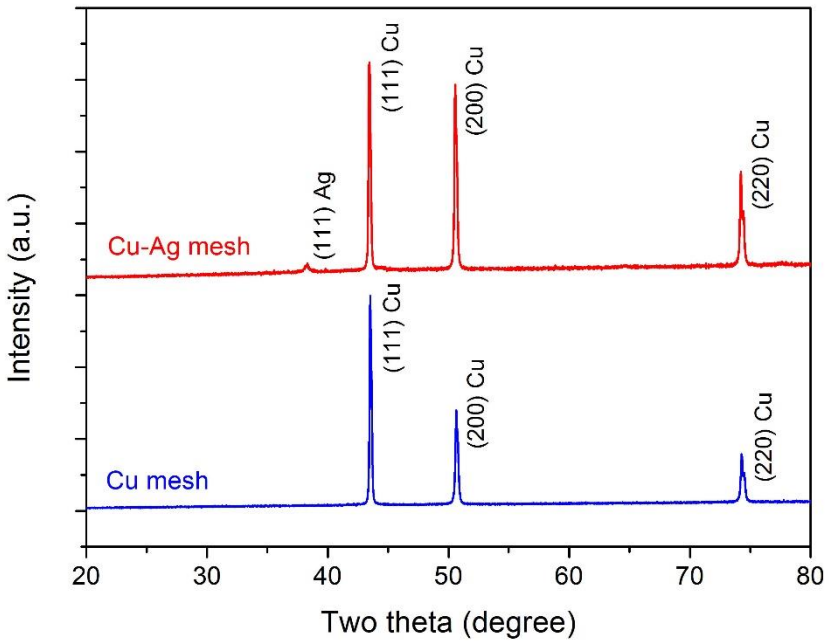

Figure S7. X-ray diffraction patterns of pristine copper mesh (blue) and Cu-Ag mesh (red).

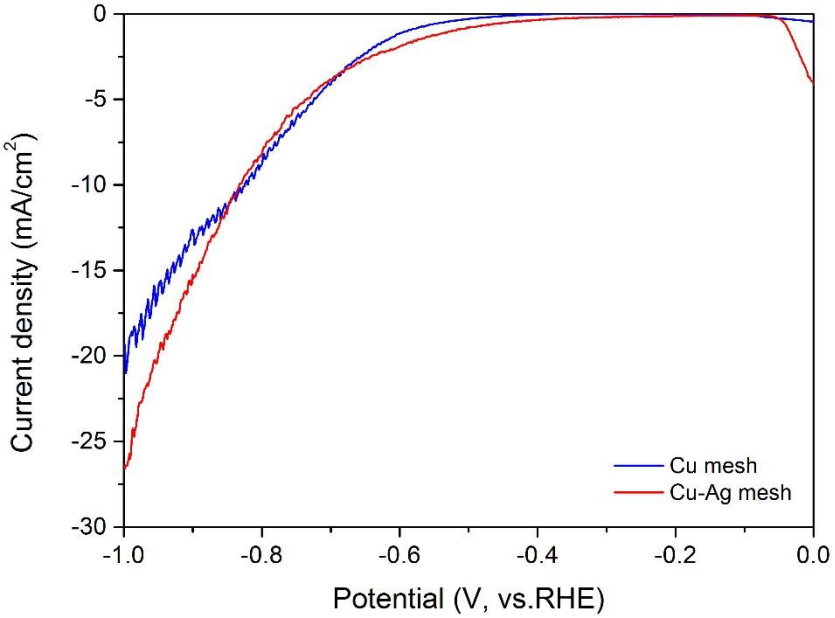

Figure S8. Linear sweep voltammetry of electrochemical CO_2 reduction of the Cu mesh and the Cu-Ag mesh.

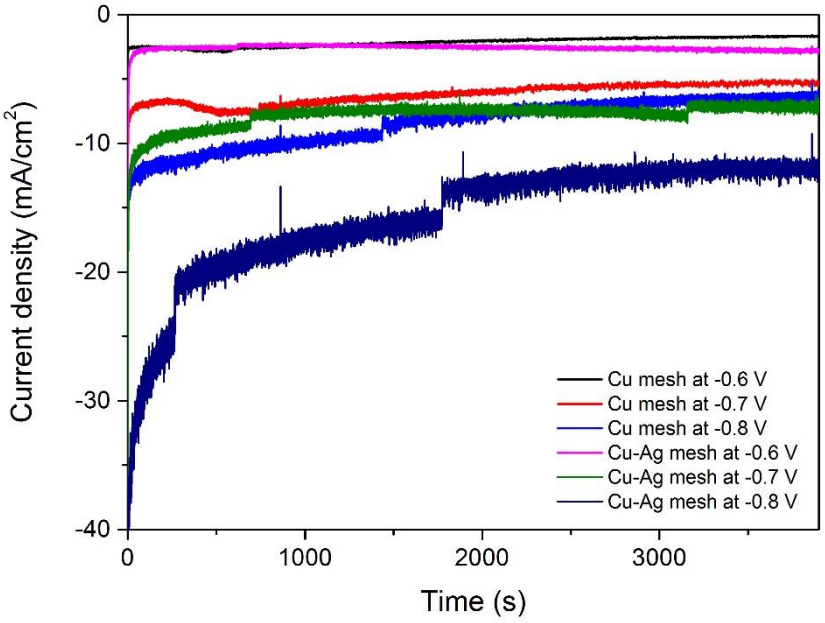

Figure S9. CO_2 reduction activity on the Cu mesh and the Cu-Ag mesh at different potentials.

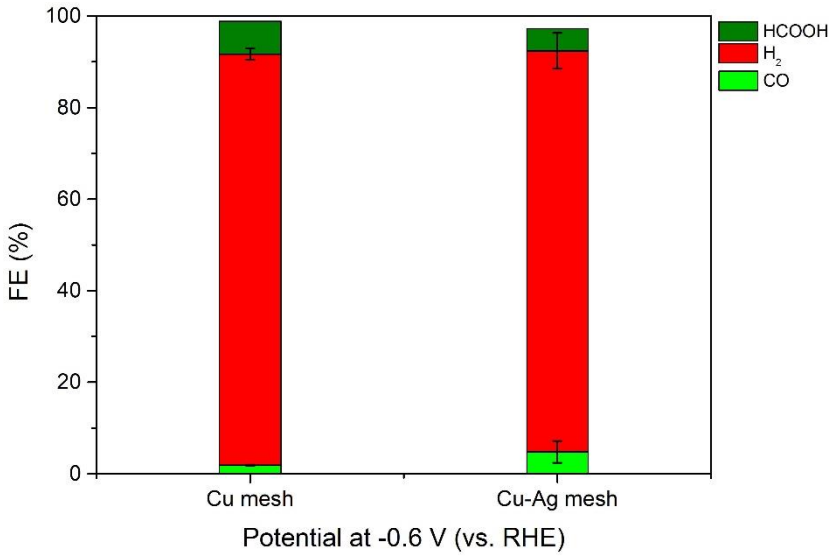

Figure S10. FE of the major products for the Cu mesh and the Cu-Ag mesh at -0.6 V (vs. RHE).

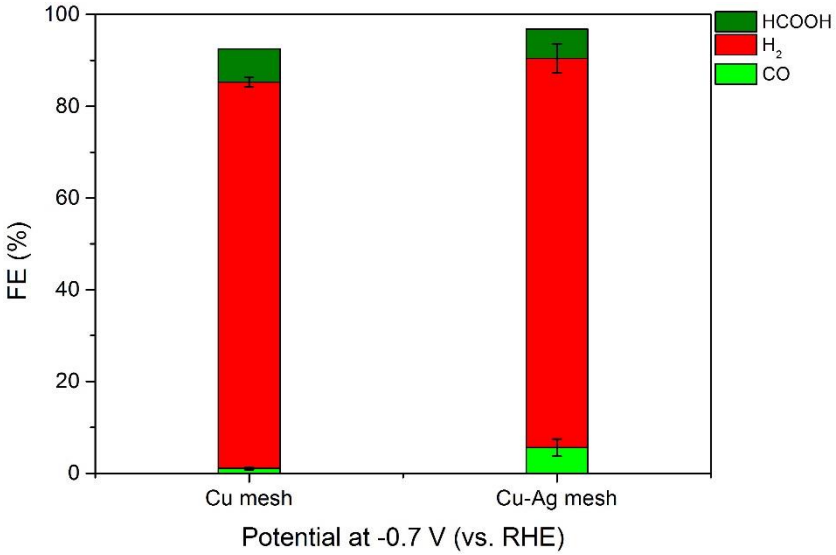

Figure S11. FE of the major products for the Cu mesh and the Cu-Ag mesh at -0.7 V (vs. RHE).

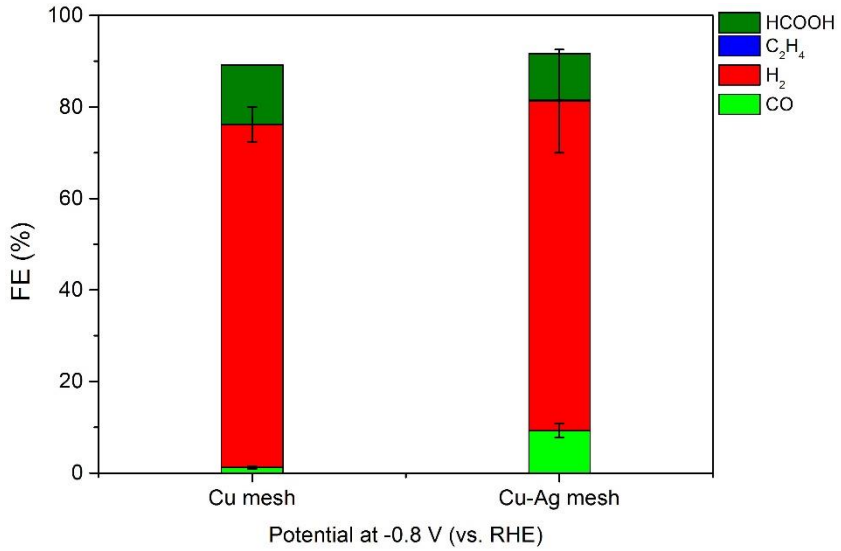

Figure S12. FE of the major products for the Cu mesh and the Cu-Ag mesh at -0.8 V (vs. RHE).