

Article

Improved Route to Linear Triblock Copolymers by Coupling with Glycidyl Ether-Activated Poly(ethylene oxide) Chains

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Abstract: Poly(ethylene oxide) block copolymers (PEO_z BCP) have been demonstrated to exhibit remarkably high lithium ion (Li⁺) conductivity for Li⁺ batteries applications. For linear poly(isoprene)-*b*-poly(styrene)-*b*-poly(ethylene oxide) triblock copolymers (PI_xPS_yPEO_z), a pronounced maximum ion conductivity was reported for short PEO_z molecular weights around 2 kg mol^{−1}. To later enable a systematic exploration of the influence of the PI_x and PS_y block lengths and related morphologies on the ion conductivity, a synthetic method is needed where the short PEO_z block length can be kept constant, while the PI_x and PS_y block lengths could be systematically and independently varied. Here, we introduce a glycidyl ether route that allows covalent attachment of pre-synthesized glycidyl-end functionalized PEO_z chains to terminate PI_xPS_y BCPs. The attachment proceeds to full conversion in a simplified and reproducible one-pot polymerization such that PI_xPS_yPEO_z with narrow chain length distribution and a fixed PEO_z block length of $z = 1.9$ kg mol^{−1} and a $\bar{D} = 1.03$ are obtained. The successful quantitative end group modification of the PEO_z block was verified by nuclear magnetic resonance (NMR) spectroscopy, gel permeation chromatography (GPC) and differential scanning calorimetry (DSC). We demonstrate further that with a controlled casting process, ordered microphases with macroscopic long-range directional order can be fabricated, as demonstrated by small-angle X-ray scattering (SAXS), scanning electron microscopy (SEM) and transmission electron microscopy (TEM). It has already been shown in a patent, published by us, that BCPs from the synthesis method presented here exhibit comparable or even higher ionic conductivities than those previously published. Therefore, this PEO_z BCP system is ideally suitable to relate BCP morphology, order and orientation to macroscopic Li⁺ conductivity in Li⁺ batteries.

Keywords: polymers; block copolymers; Polyethylene Oxide (PEO); convergent synthesis; epoxide; anionic polymerization; microphase separation; morphology

1. Introduction

As a polymer class, block copolymers (BCPs) are gaining continuous attention due to their remarkable properties such as an amphiphilic character or self-assembling ability [1–3]. For instance, an amphiphilic behavior enables in solution the formation of micelles, which are widely utilized in pharmaceutical applications, e.g., for drug delivery systems [1,2]. Moreover, their ability to self-assemble also in bulk paves the way to well-ordered morphologies, which find a wide range of applications [3–5], e.g., in lithography [6], semiconductor-based photocatalysis [4,7] and energy storage and conversion as fuel cell membranes [8], electrodes [9] or polymer electrolytes [10–12].

These aforementioned properties rely on the tailorable and unique structure of the respective BCP. The covalent binding of polar and nonpolar polymer blocks with defined block length could result in macromolecules with amphiphilic character and therefore tend to organize themselves into periodic, highly ordered, nano-sized domains, the so-called microphases [3–6,8,13–16]. The Flory–Huggins interaction parameter (χ) quantifies the incompatibility between the different blocks based on their interaction energy [13–16].

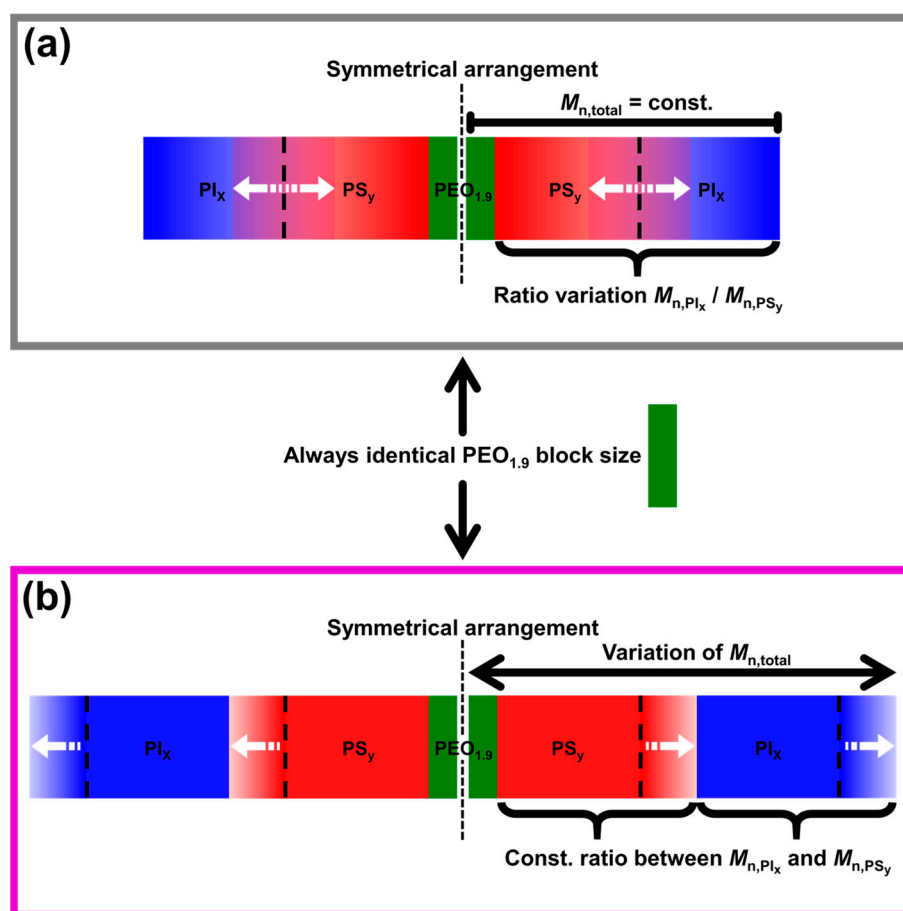
Our decision to choose the linear triblock copolymer poly(isoprene)-*b*-poly(styrene)-*b*-poly(ethylene oxide), denoted as $PI_xPS_yPEO_z$, whereby each index (x , y and z) indicates the molar mass (M_n) of the corresponding block in kg mol^{-1} , was inspired by the work of Dörr et al. [4,7–11]. They applied a synthetic route developed in the group of Bates [17–19] and demonstrated its advantageous use as a template for detailed control of mesoscopic porous 3D architectures with embedded inorganic materials [4,7–11]. In our case, the synthesized $PI_xPS_yPEO_z$ are supposed to be used as a structure-giving BCP matrix for conducting lithium ions (Li^+). Additionally, Dörr and Pelz et al. showed that lowering the PEO_z chain length down to values around 45 repeating units ($M_n \sim 2 \text{ kg mol}^{-1}$) and with a low content ($<4 \text{ vol.}\%$) in relation to total BCP size resulted in a high Li^+ conductivity [10,11].

Therefore, the PEO_z block has a significant influence on the Li^+ transport properties and is directly linked to it, as the Li^+ presumably only accumulate in this block, resulting in the formation of Li^+ conducting pathways [10,11,20]. Consequently, those Li^+ conducting pathways and, finally, the total Li^+ transport are substantially determined by the properties of the PEO_z block and its resulting domain structure such as its size, long-range order, morphology and macroscopic orientation. This means for an optimal Li^+ transfer, the structure-giving BCP matrix has to arrange itself into a long-range and highly ordered morphology, e.g., lamellar (LAM) or hexagonally close-packed cylindrical (HEX), continuously between two electrodes and an orientation connecting these electrodes. In order to obtain these properties in the $PI_xPS_yPEO_z$, the combination of always having precisely defined length and a nearly monodisperse distribution in each block, as well as a controlled self-assembly during the membrane preparation process, is important [21,22].

Hence, in this work, the influence of the structure-giving BCP matrix on the short PEO_z chain order will be investigated. Corresponding $PI_xPS_yPEO_z$ BCP will be synthesized by keeping the PEO_z chain length constant in order to vary its composition systematically and independently from each other (cf. Scheme 1) [18,19,23,24]. For this purpose, $PI_xPS_yPEO_z$ BCPs are varied in two different ways:

(1) By differing the ratio of the M_n of the PI_x to the PS_y block ($M_{n,PI_x}/M_{n,PS_y}$), while retaining the same PEO_z block proportion because the total M_n of the BCP ($M_{n,total}$) is hold constant (cf. Scheme 1a).

(2) By altering the $M_{n,total}$ and therefore the PEO_z block proportion, while keeping $M_{n,PI_x}/M_{n,PS_y} = \text{constant}$ (cf. Scheme 1b).



Scheme 1. Controlled compositional variations of the synthesized $PI_xPS_yPEO_z$ BCP by systematically and independently altering the PI_x and PS_y block lengths while keeping the PEO_z block constant. **(a)** Changing the ratio of the M_n of the PI_x to the PS_y block ($M_{n,PI_x}/M_{n,PS_y}$) by keeping the total M_n of the BCP ($M_{n,total}$) constant and therefore retaining identical PEO_z block proportion. **(b)** Variation of the PEO_z block proportion in the BCP by altering the $M_{n,total}$ and holding $M_{n,PI_x}/M_{n,PS_y} = \text{constant}$.

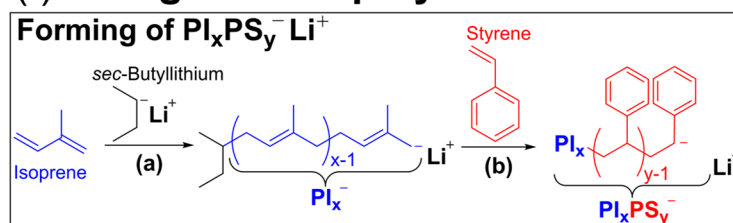
This will be achieved by a new developed synthesis route, which ensures the use of consistent identical and very short as well as commercially available prefabricated PEO_z chains for the attachment to BCPs.

Usually, such BCPs are prepared in a stepwise manner by synthesizing each polymer block sequentially [17–19,25]. Thus, each polymer block is formed from the respective monomers in a series of polymerization steps according to the principle of living sequential anionic polymerization (cf. Scheme 2I), as it offers the highest control over the polymerization process, hence the dispersity (D), and also proceeds without side reactions [5,17–19,26–29]. However, due to the explosive and highly toxic properties of ethylene oxide (EO) gas, its use implies special safety requirements [30]. Therefore, the use of a short prefabricated PEO_z block in our synthesis route leads to the fact that the handling of EO gas monomers during $PI_xPS_yPEO_z$ polymerization can be avoided. In this way, the necessary use of EO gas for the synthesis of PEO_z chains can be carried out in a separate and upstream synthesis step.

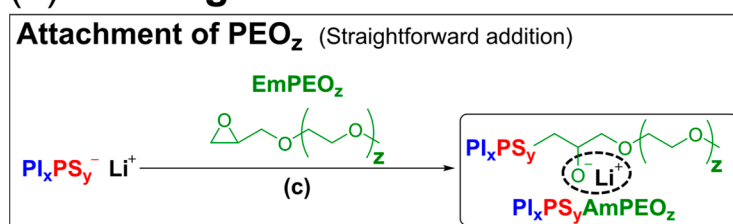
Thus, in this study, an exact defined methoxy PEO ($mPEO_z$) chain with a modified end group was chosen, enabling it accessible for direct and covalent attachment to the stable $PI_xPS_y^-$ carbanion of the living polymer chain. Considering the large variety of suitable end groups, tethering an epoxide end group to the $mPEO_z$ chain ($EmPEO_z$) enables a selective single one-step addition to the $PI_xPS_y^-$ anion by utilization of the strong Li-O interaction (cf. Scheme 2II) [5,17,31–34]. This is similar to the general strategy of using epoxides as terminating agents as reported in literature [35–41]. The strong interaction

between the hard oxygen anion and the hard Li^+ can be well explained based on the concept of “hard” and “soft” acids and bases (HSAB) [32,42]. In addition to the PEO_z chain, at the junction point only an extra alcohol group is introduced in the polymer as (poly(isoprene)-*b*-poly(styrene)-*b*-alcohol methoxy poly(ethylene oxide) = $\text{PI}_x\text{PS}_y\text{AmPEO}_z$).

(I) Living anionic polymerization



(II) Convergent termination



Scheme 2. Synthesis routes towards $\text{PI}_x\text{PS}_y\text{PEO}_z$ triblock copolymers. **(I)** Illustration of the living anionic polymerization forming the PI_xPS_y^- anion with Li^+ as counterion, whereas each block is obtained from the respective monomers. **(II)** Final linkage of the PEO_z block to the PI_xPS_y^- anion using EmPEO_z . The formation of the O-Li-ion pair inhibits further ring-opening polymerization and is highlighted by a dotted circle. For clarity, the *sec*-butyl group resulting from *sec*-butyllithium is drawn only once at the beginning.

Our convergent synthetic method based on the modular principle aims to create access to $\text{PI}_x\text{PS}_y\text{PEO}_z$ with constantly the same very short and well-defined PEO_z block, allowing the PI_x and PS_y blocks to be varied systematically and independently. Moreover, as this synthesis route only utilizes commercially available chemicals, it offers a high reproducibility and up-scaling probability, making tailored and precisely defined $\text{PI}_x\text{PS}_y\text{AmPEO}_z$ BCPs accessible for large-scale production. We have already published a prior patent application for this synthesis method [43].

2. Materials and Methods

2.1. Materials

Sodium *tert*-butoxide (NaO^tBu , 99.9%, Sigma-Aldrich, Darmstadt, Germany) was purified by sublimation [44] (105°C at $\leq 3 \times 10^{-3}$ mbar), 3 Å molecular sieves (VWR, Darmstadt, Germany) was activated by drying at 300°C under vacuum $< 1 \times 10^{-6}$ mbar and methoxy poly(ethylene oxide) ($\text{mPEO}_{1.9}$ equals to $M_n = 1.9 \text{ kg mol}^{-1}$, VWR) was dried at 30°C under vacuum $< 1 \times 10^{-6}$ mbar and all were subsequently stored inside a glovebox (MBraun Unilab, Garching, Germany, ≤ 0.1 ppm of Water (H_2O) and oxygen (O_2)) under argon atmosphere. Epichlorohydrin ($\geq 99.0\%$, Sigma-Aldrich), chloroform- d_1 (CDCl_3 , 99.8% D, VWR) and toluene ($\geq 99.85\%$, VWR) were dried using activated 3 Å molecular sieves until ≤ 1 ppm H_2O , tetrahydrofuran (THF, $\geq 99.8\%$, unstab., Alfa Aesar, Kandel, Germany) was dried using activated 3 Å molecular sieves until ≤ 5 ppm H_2O ; all were subsequently stored inside a glovebox under argon atmosphere and passed through syringe filter (polytetrafluoroethylene (PTFE) membrane, pore size = $0.2 \mu\text{m}$, VWR) prior to use. Seven days before use, isoprene ($\geq 99\%$, VWR) and styrene ($\geq 99\%$, Sigma-Aldrich) were dried (still stored in glovebox fridge) using activated 3 Å molecular sieves until ≤ 1 ppm H_2O and were distilled under vacuum directly before use. A total of 1.4 M *sec*-butyllithium

solution in cyclohexane (*sec*-BuLi, Sigma-Aldrich) was used as received. The concentration of *sec*-BuLi was directly determined by double titration, using a glass-coated magnetic stir bar and the ready-to-use reagent: 2-propanol solution in toluene with 0.2% 1,10 phenanthroline indicator titration solution for quantitative analysis of butyllithium (Sigma-Aldrich), prior to use. Diethyl ether (Et₂O, ≥99.9%, inhibitor-free, Sigma-Aldrich), dichloromethane (DCM, ≥99.5%, VWR), methanol (MeOH, ≥99.9%, VWR) and 0.5 M hydrogen chloride solution in MeOH (Sigma-Aldrich) were used as received. For quantitative water content determination, a Karl Fischer coulometric titrator C30S (Mettler-Toledo, Gießen, Germany) with a platinum generator electrode without a diaphragm was used.

2.2. End Group Modification (EmPEO_{1.9})

The complete end group modification reaction of mPEO_{1.9} to EmPEO_{1.9} was carried out at room temperature under argon atmosphere in a glovebox. NaO^tBu (1.5 equiv, 9.32 mmol, 0.896 g) was dissolved in THF (60 mL) and subsequently added to a solution of mPEO_{1.9} (1.0 equiv, 6.29 mmol, 12.0 g) in THF (60 mL). After 72 h, epichlorohydrin (8.0 equiv, 50.3 mmol, 4.65 g, 3.94 mL) was added dropwise to the reaction within 15 min and stirred for six days. Subsequently, the volatile components were removed at 55 °C under vacuum and the solid residue was dissolved in THF. Undissolved components, mainly formed sodium chloride (NaCl), were removed by centrifugation (Sigma 3-18KS, Osterode am Harz, Germany, 10,000 rpm for 10 min), followed by filtration with syringe filter (PTFE membrane, pore size = 0.2 µm, VWR) and drying at room temperature under vacuum. Afterwards, the product was dissolved in a little amount of toluene at 40 °C, precipitated into cold Et₂O and collected by centrifugation as before. This process was repeated three times. The resulting EmPEO_{1.9} was dried at 30 °C under vacuum <1 × 10^{−6} mbar (yield: 89–93%). The product was characterized by ¹H-nuclear magnetic resonance (NMR) spectroscopy, ¹³C-NMR spectroscopy, gel permeation chromatography (GPC), differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA).

2.3. Synthesis of Poly(isoprene)-*b*-poly(styrene)-*b*-alcohol Methoxy Poly(ethylene oxide) (PI_xPS_yAmPEO_{1.9})

The complete polymer synthesis reaction of PI_xPS_yAmPEO_{1.9} was carried out at room temperature under argon atmosphere in a glovebox using a glass-coated magnetic stir bar for mixing. The PI_x and PS_y block was synthesized by living sequential anionic polymerization as reported in the literature (Scheme 2I) [17–19]. All PI_xPS_yAmPEO_{1.9} materials were synthesized, as exemplarily described for the PI_{14.6}PS_{34.8}AmPEO_{1.9} in the following, where only the *sec*-BuLi, monomers (isoprene and styrene) and EmPEO_{1.9} amounts were adjusted according to the desired composition (cf. Table 1).

Table 1. Overview of *sec*-BuLi (initiator), monomers (isoprene and styrene) and EmPEO_{1.9} used for the synthesis of PI_xPS_yAmPEO_{1.9}.

Polymer	<i>n</i> (<i>sec</i> -BuLi) /mmol	<i>n</i> (Isoprene) /mmol	<i>n</i> (Styrene) /mmol	<i>n</i> (EmPEO _{1.9}) /mmol
PI _{6.8} PS _{17.3} AmPEO _{1.9}	0.256	25.6	42.5	0.282
PI _{14.6} PS _{34.8} AmPEO _{1.9}	0.249	53.1	83.1	0.274
PI _{24.8} PS _{25.0} AmPEO _{1.9}	0.132	48.0	31.5	0.145
PI _{35.1} PS _{14.8} AmPEO _{1.9}	0.233	119.9	33.2	0.257
PI _{26.1} PS _{67.3} AmPEO _{1.9}	0.148	56.6	95.3	0.163

For preparation of the PI_{14.6} block, isoprene (214 equiv, 53.1 mmol, 3.62 g) was dissolved in toluene (120 mL) followed by the addition of 1.4 M *sec*-BuLi solution in cyclohexane (1.0 equiv, 0.249 mmol, 178 µL). After stirring for 24 h, to ensure a quantitative conversion of isoprene monomers [45], styrene (334 equiv, 83.1 mmol, 8.66 g) was added to the yellowish reaction solution of the living PI_{14.6}[−] anion to build the PS_{34.8} block. After

stirring for 24 h, to ensure a quantitative conversion of styrene monomers [45], EmPEO_{1.9} (1.1 equiv, 0.274 mmol, 0.521 g) was added to the reddish reaction solution of the living PI_{14.6}PS_{34.8}[−] anion to attach the PEO_{1.9} block. After stirring for 48 h, 0.5 M hydrogen chloride solution in MeOH (1.5 equiv, 0.37 mmol, 497 μ L) was added to the colorless reaction solution. Subsequently, the volatile components were removed at 40 °C under vacuum. Then, the product was dissolved in DCM, precipitated into MeOH and this process was repeated three times. The resulting polymer was dried at 50 °C under vacuum $<1 \times 10^{-6}$ mbar (yield: 92–95%). The product was obtained as a white solid and characterized by ¹H-NMR spectroscopy, GPC, DSC and TGA.

2.4. Sample Preparation of PI_xPS_yAmPEO_{1.9} Membranes for Morphological Characterization

The complete PI_xPS_yAmPEO_{1.9} membrane casting process, for preparing the samples for the morphological characterization, was performed under argon atmosphere. In a glovebox, an 8 wt.% solution of the dried PI_xPS_yAmPEO_{1.9} in THF was prepared and followed by transferring the mixture into a PTFE crucible. The filled PTFE crucible was placed in a Schlenk vessel and connected to the argon of a Schlenk line. A very low and constant argon flow over 6 days at room temperature in a THF-saturated atmosphere was used to allow a controlled evaporation to achieve distinct microphase separation.

Subsequently, for small-angle X-ray scattering (SAXS) measurement, the resulting membrane (cf. Figure S17) was carefully broken into smaller pieces in order to fit the sample into the glass capillary (borosilicate glass, outer diameter = 2.1 mm, wall thickness = 0.05 mm, Hilgenberg, Malsfeld, Germany). The filled glass capillary was sealed tightly.

For scanning electron microscopy (SEM) and transmission electron microscopy (TEM) investigations, a small piece of the resulting membrane (cf. Figure S17) was cut into ultrathin sections of about 50–100 nm using a cryo-ultramicrotome (cf. Scheme S2).

2.5. Nuclear Magnetic Resonance (NMR) Spectroscopy

The NMR spectra of PI_{14.6}PS_{34.8}AmPEO_{1.9}, PI_{35.1}PS_{14.8}AmPEO_{1.9} and EmPEO_{1.9} were recorded using an AVANCE NEO 500 MHz (Bruker, Billerica, MA, USA), that of mPEO_{1.9}, PI_{24.8}PS_{25.0}AmPEO_{1.9} and PI_{26.1}PS_{67.3}AmPEO_{1.9} using an AVANCE NEO 400 MHz (Bruker) and for PI_{6.8}PS_{17.3}AmPEO_{1.9} using an AVANCE III 400 MHz (Bruker). NMR measurements with number of scans = 64 were recorded, and the recycle delay *D*₁ between transients was set to 30 s to ensure full relaxation to equilibrium magnetization and thus the acquisition of quantitative spectra, except for mPEO_{1.9}. All NMR spectra were recorded at 300 K using CDCl₃ as deuterated solvent, where all signals were referenced to CDCl₃ (δ = 7.3 ppm for ¹H and δ = 77.2 ppm for ¹³C relative to tetramethylsilane) [46]. The spectra were analyzed with the software MestReNova (version: 12.0.4-22023, Mestrelab Research, Santiago de Compostela, Spain).

2.6. Gel Permeation Chromatography (GPC)

GPC was carried out in THF using an HLC-8320GPC EcoSEC (Tosoh Bioscience, Griesheim, Germany) system equipped with three PSS SDV columns (100, 1000, 100000 Å) (8 × 300 mm) of 5 μ m, a UV and a differential refractive index (RI) detector. The operation temperature was set to 35 °C with a flow rate of 1 mL min^{−1}. Calibration of the system was carried out with poly(styrene) standards ranging from 800 to 2.2×10^6 g mol^{−1}. Typically, 50 μ L of a 2.0 mg mL^{−1} sample solution was injected onto the columns.

2.7. Differential Scanning Calorimetry (DSC)

DSC was conducted using a heat flux calorimeter DSC-Q2000 (TA Instruments, New Castle, DE, USA) with LNCS (Liquid Nitrogen Cooling System) and the Tzero®-technology for the precise recording of the baseline. Under argon atmosphere (inside glovebox), ~10 mg of sample was enclosed in hermetically sealed Tzero® aluminum pans. Two heating ramps in the temperature range from −140 °C to 190 °C with a heating rate of 10 K

min^{-1} under helium as sample purge (25 mL min^{-1}) were measured for all samples. The DSC signals were analyzed with the Universal Analysis 2000 software (version: 4.5A, Build 4.5.0.5., TA Instruments).

2.8. Thermogravimetric Analysis (TGA)

For TGA, an aluminum pan was filled with $\sim 10 \text{ mg}$ of sample and hermitically sealed under argon atmosphere in a glovebox, subsequently loaded to the device without contact to ambient air and pierced in the furnace under helium atmosphere. The measurement was carried out on a TGA-5500 with IR furnace (TA Instruments) under helium flow (25 mL min^{-1}) with a constant heating rate of 2 K min^{-1} from 30°C to 600°C . The TGA signals were analyzed with the TA Instruments TRIOS software (version: 5.1.1.46572, TA Instruments).

2.9. Small-Angle X-ray Scattering (SAXS)

The instruments “Ganesha-Air” from SAXSLAB/XENOCs (Grenoble, France) and Gallium Anode Low-Angle X-ray Instrument (GALAXI) were used. The X-ray source of the laboratory-based “Ganesha-Air” system is a D2-MetalJet (Excillum, Kista, Sweden) with a liquid metal anode operating at 70 kV and 3.57 mA with $\text{Ga-K}\alpha$ radiation (wavelength $\lambda = 0.134 \text{ nm}$). The beam is further focused with a focal length of 55 cm , using especially made X-ray optics (Xenocs) to provide a very narrow and intense beam at the sample position. Two pairs of scatterless slits are used to adjust the beam size depending on the detector distance. The data were acquired with a position-sensitive detector (PILATUS 300 K, Dectris, Baden-Daettwil, Switzerland). After calibration with silver behenate, the distance from the sample to the detector was set to 950 and 350 mm resulting in a Q -range $0.13\text{--}6.00 \text{ nm}^{-1}$. All samples were sealed in glass capillaries of 2 mm inner diameter. Data reduction and background subtraction were performed using the Python-based project Jscatter [47]. Fitting of radially averaged SAXS curves was done using Scatter [48].

2.10. Cryo-Ultramicrotomy

To prepare ultra-thin sections for electron microscopy, a Leica EM UC7 ultramicrotome (Wetzlar, Germany) equipped with an EM FC7 cryo-chamber was used. Temperature of sample, knives and chamber was set to -80°C . The samples were trimmed with a diamond trimming knife from Diatome (trim 45, Nidau, Switzerland), and ultra-thin sections were made with a cryo-immuno diamond knife also from Diatome (cf. Scheme S2). Ultra-thin sections were collected dry on carbon-coated copper grids and section thickness was set to 50 nm .

2.11. Scanning Electron Microscope (SEM)

SEM measurements were performed using a Thermo-Fisher Volumescope (Waltham, MA, USA). Images were taken on unstained samples in high vacuum at an accelerating voltage of 30 kV and a working distance of 10 mm using an annular ring STEM detector in bright field mode at room temperature.

2.12. Transmission Electron Microscope (TEM)

For TEM measurements, a JEOL JEM-F200 (Freising, Germany) with field emission gun (FEG) operating at an accelerating voltage of 200 kV was used. Images were taken of the unstained sample with a STEM bright field detector at room temperature.

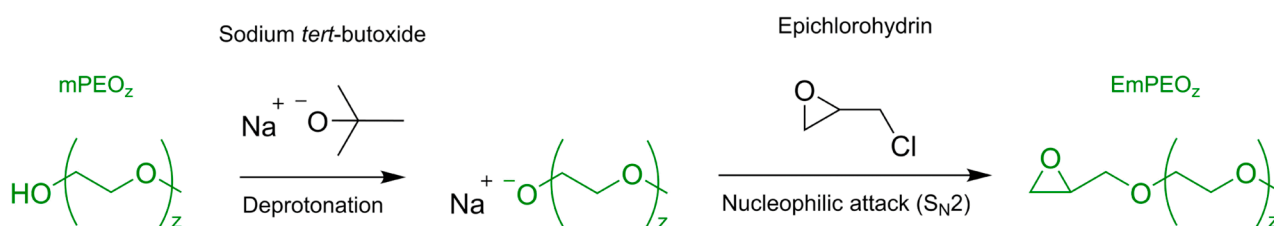
3. Results and Discussion

3.1. Strategy

In line with the introductory part, the main focus of this work is to develop a convergent synthesis route which enables the independent and systematic PI_x and PS_y block variation in BCPs with constantly the same PEO_z block size, whereby the complete control over morphology and orientation of the BCP, especially of the very short PEO_z block, should be obtained, as this is crucial for its possible application. For this purpose, PEO_z blocks with an identical chain length were chosen and introduced into $PI_xPS_yPEO_z$, in which the PI_x and PS_y blocks are synthesized by anionic polymerization of the monomers and therefore easily and precisely modified.

To do so, the commercially available and prefabricated $mPEO_z$ ($M_n = 1.9 \text{ kg mol}^{-1}$) was activated for the attachment by functionalization with an epoxide end group (yielding $EmPEO_z$). In our synthesis route, the $EmPEO_z$ can be attached and thus covalently linked to the previously anionically synthesized $PI_xPS_y^-$ carbanion by a simple one-step addition (cf. Scheme 2). Herein, it is important that an appropriate end group is selected, because it will have a great influence on the properties of the whole PEO_z block of the BCP, especially in our focus of very short PEO_z chains [49]. Thus, the use of an epoxide as an end group ensures that the formed junction point in $PI_xPS_yAmPEO_z$ hardly differs from the $PI_xPS_yPEO_z$ obtained by using EO gas, i.e., only by an additional OH-group. This smallest possible difference enables the selective consideration solely of the variation of the two nonpolar blocks between the different synthesized $PI_xPS_yAmPEO_z$. There are approaches that have been reported in literature thus far in which, for instance, benzophenone or diphenylethylene were attached to the PEO_z chain as terminal groups [49]. However, regarding the reported end group modifications, multi-step synthesis has to be employed, making the overall synthesis of a $PI_xPS_yPEO_z$ more complex, and these groups are more chemically different compared to the ether and OH-group with unintended influence on the polarity [50]. Especially in the case of short PEO_z chains, this could have a great impact on its domain structure formation.

Considering the end group modification shown in Scheme 3, the PEO_z block was end-capped with an epoxide group in a simple literature-known and modified one-pot reaction [51,52]. In the first modification step, the terminal OH-group of $mPEO_z$ is selectively deprotonated by a strong base, i.e., NaO^tBu . Afterwards, the formed $mPEO_z$ alkoxide attacks the epoxide group of the added epichlorohydrin via a nucleophilic (S_N2) attack to form the desired product $EmPEO_z$ (Scheme 3). This method can also be extended for longer ($M_n = 6 \text{ kg mol}^{-1}$) PEO_z chains as reported by van Butsele et al. [52]. Here it should be mentioned, that the PEO_z block has to be selectively functionalized at only one terminal OH-group. End-capping both terminal OH-groups would lead to the undesired formation of a symmetric five BCP ($PI_xPS_yPEO_zPS_yPI_x$). As $mPEO_z$ only possess one terminal OH-group allowing for a distinct functionalization, the linkage to the $PI_xPS_y^-$ anion happens in a selective manner, as indicated in (Scheme 3).



Scheme 3. End group modification of $mPEO_z$ to selectively attach a terminal epoxide group to yield $EmPEO_z$.

Another important aspect of the above-described usage of the prefabricated $mPEO_z$ block is the accurate characterization prior to its linkage, because a PEO_z block that is

already attached to the PI_xPS_y block is relatively small compared to the total BCP and therefore difficult to characterize and control accurately, especially in terms of its D .

The straightforward addition of EmPEO_z to the reaction solution furthermore simplifies the polymerization by terminating the living PI_xPS_y^- anion because it allows a one-step attachment in the presence of the Li^+ counterion in nonpolar solvent (Scheme 2II). After the addition of the EmPEO_z to the reaction solution, the PI_xPS_y^- anion attacks the epoxy group of the EmPEO_z via a ring-opening reaction forming an alcoholate group and a covalent bond between the PI_xPS_y block and the AmPEO_z . Hence, the charge of the carbanion is transferred to the O-atom of the alcoholate group. The negative charge on the O-atom of the alcoholate group is directly blocked by the Li^+ counterion, which strongly reduces the nucleophilicity. Therefore, a new ring-opening reaction with another epoxide group of a second EmPEO_z molecule is inhibited due to the strong O-Li-ion pair (Scheme S1), thus blocking a further polymerization, or any other undesired crosslinking or side reaction and highlighting the advantage of this one-pot synthesis route (Scheme 2) [17,31,37–40]. Especially, a reaction of the living anions with O_2 is suppressed, so that immediately after the AmPEO_z block attachment, the product is stable in air, providing an advantage in terms of reproducibility and up-scalability [53–55].

Moreover, the epoxide functionalization allows the EmPEO_z to be added to the reaction solution in minimal excess (cf. Table 1), so that the PI_xPS_y^- anions react stoichiometrically to form the $\text{PI}_x\text{PS}_y\text{AmPEO}_z$ and the excess and unbound $\text{EmPEO}_{1.9}$ is easily washed out after the polymerization reaction.

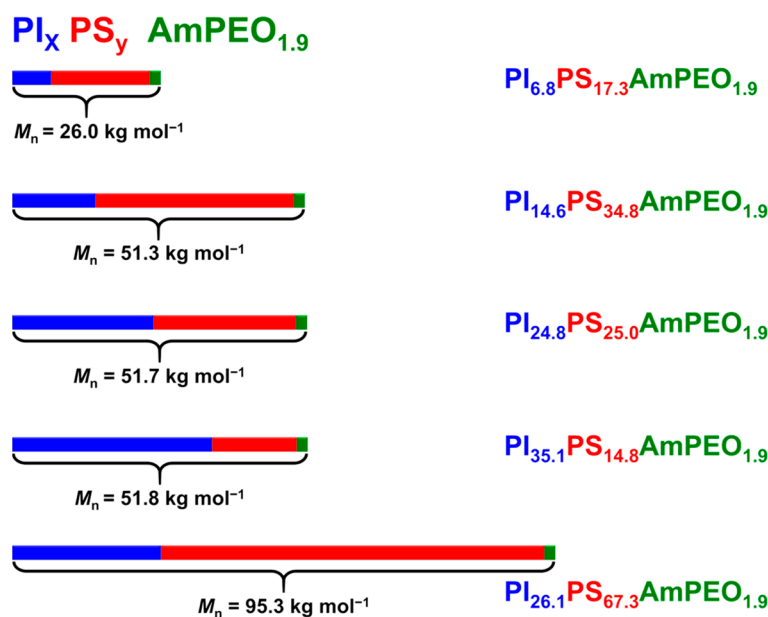
3.2. Synthesis

Triblock copolymers were synthesized in one-pot polymerization in three sequential steps by living anionic polymerization. As indicated by Scheme 2Ia, the PI_x block was prepared from isoprene monomers using *sec*-BuLi serving as initiator for formation of the living carbanion. Subsequently, the living PI_x^- anion was used for chain extension by addition of styrene (Scheme 2Ib). In the third step of this synthesis route, the PEO_z block was covalently attached to the living PI_xPS_y^- anion via a one-step reaction of the epoxide group-functionalized EmPEO_z and thus terminated the polymerization to the BCP (Scheme 2II). The coupling of EmPEO_z to a PI_xPS_y^- anion enabled the selective and highly controllable preparation of $\text{PI}_x\text{PS}_y\text{AmPEO}_z$ in which the $\text{PEO}_{1.9}$ block features an identical chain length and a D very close one in all cases. The respective $M_{n,\text{PI}_x}/M_{n,\text{PS}_y}$ and thus the $M_{n,\text{total}}$ of the BCP can be specifically tailored by systematically varying the used amount of initiator and the PI_x and PS_y monomers. However, as the living anionic polymerization is highly sensitive to impurities, because they directly lead to the termination of the corresponding living anion, special requirements for the synthesis conditions must always be ensured (cf. experimental section).

Here, it is worth noticing that the reaction of mPEO_z to EmPEO_z has to be quantitative, as no separation of both can be conducted due to the strong chemical similarity of mPEO_z and EmPEO_z . Even more, since acidic protons of residual mPEO_z would protonate the PI_xPS_y^- anions upon addition and therefore inhibit the reaction of EmPEO_z and the PI_xPS_y^- anions, the importance of a quantitative functionalization reaction is emphasized. In order to obtain a quantitative end group modification of mPEO_z , the base used for deprotonation of the OH-group has to be a weak nucleophile, as otherwise it would compete with the formed mPEO_z alcoholate regarding the $\text{S}_\text{N}2$ reaction with epichlorohydrin, thus minimizing the yield of the desired EmPEO_z . Additionally, the following conditions have been optimized for the quantitative formation of EmPEO_z : (1) The use of NaO^tBu and epichlorohydrin in excess (relative to the mPEO_z). (2) The reaction has to be carried out under argon atmosphere using anhydrous reactants to ensure that the precipitation of the formed NaCl in THF is quantitatively [56]. Furthermore, the absence of H_2O prohibits the formation of nucleophilic OH $^-$ -ions, which would also compete with the formed mPEO_z alcoholate regarding the reaction with epichlorohydrin.

By the use of EmPEO_{1.9}, different BCPs with constant PEO_{1.9} block size were synthesized, using this universally applicable and simplified synthesis method, by systematically varying the M_n of their PI_x and PS_y blocks (cf. Scheme 1), namely PI_{6.8}PS_{17.3}AmPEO_{1.9}, PI_{14.6}PS_{34.8}AmPEO_{1.9}, PI_{24.8}PS_{25.0}AmPEO_{1.9}, PI_{35.1}PS_{14.8}AmPEO_{1.9} and PI_{26.1}PS_{67.3}AmPEO_{1.9} (cf. Scheme 4). While the $M_{n,total}$ of the BCP and thus the PEO_{1.9} block fraction remained constant, the $M_{n,PI_x}/M_{n,PS_y}$ was varied (cf. Scheme 1a). According to this, for PI_{24.8}PS_{25.0}AmPEO_{1.9} the M_{n,PI_x} is equal to the M_{n,PS_y} , in the case of PI_{35.1}PS_{14.8}AmPEO_{1.9} the M_{n,PI_x} is doubled compared to the M_{n,PS_y} , whereas for PI_{14.6}PS_{34.8}AmPEO_{1.9} the M_{n,PS_y} is two times larger than the M_{n,PI_x} , which is exactly the opposite compared to PI_{35.1}PS_{14.8}AmPEO_{1.9}.

Block weight ratio



Scheme 4. Comparison of the block weight ratio of the prepared PI_xPS_yAmPEO_{1.9}, which are systematically and independently varied with constantly always the same PEO_{1.9} block size. The polymer chain length of each block is shown in relation to the theoretical $M_{n,calc}$ calculated from the ratio of the used masses of monomers to initiator (cf. Table 1).

In addition, by keeping a $M_{n,PI_x}/M_{n,PS_y}$ constant, in which the M_{n,PS_y} is doubled compared to the M_{n,PI_x} , the PEO_{1.9} block proportion was varied by changing the $M_{n,total}$ of the BCP (cf. Scheme 1b). Therefore, by halving the $M_{n,total}$ of PI_{14.6}PS_{34.8}AmPEO_{1.9}, the PEO_{1.9} block fraction was doubled in PI_{6.8}PS_{17.3}AmPEO_{1.9}. In contrast, for PI_{26.1}PS_{67.3}AmPEO_{1.9} the $M_{n,total}$ was varied exactly in the opposite way, i.e., the $M_{n,total}$ of PI_{14.6}PS_{34.8}AmPEO_{1.9} was doubled, resulting in a reduced PEO_{1.9} block proportion.

3.3. NMR Characterization

3.3.1. EmPEO_{1.9}

First, the EmPEO_{1.9} obtained from the functionalization reaction of mPEO_{1.9} (Scheme 3) was analyzed for its quality via ¹H- and ¹³C-NMR to verify that the end group modification was quantitative as required. The ¹H-NMR spectrum of EmPEO_{1.9} (Figure 1) shows the proton and satellite peaks (highlighted by the grey box) from the polyether chain in the chemical shift region of $\delta = 3.4$ – 3.7 ppm (c in orange circle) [51,52]. The sharp singlet at $\delta = 3.3$ ppm (d in grey circle) can be assigned to the three protons of the terminal methoxy group [51,52]. The three protons of the epoxide group attached by functionalization reaction split into three characteristic signals located at $\delta = 2.5$ ppm (a in blue circle), $\delta = 2.7$ ppm (a' in blue circle) and $\delta = 3.1$ ppm (b in yellow circle) [51,52]. The integrals of these three signals are equal and have a value of one with respect to the three protons of the

terminal methoxy group, indicating that the attachment of the epoxide group took place quantitatively.

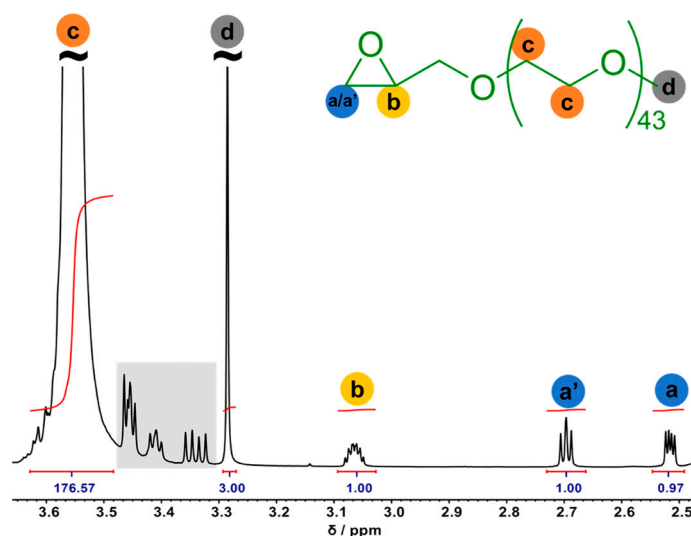


Figure 1. ^1H -NMR spectrum of the reaction product EmPEO_{1.9} obtained by the functionalization reaction of mPEO_{1.9} with epichlorohydrin (Scheme 3). The integrated signals indicate a quantitative formation of EmPEO_{1.9}. The protons of the bridging CH_2 -group (between epoxide group and the polyether chain) as well as the satellite signals of the polyether chain are highlighted by the grey box.

The comparison of the ^{13}C -NMR spectra of mPEO_{1.9} and EmPEO_{1.9} (Figure 2) highlights the characteristic carbon signal changes due to the modification reaction (cf. for entire spectrum of mPEO_{1.9} Figure S4 and EmPEO_{1.9} Figure S2). For mPEO_{1.9} (Figure 2a), the carbon atom signal at $\delta = 61.7$ ppm (a in dark blue circle) corresponds to the carbon atom directly linked to the OH-group and the signal at $\delta = 72.5$ ppm to its directly adjacent carbon atom (b in yellow circle). Both carbon atom signals are not detected in the ^{13}C -NMR spectrum of the EmPEO_{1.9} (Figure 2b), confirming that the functionalization of all OH-groups with epoxide groups to form EmPEO_{1.9} is quantitatively [51]. The carbon atoms of the polyether chain at $\delta = 70.5$ ppm (c in orange circle) and the terminal methoxy group at $\delta = 59.0$ ppm (e in light green circle) are not affected during the modification reaction, thus the respective signals can be found in both ^{13}C -NMR spectra [51]. In contrast to the ^{13}C -NMR spectrum of mPEO_{1.9} (Figure 2a), two carbon atom signals appear at $\delta = 44.1$ ppm (f in light blue circle) and $\delta = 50.7$ ppm (g in brown circle) in ^{13}C -NMR spectrum of EmPEO_{1.9} (Figure 2b) which are characteristic for the attached epoxide group [51].

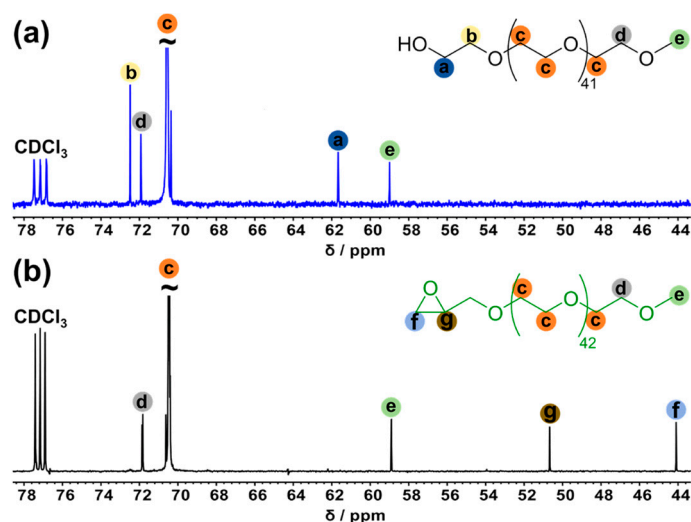


Figure 2. (a) ^{13}C -NMR spectrum of mPEO_{1.9} and (b) of EmPEO_{1.9}. The absence of signals a and b in ^{13}C -NMR spectrum of EmPEO_{1.9} indicates the quantitative formation. For better clarity, the most intense signal (marked by c in filled orange circle) was cut off, symbolized by inserted tilde.

The successful and quantitative attachment of the epoxide group to the mPEO_{1.9} molecule was verified by the integrals in ^1H -NMR as well as the altered signals in ^{13}C -NMR. Furthermore, the absence of additional signals in the ^1H - and ^{13}C -NMR spectra of EmPEO_{1.9} (Figures 1 and 2b) indicated that during the modification reaction, no side reaction occurred, and no impurities were introduced (cf. Figures S1 and S2 for entire spectrum).

3.3.2. PI_xPS_yAmPEO_{1.9}

The quantitatively epoxy-functionalized EmPEO_{1.9} was used for the synthesis of different PI_xPS_yAmPEO_{1.9}, which were subsequently characterized by ^1H -NMR measurements. In the following, PI_{14.6}PS_{34.8}AmPEO_{1.9} is discussed exemplarily based on its ^1H -NMR result, whereas in Figure S10 and Table S1, the results of all synthesized PI_xPS_yAmPEO_{1.9} are summarized, while the individual spectra are depicted in Figure S5–S9.

The ^1H -NMR spectrum of PI_{14.6}PS_{34.8}AmPEO_{1.9} is displayed in Figure 3. The signals in the range of $\delta = 7.4$ – 6.2 ppm can be assigned to the five aromatic protons of the phenyl group of PS_{34.8} block (red circles). The signal of the olefinic proton from the 1,4-PI_{14.6} block (blue circle) is located at $\delta = 5.2$ ppm and the two terminal protons of the 3,4-PI_{14.6} block (blue circle) are located at $\delta = 4.8$ and 4.7 ppm, respectively [57]. The integral ratio of the 1,4-PI_{14.6} block to the 3,4-PI_{14.6} block is 1.00: 0.15, being the same for all prepared PI_xPS_yAmPEO_{1.9} [57]. The protons of the polyether chain of the AmPEO_{1.9} (dark green circle) are localized at $\delta = 3.7$ ppm. The signal at $\delta = 3.4$ ppm can be attributed to the terminal methoxy group of AmPEO_{1.9} (light green circle), and the alkyl backbone protons of the entire PI_{14.6}PS_{34.8}AmPEO_{1.9} (light grey circles) are located in the range of $\delta = 2.3$ – 1.3 ppm. The fact that the proton signals from AmPEO_{1.9} are still present after the purification procedure of PI_{14.6}PS_{34.8}AmPEO_{1.9} indicates that AmPEO_{1.9} was covalently linked to the PI_{14.6}PS_{34.8}[−] anion by a nucleophilic attack on the epoxide group of EmPEO_{1.9}, as described previously (cf. Scheme 2II). The nonbonded EmPEO_{1.9} was removed during the purification procedure because it dissolves in MeOH, which is used in the purification process [58].

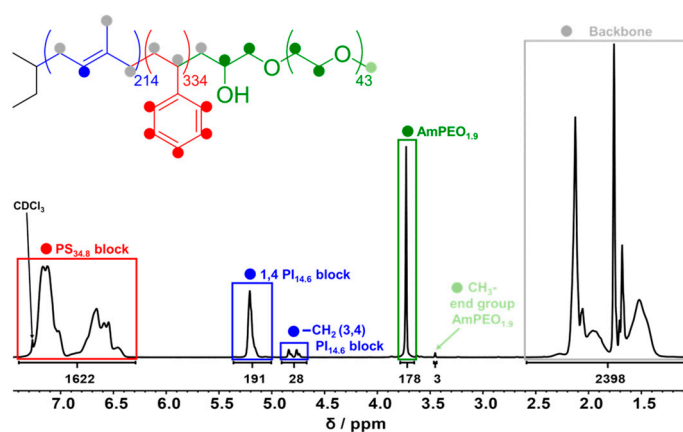


Figure 3. ^1H -NMR spectrum of the PI_{14.6}PS_{34.8}AmPEO_{1.9} obtained by the synthesis route with EmPEO_{1.9} (Scheme 2) after precipitation and drying. The presence of the AmPEO_{1.9} proton signals after purification procedure indicates that the AmPEO_{1.9} was covalently attached to the PI_{14.6}PS_{34.8}[−] anion. The integral of the protons fit to the theoretical molar mass ($M_{n,\text{calc.}}$) of the individual blocks, which implies 100% reaction efficiency. The protons of the butyl group (shown in the structural formula, originating from the *sec*-BuLi) were not displayed in the spectrum for a better overview.

The theoretical molar mass ($M_{n,calc.}$) of the individual blocks shown in Table 2, which was calculated from the ratio of the used masses of monomers (isoprene and styrene) to *sec*-BuLi (initiator) (cf. Table 1), fit to the respective ^1H -NMR integrals of the characteristic protons from the PI_x, PS_y and AmPEO_{1.9} blocks for all synthesized PI_xPS_yAmPEO_{1.9} (Table 3, Figure S10, Table S1).

Table 2. Theoretical constitutional repeating units ($\text{CRU}_{calc.}$) and molar masses ($M_{n,calc.}$) of the individual blocks from the synthesized PI_xPS_yAmPEO_{1.9}, calculated from the ratio of the used masses of monomers to initiator (cf. Table 1).

Polymer	$M_{n,calc.}$ /kg mol ⁻¹	PI _x CRU _{calc.}	$M_{n,calc.,PI_x}$ /kg mol ⁻¹	PS _y CRU _{calc.}	$M_{n,calc.,PS_y}$ /kg mol ⁻¹	PEO _{1.9} CRU _{calc.}	$M_{n,calc.,PEO_{1.9}}$ /kg mol ⁻¹
PI _{6.8} PS _{17.3} AmPEO _{1.9}	26.0	100	6.8	166	17.3	43	1.9
PI _{14.6} PS _{34.8} AmPEO _{1.9}	51.3	214	14.6	334	34.8	43	1.9
PI _{24.8} PS _{25.0} AmPEO _{1.9}	51.7	365	24.8	240	25.0	43	1.9
PI _{35.1} PS _{14.8} AmPEO _{1.9}	51.8	516	35.1	142	14.8	43	1.9
PI _{26.1} PS _{67.3} AmPEO _{1.9}	95.3	384	26.1	646	67.3	43	1.9

Considering the fact that the molar masses ($M_{n,NMR}$) of the individual blocks shown in Table 3 determined from the respective proton number of ^1H -NMR integrals (cf. Table S1) agree with the corresponding $M_{n,calc.}$ indicates a nearly stoichiometric conversion of the utilized masses listed in Table 1 to the desired PI_xPS_yAmPEO_{1.9} without any side reactions. This close to quantitative polymerization process is due to the high reaction control by living anionic polymerization, the controlled synthesis conditions and the use of a slight excess of the EmPEO_{1.9} during the polymer synthesis. The slight excess (10%) of EmPEO_{1.9} ensures that all PI_xPS_y[−] anions are stoichiometrically saturated with the corresponding AmPEO_{1.9} blocks, whereby the chosen epoxide end group modification ensures that not more than one EmPEO_{1.9} molecule can be covalently attached per PI_xPS_y[−] anion (cf. Schemes 2 and S1). Therefore, the excess of unattached EmPEO_{1.9} was removed during the purification process. In addition, the strong O-Li-ion pair acts as a kind of protecting group, avoiding unwanted crosslinking, side and termination reactions.

Table 3. Practical constitutional repeating units (CRU_{NMR}) and molar masses ($M_{n,NMR}$) of the individual polymer blocks from the synthesized PI_xPS_yAmPEO_{1.9} determined from the respective protons number of ^1H -NMR integrals.

Polymer	$M_{n,NMR}$ /kg mol ⁻¹	PI _x CRU _{NMR}	M_{n,NMR,PI_x} /kg mol ⁻¹	PS _y CRU _{NMR}	M_{n,NMR,PS_y} /kg mol ⁻¹	PEO _{1.9} CRU _{NMR}	$M_{n,NMR,PEO_{1.9}}$ /kg mol ⁻¹
PI _{6.8} PS _{17.3} AmPEO _{1.9}	26.8	101	6.9	174	18.1	42	1.9
PI _{14.6} PS _{34.8} AmPEO _{1.9}	49.7	206	14.0	324	33.8	44	1.9
PI _{24.8} PS _{25.0} AmPEO _{1.9}	49.7	346	23.6	231	24.1	45	2.0
PI _{35.1} PS _{14.8} AmPEO _{1.9}	49.6	489	33.3	137	14.3	45	2.0
PI _{26.1} PS _{67.3} AmPEO _{1.9}	106.2	460	31.3	702	73.1	41	1.8

3.4. GPC Measurements

The GPC measurements were conducted to determine the molar mass ($M_{n,GPC}$), to confirm the $M_{n,NMR}$ results, the efficiency of the polymer synthesis and to determine the $\bar{D}$ of the EmPEO_{1.9} as well as the synthesized PI_xPS_yAmPEO_{1.9}. The corresponding GPC traces are shown in Figure 4, and Table 4 lists the results (cf. Figures S11–S15 for the complete traces). From the GPC traces, it can be seen that all synthesized PI_xPS_yAmPEO_{1.9} polymers and EmPEO_{1.9} have a narrow, unimodal shape and as a result, a $\bar{D}$ smaller than 1.10. These are: PI_{6.8}PS_{17.3}AmPEO_{1.9} ($\bar{D}$ = 1.02), PI_{14.6}PS_{34.8}AmPEO_{1.9} ($\bar{D}$ = 1.02), PI_{24.8}PS_{25.0}AmPEO_{1.9} ($\bar{D}$ = 1.03), PI_{35.1}PS_{14.8}AmPEO_{1.9} ($\bar{D}$ = 1.03), PI_{26.1}PS_{67.3}AmPEO_{1.9} ($\bar{D}$ = 1.08) with an additional minor signal, and in comparison, EmPEO_{1.9} ($\bar{D}$ = 1.03). Moreover, it can be clearly observed

by the overlapping of the traces that the $M_{n, \text{GPC}}$ of $\text{PI}_{14.6}\text{PS}_{34.8}\text{AmPEO}_{1.9}$, $\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$ and $\text{PI}_{35.1}\text{PS}_{14.8}\text{AmPEO}_{1.9}$ are almost identical and in the range of $M_{n, \text{GPC}} = 62.4\text{--}65.3 \text{ kg mol}^{-1}$, whereas the $M_{n, \text{GPC}}$ of $\text{PI}_{6.8}\text{PS}_{17.3}\text{AmPEO}_{1.9}$ is shifted to a lower value at $M_{n, \text{GPC}} = 31.2 \text{ kg mol}^{-1}$ and for $\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$ to a higher value at $M_{n, \text{GPC}} = 107.6 \text{ kg mol}^{-1}$, which corresponds to their respective $M_{n, \text{calc.}}$ and $M_{n, \text{NMR}}$ (cf. Figure 4, Table 4). The deviation towards higher molar mass in the $M_{n, \text{GPC}}$ compared to the $M_{n, \text{NMR}}$ values is attributed to the fact that the GPCs used polystyrene calibration standards. Furthermore, the GPC trace of $\text{EmPEO}_{1.9}$ with $M_{n, \text{GPC}} = 2.8 \text{ kg mol}^{-1}$ shows a very narrow chain length distribution of $D = 1.03$. Thus, it was successfully characterized prior to linkage, showing a major advantage over the short PEO_z block control.

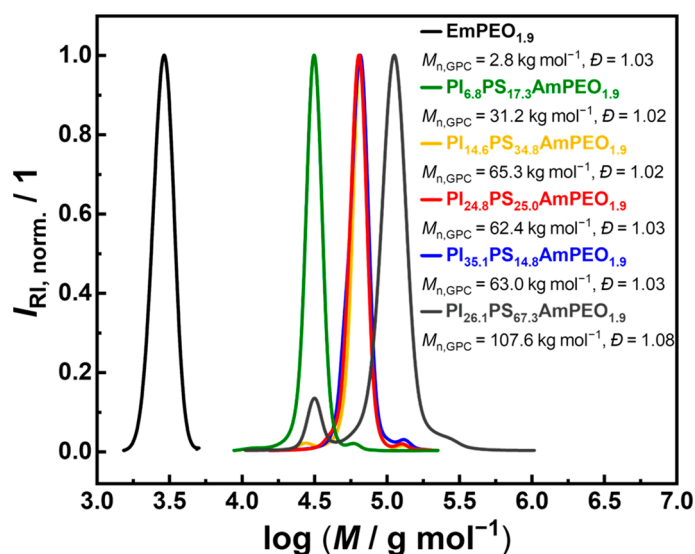


Figure 4. GPC traces of $\text{EmPEO}_{1.9}$ as well as of the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ polymers after the linkage of $\text{EmPEO}_{1.9}$. The same narrow signal shape of all synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ as well as of the $\text{EmPEO}_{1.9}$ indicate their nearly monodisperse polymer chain lengths ($D \leq 1.08$). For $\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$, a very small additional signal was detected.

By linking the PEO_z block via EmPEO_z , it prevents unwanted chain termination caused by the introduction of impurities, which has a considerable influence on D .

The result that the D values are close to one ($D \leq 1.08$) indicates that the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_z$ polymer chains have nearly the same length, which means that there are almost no chain terminations during the polymerization process. This is consistent with the $^1\text{H-NMR}$ results that the polymerization process is nearly quantitative and thus $M_{n, \text{calc.}} = M_{n, \text{NMR}}$. Therefore, this simplified polymerization procedure allows a fast and efficient polymer synthesis of tailored $\text{PI}_x\text{PS}_y\text{AmPEO}_z$ BCPs in a controlled and reproducible way.

Table 4. Overview of the theoretical molar mass ($M_{n, \text{calc.}}$), the molar mass determined via $^1\text{H-NMR}$ measurement ($M_{n, \text{NMR}}$) and the molar mass obtained from GPC measurement ($M_{n, \text{GPC}}$) as well as the dispersity (D) of $\text{EmPEO}_{1.9}$ and the different synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ polymers.

Polymer	$M_{n, \text{calc.}}^1$ /kg mol $^{-1}$	$M_{n, \text{NMR}}^2$ /kg mol $^{-1}$	$M_{n, \text{GPC}}^3$ /kg mol $^{-1}$	D^3
$\text{EmPEO}_{1.9}$	1.9	2.0	2.8	1.03
$\text{PI}_{6.8}\text{PS}_{17.3}\text{AmPEO}_{1.9}$	26.0	26.8	31.2	1.02
$\text{PI}_{14.6}\text{PS}_{34.8}\text{AmPEO}_{1.9}$	51.3	49.7	65.3	1.02
$\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$	51.7	49.7	62.4	1.03
$\text{PI}_{35.1}\text{PS}_{14.8}\text{AmPEO}_{1.9}$	51.8	49.6	63.0	1.03
$\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$	95.3	106.2	107.6	1.08

¹ Calculated from the ratio of the used masses of monomers to initiator (cf. Table 2). ² Determined from the respective proton number of ¹H-NMR integrals (cf. Figure S10, Table S1). ³ Determined by GPC measurement in THF using polystyrene standards (cf. Figure 4).

3.5. Thermal Analysis

The synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ polymers were analyzed by DSC (cf. Figure 5a). Besides the thermal induced phase transitions, these measurements were carried out in order to obtain relevant information regarding the microphase separation.

The occurrence of the three characteristic thermal phase transitions, i.e., the glass transition temperature (ϑ_g) for the PI_x phase $\vartheta_{g,\text{PI}_x}$ in the range of -68 to -60 °C, the $\vartheta_{g,\text{PS}_y}$ for the PS_y phase in the range of 75 to 96 °C, as well as the melting point $\vartheta_{\text{mp,PEO}_{1.9}}$ of the $\text{PEO}_{1.9}$ (from $\text{AmPEO}_{1.9}$) at ~ 50 °C (cf. Figure S16 for DSC measurement of pure $\text{EmPEO}_{1.9}$), evidence a microphase separation of the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ BCPs (Table 5) [59]. The existence of $\vartheta_{\text{mp,PEO}_{1.9}}$ indicates the phase separation of the $\text{PEO}_{1.9}$ block from both non-polar blocks. An exception is $\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$, where the $\vartheta_{\text{mp,PEO}_{1.9}}$ is not detected. The reason for the absence of the ordered crystalline structure of the $\text{PEO}_{1.9}$ chains in $\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$ can be attributed either to their low content of 1.7 wt.% or to the minimal contamination, which was detected in GPC measurement, so that the entire $\text{PEO}_{1.9}$ block is amorphous. Moreover, the two ϑ_g and the $\vartheta_{\text{mp,PEO}_{1.9}}$ are equal to those of the pure polymer blocks, and the absence of additional ϑ_g and ϑ_{mp} suggests the exclusion of any mixing of the individual blocks even at the phase boundaries. In consequence of the high tendency of phase separation (affected by the nearly monodisperse polymer chain lengths ($D \leq 1.08$)), a correlation between the chain length of the individual polymer blocks and the respective ϑ_g is observed. An increase of the block length, which corresponds to a higher M_n of the corresponding polymer block, leads to a shift of the ϑ_g to higher temperature; for instance, an increase in M_{n,PS_y} by ~ 10 kg mol⁻¹ leads to a $\vartheta_{g,\text{PS}_y}$ temperature increase of ~ 5 °C (cf. Table 5). Furthermore, the fact that there is no mixing between different polymer blocks indicates that a linear dependence of change of heat capacity ΔC_p with the weight fraction of the respective polymer block at ϑ_g is detected. For example, an increase in M_{n,PI_x} of ~ 20 wt.% leads to a rise in the $\Delta C_{p,\text{PI}_x}$ by ~ 0.12 J (g K)⁻¹ for the $\vartheta_{g,\text{PI}_x}$ (Table 5).

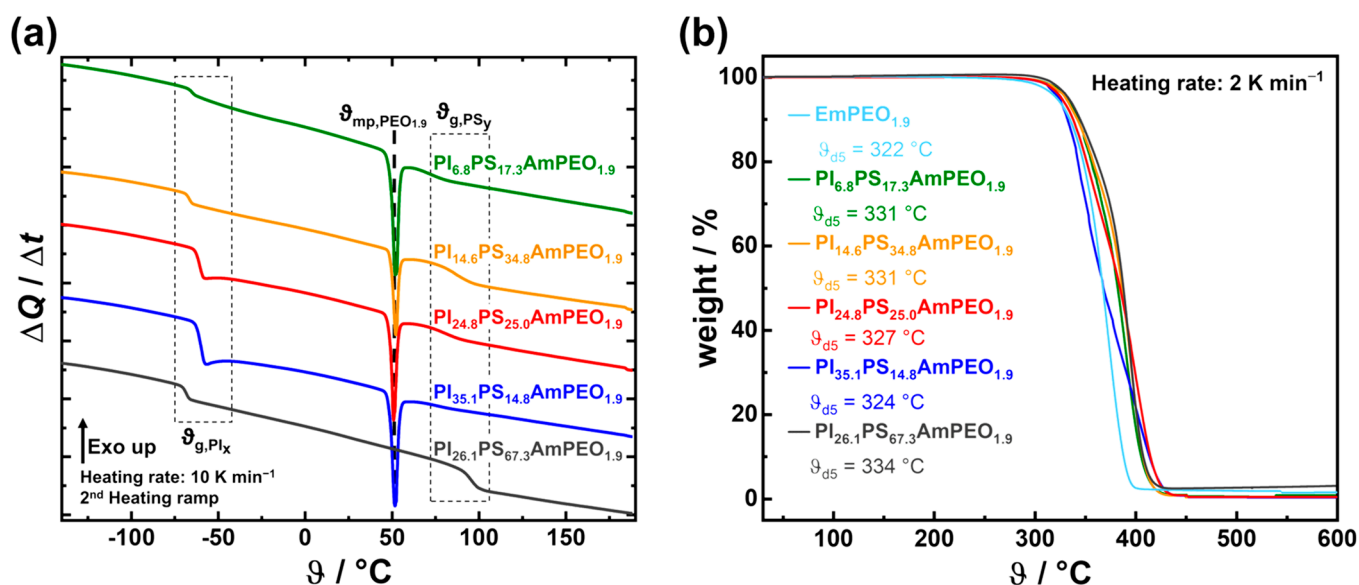


Figure 5. (a) DSC measurements of the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ polymers, which show that they are microphase-separated as proven by the distinct ϑ_g (glass transition temperature) of the PI_x and PS_y phase as well as the ϑ_{mp} (melting point) of the $\text{PEO}_{1.9}$ phase. An exception is $\text{PI}_{26.1}\text{PS}_{67.3}\text{AmPEO}_{1.9}$, where the $\vartheta_{\text{mp,PEO}_{1.9}}$ is not detected. (b) TGA curves of the $\text{EmPEO}_{1.9}$ and the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$ BCPs and their respective decomposition temperature at 5% weight loss (ϑ_{d5}).

In Figure 5b, the decomposition temperature at a weight loss of 5% (Θ_{d5}) of EmPEO_{1.9} and the synthesized PI_xPS_yAmPEO_{1.9} polymers as determined by TGA measurements are shown. The Θ_{d5} values are similar and independent of the polymer composition, with a slight trend of decreasing Θ_{d5} with higher PI_x content and range from 322 to 334 °C. Therefore, they have a high thermal stability for the application as polymer electrolyte templates.

Table 5. Overview of the weight fractions (wt.%), the thermal induced phase transitions Θ_{mp} (melting point), Θ_g (glass transition temperature) and the ΔC_p (heat capacity changes) of the individual polymer blocks well as the Θ_{d5} (decomposition temperature at 5% weight loss) of the different PI_xPS_yAmPEO_{1.9}.

Polymer	PI _x ¹ /wt. %	PS _y ¹ /wt. %	PEO _{1.9} ¹ /wt. %	Θ_{g,PI_x} ² /°C	$\Delta C_{p,PI_x}$ ² /J (g K) ^{−1}	Θ_{g,PS_y} ² /°C	$\Delta C_{p,PS_y}$ ² /J (g K) ^{−1}	$\Theta_{mp,PEO_{1.9}}$ ² /°C	Θ_{d5} ³ /°C
PI _{6.8} PS _{17.3} AmPEO _{1.9}	25.6	67.5	6.9	−63.7	0.13	74.6	0.12	50.2	331
PI _{14.6} PS _{34.8} AmPEO _{1.9}	28.2	68.0	3.9	−65.9	0.14	86.8	0.24	50.6	331
PI _{24.8} PS _{25.0} AmPEO _{1.9}	47.5	48.5	4.0	−60.7	0.27	80.9	0.14	49.9	327
PI _{35.1} PS _{14.8} AmPEO _{1.9}	67.2	28.8	4.0	−60.1	0.37	76.1	0.06	49.9	324
PI _{26.1} PS _{67.3} AmPEO _{1.9}	29.5	68.8	1.7	−67.6	0.14	95.8	0.21	−	334

¹ Calculated using $M_{n,NMR}$ (cf. Table 3). ² Determined by DSC measurement (cf. Figure 5a). ³ Determined by TGA measurement (cf. Figure 5b).

3.6. Morphological Characterization

Based on the previous results, it is evident that this polymerization route provides precise control over the chain length distribution respectively structure on a molecular level. In the following, it is investigated how this in combination with the controlled solvent casting process affects the morphology of the membrane. In particular, the self-assembly induced highly ordered microphase separation, which was already observed from DSC measurement, is crucial for the properties of the PI_xPS_yAmPEO_{1.9} polymers and therefore plays a key role in their application. The morphology of all synthesized PI_xPS_yAmPEO_{1.9} is determined using SAXS, SEM and TEM measurements. The results of above-mentioned morphological characterizations performed on the PI_xPS_yAmPEO_{1.9} membranes are shown in Figures 6–9 and Table 6, which were controlled-cast from THF without annealing them later.

SAXS measurements were performed at room temperature on the as-cast membranes for all synthesized PI_xPS_yAmPEO_{1.9}. Figure 6 and Table 6 show the results. All of the synthesized PI_xPS_yAmPEO_{1.9} specimens exhibited strong scattering, as seen in the 2D SAXS patterns in Figures 7, 8 and S18, and a clear first scattering peak (q^*), which was marked by a yellow filled circle in the SAXS curves in Figure 6, confirming the microphase separation [60,61].

In the SAXS curve of PI_{6.8}PS_{17.3}AmPEO_{1.9}, in addition to the $q^* = 0.32 \text{ nm}^{-1}$, additional peaks were measured at a relative peak position at q/q^* of $\sqrt{1}$, $\sqrt{3}$, $\sqrt{4}$ and $\sqrt{7}$, indicating a HEX structure (cf. Figure 6) [18]. The same characteristic scattering peaks of a microphase-separated HEX morphology are detected for the membranes consisting of PI_{14.6}PS_{34.8}AmPEO_{1.9}, PI_{35.1}PS_{14.8}AmPEO_{1.9} and PI_{26.1}PS_{67.3}AmPEO_{1.9}, as shown in Figure 6 [18]. However, all q values of these polymers are shifted in comparison to PI_{6.8}PS_{17.3}AmPEO_{1.9} to lower q values, whereby PI_{26.1}PS_{67.3}AmPEO_{1.9} has the lowest q^* value with 0.11 nm^{-1} . In the case of PI_{24.8}PS_{25.0}AmPEO_{1.9}, the q^* value = 0.20 nm^{-1} is the same as for PI_{14.6}PS_{34.8}AmPEO_{1.9} (0.20 nm^{-1}) and PI_{35.1}PS_{14.8}AmPEO_{1.9} (0.19 nm^{-1}). In addition to the q^* peak, further peaks at a relative position of q/q^* 2 and 3 are clearly observed for PI_{24.8}PS_{25.0}AmPEO_{1.9}, indicating a LAM morphology (cf. Figure 6) [18,60].

The q^* values were used to calculate the respective average domain spacing (d) for all PI_xPS_yAmPEO_{1.9}; this means for a HEX structure, the distance is from cylinder to adjacent cylinder, and for a LAM structure, from center to the next center, which is listed in Table

6. In addition, from SAXS measurement of $\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$ the layer thickness (t) and for samples with the HEX morphology, the cylinder radius (r_{cylinder}) were determined and summarized in Table 6. Details concerning the fits to the radially averaged SAXS data in Figure 6 and the values in Table 6 are given in the supporting information.

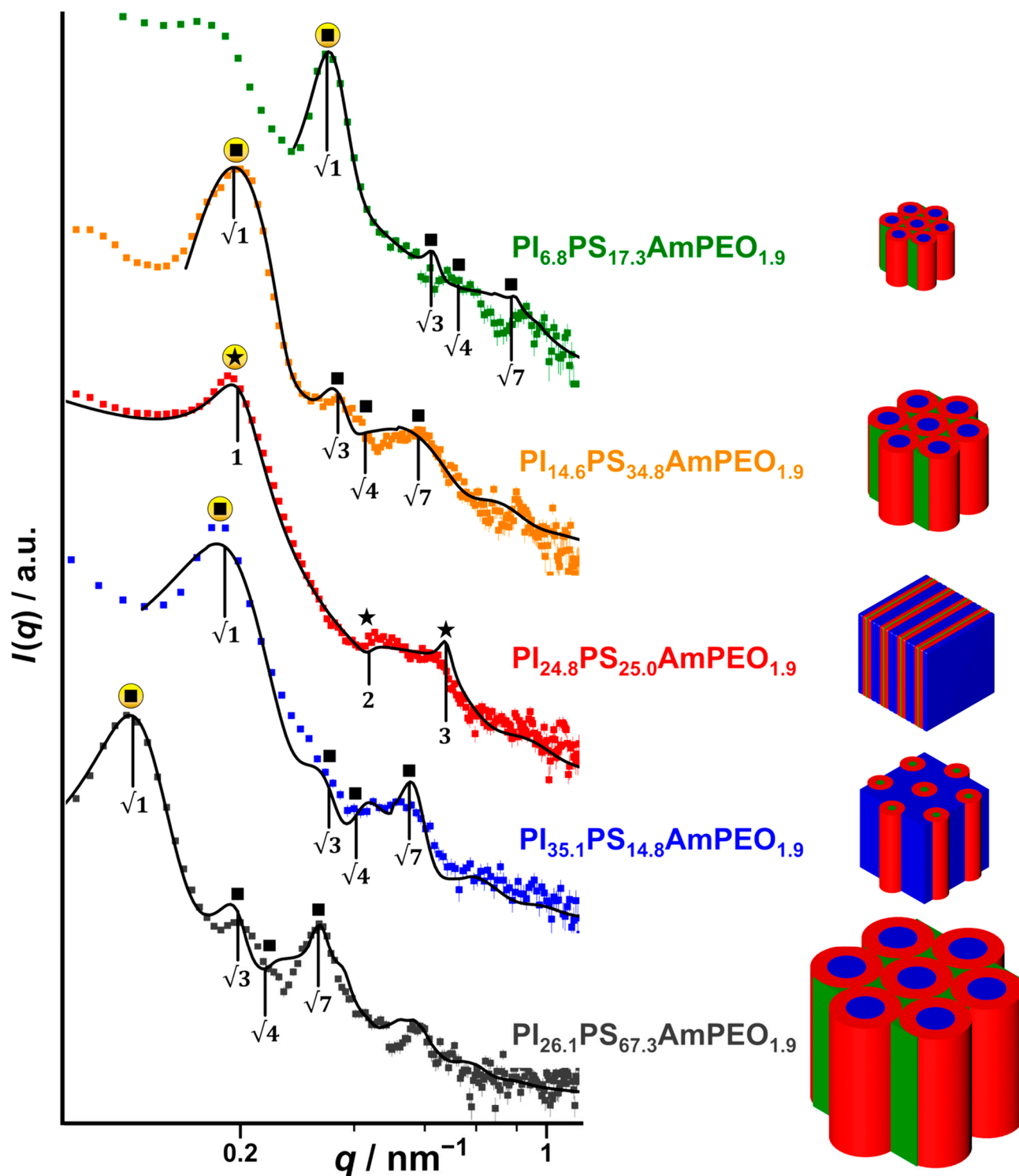


Figure 6. SAXS curves measured at room temperature of as-cast membranes from the synthesized $\text{PI}_x\text{PS}_y\text{AmPEO}_{1.9}$, by plotting the scattering intensity as a function of the magnitude of the scattering vector (q). The first reflection peak (q^*) determined from SAXS measurement is marked by a yellow filled circle. Black squares with the values below indicate the calculated relative peak positions at q/q^* due to hexagonally close-packed cylindrical structure and black stars indicate the respective peaks due to the lamellar structure. The theoretical fit for the morphology (black lines) and the curve

obtained from SAXS measurement, as well as the calculated peak positions, match almost perfectly. On the right side are 3D drawings of the corresponding structure, therein blue represents PI_x, red PS_y and green PEO_{1.9} volume fraction and the size ratios of the drawing are equal to the average domain spacing (d) obtained from SAXS measurement.

The extra hump in the SAXS curve from PI_{6.8}PS_{17.3}AmPEO_{1.9} at low q value = 0.16 nm⁻¹, which therefore precedes the q^* value = 0.32 nm⁻¹ and was not fitted, comes most probably from heterogeneities in the structure on length scales larger than the unit cell. The reason that only the one peak is seen and the others are suppressed lies in the ratio of r_{cylinder} and d because in this combination, the peaks fall on the minima and are suppressed.

For all investigated polymers, the PEO_{1.9} block is with a volume fraction (φ) of 2–6% too small to be detected by SAXS measurements. Thus, it was not necessary to include the PEO_{1.9} block in the fit, as from DSC measurements it is known to be phase-separated from the nonpolar blocks, so a simplified diblock model was used.

For all samples, the fit used to determine the morphology matches very well to the measured SAXS curve shapes (cf. plotted black line in the SAXS curves in Figure 6). Therefore, the calculated relative peak positions at q/q^* (cf. black squares for HEX and black stars for LAM structure in Figure 6) fit almost perfectly to the maxima and minima in the respective obtained SAXS curves. The model used for the fitting is either using homogeneous cylinder form factor for the HEX or platelets form factor for the LAM phases. Even with this rather simple model, a precise fit of the radially averaged scattering data was achieved. The φ for different blocks were calculated from the ratio of r_{cylinder} or t to the fitted d and are listed in Table 6.

The d values calculated from measured q^* correlate with the total $M_{n,NMR}$ of the PI_xPS_yAmPEO_{1.9}. This means that PI_{6.8}PS_{17.3}AmPEO_{1.9} possesses not only the smallest structure size with $d = 22.4$ nm, but also the fact that it is about half of the size compared to PI_{14.6}PS_{34.8}AmPEO_{1.9} with $d = 35.7$ nm, PI_{24.8}PS_{25.0}AmPEO_{1.9} with $d = 31.7$ nm and PI_{35.1}PS_{14.8}AmPEO_{1.9} with $d = 39.2$ nm. The same behavior, but this time reversed, is seen for PI_{26.1}PS_{67.3}AmPEO_{1.9}, which not only has the largest structure size with $d = 63.6$ nm, but also its size is in comparison to PI_{14.6}PS_{34.8}AmPEO_{1.9}, PI_{24.8}PS_{25.0}AmPEO_{1.9} and PI_{35.1}PS_{14.8}AmPEO_{1.9} about two times larger. Therefore, the structure size ratios of the different PI_xPS_yAmPEO_{1.9} to each other are in accurate accordance with their respective $M_{n,NMR}$ (cf. Scheme 4, Table 4). Also, the measured r_{cylinder} have the same relationship to each other as the previously described d (cf. Table 6).

Using the $M_{n,NMR}$ of the individual polymer blocks (cf. Table 3) and on basis of published homopolymer densities ($\rho_{PI_x} = 0.830$, $\rho_{PS_y} = 0.969$, $\rho_{PEO_z} = 1.064$ in g cm⁻³, note that these densities were determined for 140 °C, which may cause slight deviations) [61], the monomer volume fractions (f_i) of the PI_x block (f_{PI_x}), the PS_y block (f_{PS_y}) and the PEO_{1.9} block ($f_{PEO_{1.9}}$) were calculated and listed in Table 6. For PI_{6.8}PS_{17.3}AmPEO_{1.9}, PI_{14.6}PS_{34.8}AmPEO_{1.9} and PI_{26.1}PS_{67.3}AmPEO_{1.9}, the volume ratio of the PI_x block to the PS_yAmPEO_{1.9} block always remains the same and was in the range of ~29–33% for f_{PI_x} and ~68–71% for $f_{PS_y\text{AmPEO}_{1.9}}$. The volume fraction of the cylinder ($\varphi_{\text{cylinder}}$) was equal for all with ~28–34%, which matches very well with the f_{PI_x} . Accordingly, it was obvious that they have the same HEX structure. Consequently, the cylinder from the HEX structure corresponds to the PI_x block, because the r_{cylinder} ratio of the different polymers, i.e., PI_{6.8}PS_{17.3}AmPEO_{1.9}, PI_{14.6}PS_{34.8}AmPEO_{1.9} and PI_{26.1}PS_{67.3}AmPEO_{1.9}, agrees exactly with the respective M_{n,NMR,PI_x} ratio. As a result, the HEX structures differ from each other only in terms of size.

The volume ratio for PI_{35.1}PS_{14.8}AmPEO_{1.9} with $f_{PI_{35.1}} = 71\%$ and $f_{PS_{14.8}\text{AmPEO}_{1.9}} = 29\%$ is exactly the opposite compared to the previously discussed PI_xPS_yAmPEO_{1.9}. The determined $\varphi_{\text{cylinder}}$ with 31% matches quite closely with the $f_{PS_{14.8}\text{AmPEO}_{1.9}}$. Consequently, the detected HEX structure fits accurately; however, due to the inverted volume fractions, a reverse phase with PS_{14.8}AmPEO_{1.9} cylinders exists. Hence, r_{cylinder} with a value of 10.9 nm corresponds to the $M_{n,NMR}$ of PS_{14.8} and AmPEO_{1.9} block and is the same size as r_{cylinder} of

PI_{14.6}PS_{34.8}AmPEO_{1.9} as well as its $M_{n,NMR}$ of the PI_{14.6} block. This fact confirms that PI_{35.1}PS_{14.8}AmPEO_{1.9} has the reverse phase of the HEX structure of the PI_{14.6}PS_{34.8}AmPEO_{1.9}.

The measured LAM morphology for PI_{24.8}PS_{25.0}AmPEO_{1.9} with the volume ratio of $f_{PI_{24.8}} = 52\%$ and $f_{PS_{25.0}AmPEO_{1.9}} = 48\%$ fits exactly with the determined almost equal volume fraction of lamella (ϕ_{LAM}) of the PI_{24.8} block with $\phi_{LAM} = 52\%$ and 48% for the PS_{25.0}AmPEO_{1.9} block.

Table 6. Summary of SAXS results determined at room temperature of the as-cast membranes of the different PI_xPS_yAmPEO_{1.9}.

Polymer	$f_{PI_x}^1$ /%	$f_{PS_y}^1$ /%	$f_{PEO_{1.9}}^1$ /%	q^*^2 /nm ⁻¹	Phase ³	d^4 /nm	$r_{cylinder}^5$ /nm	$\phi_{cylinder}^6$ /%	Domain Size ⁹ /nm
PI _{6.8} PS _{17.3} AmPEO _{1.9}	29	65	6	0.32	HEX	22.4	6.1	29 (PI _{6.8})	90
PI _{14.6} PS _{34.8} AmPEO _{1.9}	32	65	3	0.20	HEX	35.7	10.4	34 (PI _{14.6})	122
PI _{24.8} PS _{25.0} AmPEO _{1.9}	52	45	3	0.20	LAM	31.7	$t^7 = 16.4$	52 (PI _{24.8}) _{LAM} ⁸	86
PI _{35.1} PS _{14.8} AmPEO _{1.9}	71	26	3	0.19	HEX	39.2	10.9	31 (PS _{14.8} AmPEO _{1.9})	63
PI _{26.1} PS _{67.3} AmPEO _{1.9}	33	66	2	0.11	HEX	63.6	16.9	28 (PI _{26.1})	147

¹ f_i = monomer volume fraction on basis of published homopolymer densities ($\rho_{PI_x} = 0.830$, $\rho_{PS_y} = 0.969$, $\rho_{PEO_z} = 1.064$ in g cm⁻³) [61] and calculated using $M_{n,NMR}$ (cf. Table 3). Note that these densities were determined for 140 °C, which may cause deviations. ² q^* is the position of the first scattering peak, determined by SAXS. ³ HEX = hexagonally close-packed cylindrical and LAM = lamella structure, determined by SAXS measurement. ⁴ d = average domain spacing. ⁵ $r_{cylinder}$ = radius of cylinder determined by SAXS. ⁶ $\phi_{cylinder}$ = volume fraction of cylinder based on ratio of $r_{cylinder}$ to d . ⁷ t = layer thickness determined by SAXS. ⁸ Volume fraction of lamella based on ratio of t to d . ⁹ Domain size = determined by SAXS.

Indeed, as expected from the PI_xPS_y phase diagram, the nearly symmetric PI_{24.8}PS_{25.0}AmPEO_{1.9} forms a LAM, while all other, more asymmetric PI_xPS_yAmPEO_{1.9} show a HEX morphology [18,24].

In order to visualize the previously described structures more clearly, corresponding 3D drawings were plotted based on the parameters obtained from the SAXS measurements with the assumption that the PEO_{1.9} block is fully phase-separated (cf. 3D drawings on the right site in Figure 6).

The obtained SAXS results match almost perfectly for all PI_xPS_yAmPEO_{1.9} with their respective polymer composition determined from NMR, thermal analysis and GPC measurements, and in particular, when they are compared with each other.

For all PI_xPS_yAmPEO_{1.9}, the SAXS measurement results indicate a high degree of long-range order in the structure due to the strong scattering in the respective 2D SAXS pattern, but also due to the fact that the curve shape shows a number of clear peaks that agree very well with the fit as well as with the calculated peak positions. For instance, in PI_{24.8}PS_{25.0}AmPEO_{1.9} the presence of $2q^*$ and $3q^*$ suggests that the LAM morphology is present with a highly long-range arrangement [60,62].

In order to determine a more detailed, local long-range order as well as the orientation of the structure, SEM measurements were carried out. In particular, these two structural parameters of the membrane are crucial for the application of BCP as a structure-giving polymer matrix in electrolyte. In addition, the structure determined from the SAXS results is also checked. For this purpose, exemplary SEM measurements were performed on unstained ultra-thin sections of about 50 to 100 nm thickness from PI_{14.6}PS_{34.8}AmPEO_{1.9} for the HEX structure and from PI_{24.8}PS_{25.0}AmPEO_{1.9} for the LAM structure. In Scheme S2, the procedure of sample preparation by cryo-ultramicrotomy as well as sample placement onto the grid is schematically shown.

Results of PI_{14.6}PS_{34.8}AmPEO_{1.9} are shown in Figure 7; the ones for PI_{24.8}PS_{25.0}AmPEO_{1.9} are displayed in Figure 8. In (a), an overview image of several μm^2 in size is shown. The marked area in a) was enlarged as inset in b) to make the structure visible in more detail (i.e., periodic lamellae or hexagonal cylinder arrangement). In (c), the fast Fourier

transformation (FFT) of the area in (a) in case of the LAM and of the inset (b) in case of the HEX structure is added to indicate the average orientation of the structure. The good contrast of the sample is achieved through the relatively low accelerating voltage of 30 kV used for STEM-imaging in the SEM. Hence, samples can be measured without staining. Additionally, in both figures under (d), the corresponding SAXS pattern is shown to compare it with the FFT spectrum.

For $\text{PI}_{14.6}\text{PS}_{34.8}\text{AmPEO}_{1.9}$, a long-range and highly ordered HEX structure of cylinders is clearly visible in the overview image of Figure 7a, respectively, in the enlarged inset in (b) and corresponds with the SAXS result. In particular, it is remarkable that the hexagonal cylinder pattern of the HEX structure is clearly visible and almost identical in the FFT spectrum in (c) as well as in the SAXS pattern in (d). This means that the cylinders have the same arrangement and distance to each other in the macroscopic sample volume (2 mm sample diameter in the SAXS capillaries) and in the microscopic sample area in the SEM image and are aligned straight through the entire volume. Furthermore, from the SEM image, it is evident that for the cast film HEX structure, the cylinder axes are located in plane.

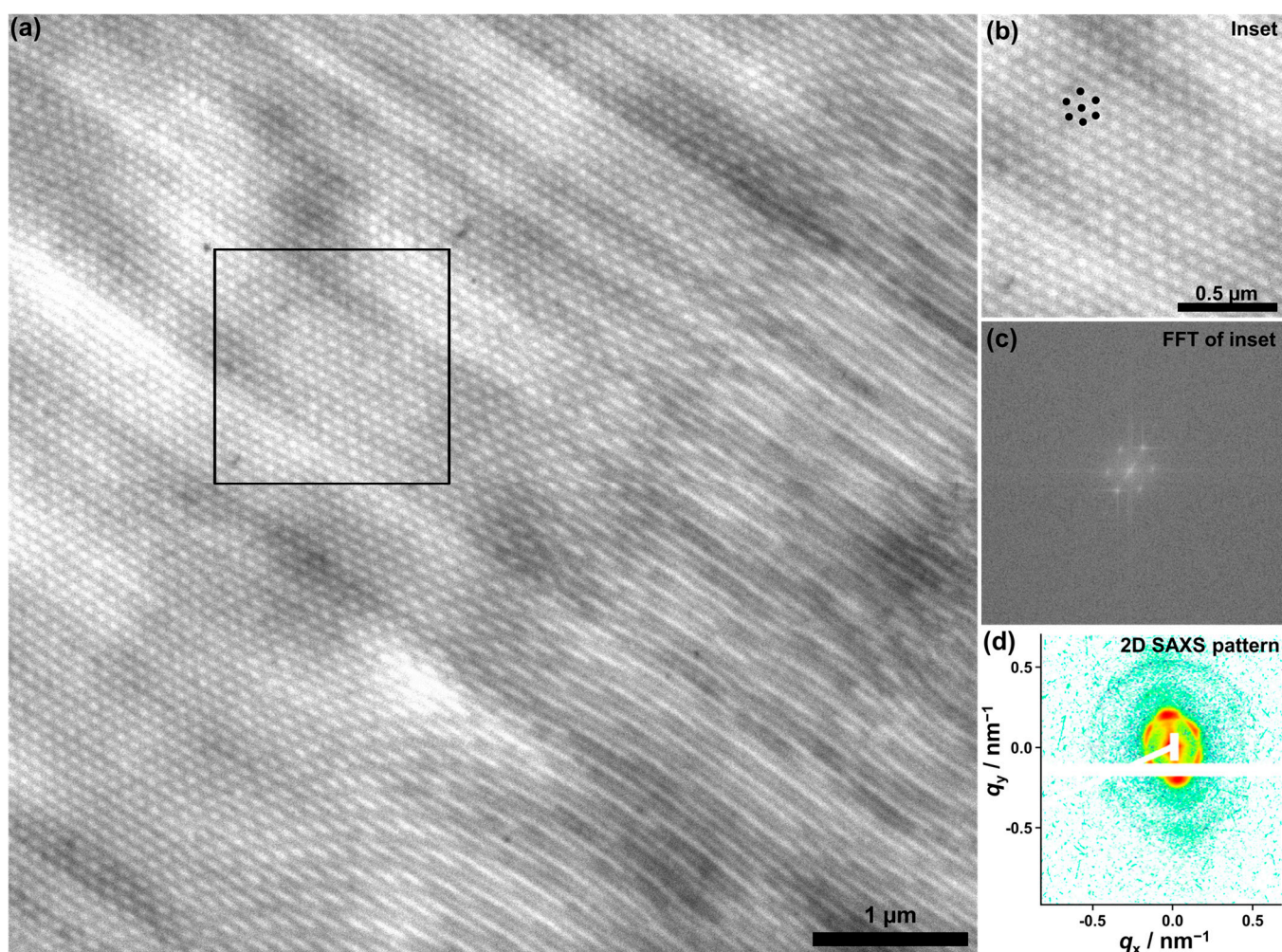


Figure 7. SEM images obtained at room temperature of self-assembled hexagonally close-packed cylindrical (HEX) domains in a 50 nm thick section of unstained as-cast membrane of $\text{PI}_{14.6}\text{PS}_{34.8}\text{AmPEO}_{1.9}$ without subsequent annealing. (a) Sample overview. (b) Zoomed area of inset marked in (a). For better visibility, the hexagonal cylinder pattern of the HEX structure is marked in black as an example. (c) Two-dimensional FFT spectrum of inset area. (d) Two-dimensional SAXS pattern obtained at room temperature from the membrane in as-cast condition with a sample diameter of 2 mm.

In case of $\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$ in the large overview image in Figure 8a, respectively, in the inset in b), a long-range and highly ordered LAM structure is clearly evident, which is also consistent with the SAXS result and corroborated by the 2D FFT result. The obtained 2D FFT spectrum reveals two “beam-like shapes” according to the LAM structure going vertical through the sample (cf. Figure 8c). Also, the SAXS pattern in Figure 8d shows the two “beam-like shapes” which differ from those in the FFT nearly only by the fact that they are slightly tilted. Most likely this is due to the placement of the sample into the SAXS capillary. Considering that the FFT spectrum and the SAXS pattern look almost the same, the lamellae are microscopically as well as macroscopically equally oriented and have the same distance to each other.

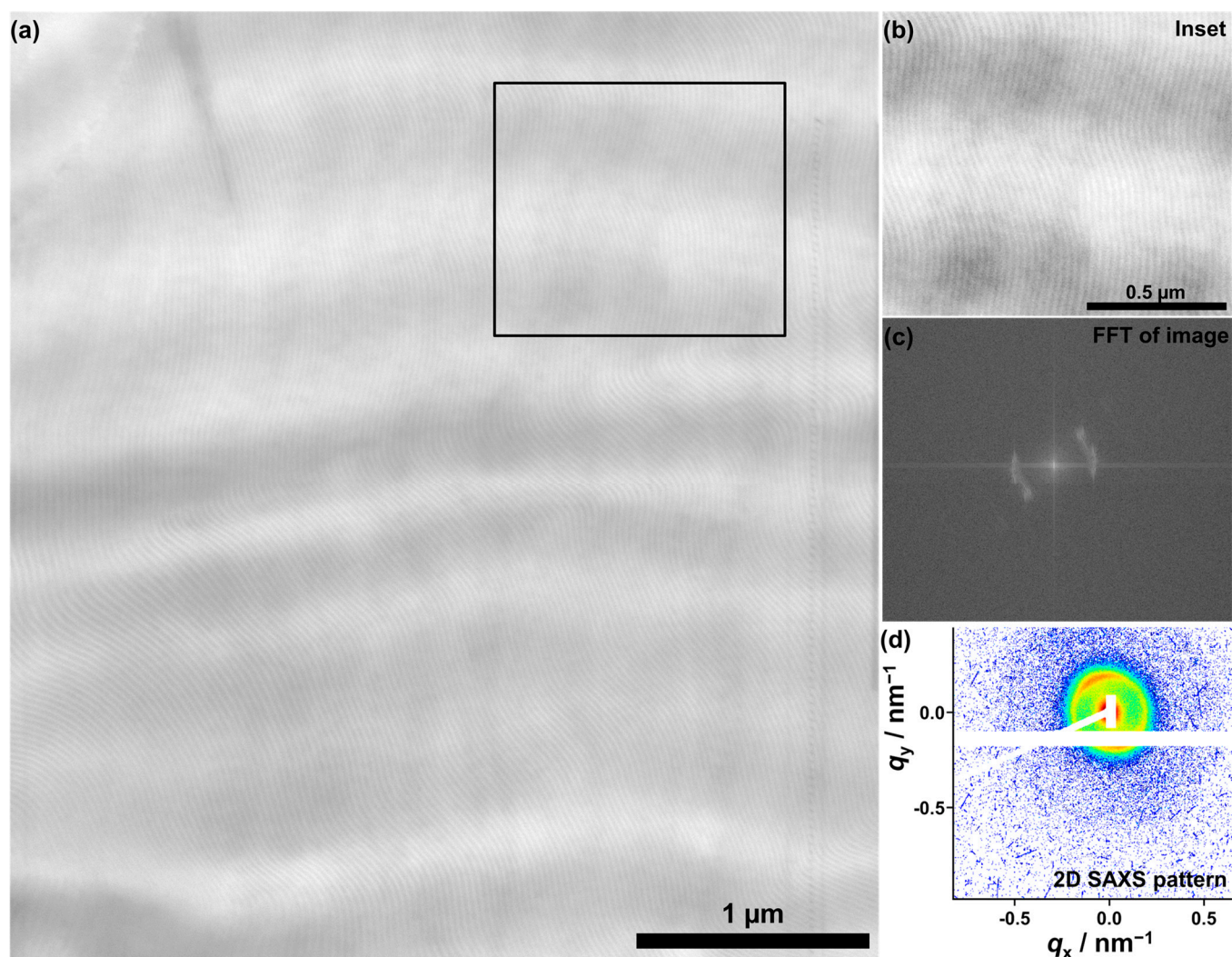


Figure 8. SEM images obtained at room temperature from an ultra-thin section of self-assembled lamellar domains of unstained as-cast membrane of $\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$ without subsequent annealing. (a) Sample overview. (b) Zoomed image of area marked in the overview image in (a). (c) Two-dimensional FFT spectrum over the entire overview image (a). (d) Two-dimensional SAXS pattern obtained at room temperature from the membrane in as-cast condition with a sample diameter of 2 mm.

Moreover, a TEM image of $\text{PI}_{24.8}\text{PS}_{25.0}\text{AmPEO}_{1.9}$ of the unstained membrane was obtained in STEM mode. Usually, stained samples are measured [18,24,63,64]. The TEM result shown in Figure 9 displays the same LAM structure as already seen in SEM.

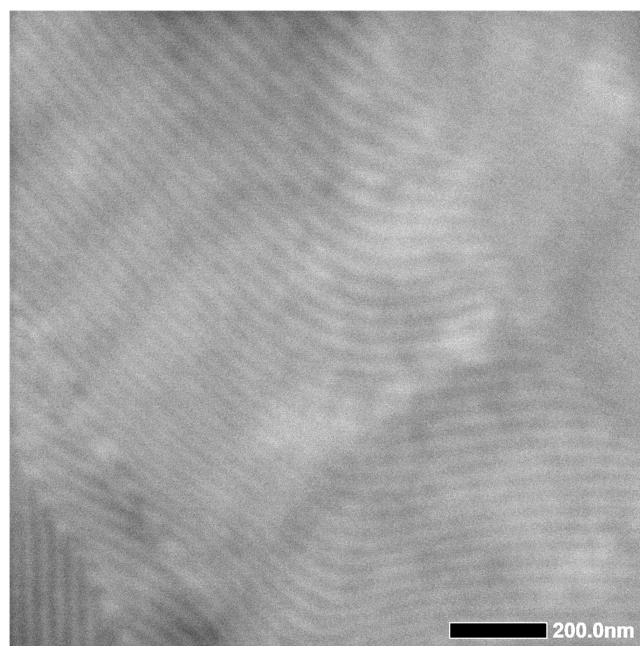


Figure 9. STEM image obtained at room temperature from a 50 nm thick section of self-assembled lamellar domains of an unstained as-cast membrane of PI_{24.8}PS_{25.0}AmPEO_{1.9} without subsequent annealing.

As a result, PI_{14.6}PS_{34.8}AmPEO_{1.9} and PI_{24.8}PS_{25.0}AmPEO_{1.9} possess a nearly perfectly ordered HEX respectively LAM structure, at least over the total SEM and TEM measurement μm size area as well as the 2 mm thick sample volume for the SAXS measurement. Therefore, they show an extraordinary long-range orientation. The strongly distinct microphase separation, most likely caused by the very narrow polymer chain length distribution, in combination with the controlled solution casting process ensuring enough time for self-assembly, leads to these exceptional morphological properties. These properties are crucial for an excellent polymer electrolyte matrix, because the conductive pathways and finally the Li⁺ transport should be aligned almost optimally to each other over an extremely long range inside the BCP membrane.

Furthermore, particularly in case of PI_{24.8}PS_{25.0}AmPEO_{1.9}, horizontal cutting artifacts can be recognized in the sample overview image as brighter and darker waves. This is a common cutting artifact of cryo-ultramicrotomy, which results from compression of the section during sectioning and extends perpendicular to the cutting direction [65]. As the cutting direction was perpendicular to the membrane and a ribbon of sections could be imaged on the grid by SEM, it is possible to determine the orientation of the lamellae within the membrane over a larger sample surface area and at several sample positions (cf. Scheme S2). The lamellae are arranged from top to bottom, i.e., vertically across the sample. Compression artefacts are perpendicular to the lamellae, indicating that the cutting direction across the membrane is the same as the orientation of the lamellae in the BCP. In other words, the lamellae are continuous aligned orthogonally to the polymer surface and therefore also at the potential use in electrolytes as a structure-giving BCP matrix to the electrode interface, i.e., connecting both electrodes to each other. It has already been shown in our patent that these BCPs can be used as a structure-giving BCP matrix in electrolytes to obtain very high ionic conductivities [43]. Such electrolytes provide comparable, and in some cases even better, ionic conductivities than those reported by Dörr and Pelz et al. [10,11,37].

4. Conclusions

Herein, we introduced a convergent synthesis method based on the modular principle, which ensures the access to well-defined $PI_xPS_yPEO_z$ linear triblock copolymers with consistently the same very short and precise PEO_z block. For this purpose, a prefabricated and commercially available $mPEO_z$ block is used, which is selectively functionalized with an epoxy end group to $EmPEO_z$. The $EmPEO_z$ block is covalently attached to the $PI_xPS_y^-$ anion synthesized by living anionic polymerization and thus terminated to the corresponding BCP. Thereby, by utilizing the O-Li ion pair formation between the epoxide group of the $EmPEO_z$ and the living anion, it is ensured that only a single ring-opening reaction occurs. Therefore, the handling of EO gas monomers during the BCP polymerization can be avoided.

The systematic variation of the block length reveals two major and independent influences of the polymer structure. By varying the $M_{n,PI_x}/M_{n,PS_y}$ at a constant $M_{n,total}$, the morphology could be precisely controlled. Alternatively, by changing the $M_{n,total}$ at the same $M_{n,PI_x}/M_{n,PS_y}$, only the PEO_z block fraction could be altered and adjusted, while keeping the morphology constant.

This simplified and reproducible one-pot polymerization, with 100% reaction efficiency, obtains highly ordered $PI_xPS_yAmPEO_z$ BCPs whose morphology can be largely controlled within the polymer membrane on microscopic and macroscopic levels. Largely controlled means at the microscopic or molecular level that the polymer chain lengths can be precisely selected, are nearly monodisperse and the respective BCP possesses a very high degree of phase separation. Moreover, the highly uniform polymer blocks combined with the controlled solution casting process lead to $PI_xPS_yAmPEO_z$ membranes with a very high and exceptionally long-range ordered structure up to the mm scale.

Overall, due to the always constant PEO_z block size, the high control over the order and orientation of the morphology as well as the PEO_z block fraction, these BCPs are suitable for subsequent targeted investigation of their influence as a structure-giving matrix in corresponding solid polymer electrolytes with respect to Li^+ transport up to the macroscale level.

Supplementary Materials: The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/polym15092128/s1>, Scheme S1: Illustration of the utilization of the blocking effect for single $EmPEO_z$ chain attachment to the $PI_xPS_y^-$ anion with Li^+ as counterion due to the strong O-Li affinity. Figure S1: 1H -NMR spectrum of $EmPEO_{1.9}$. Figure S2: ^{13}C -NMR spectrum of $EmPEO_{1.9}$. Figure S3: 1H -NMR spectrum of $mPEO_{1.9}$. Figure S4: ^{13}C -NMR spectrum of $mPEO_{1.9}$. Figure S5: 1H -NMR spectrum of $PI_{6.8}PS_{17.3}AmPEO_{1.9}$. Figure S6: 1H -NMR spectrum of $PI_{14.6}PS_{34.8}AmPEO_{1.9}$. Figure S7: 1H -NMR spectrum of $PI_{24.8}PS_{25.0}AmPEO_{1.9}$. Figure S8: 1H -NMR spectrum of $PI_{35.1}PS_{14.8}AmPEO_{1.9}$. Figure S9: 1H -NMR spectrum of $PI_{26.1}PS_{67.3}AmPEO_{1.9}$. Figure S10: Comparison of the characteristic signals in the 1H -NMR spectra of the synthesized $PI_xPS_yAmPEO_{1.9}$. Table S1: Comparison of the number of protons determined from 1H -NMR integrals of the individual polymer blocks of the synthesized $PI_xPS_yAmPEO_{1.9}$. Figure S11: Full GPC trace of $PI_{6.8}PS_{17.3}AmPEO_{1.9}$. Figure S12: Full GPC trace of $PI_{14.6}PS_{34.8}AmPEO_{1.9}$. Figure S13: Full GPC trace of $PI_{24.8}PS_{25.0}AmPEO_{1.9}$. Figure S14: Full GPC trace of $PI_{35.1}PS_{14.8}AmPEO_{1.9}$. Figure S15: Full GPC trace of $PI_{26.1}PS_{67.3}AmPEO_{1.9}$. Figure S16: DSC measurement of $EmPEO_{1.9}$. Figure S17: Exemplary photo of a colorless and transparent $PI_xPS_yAmPEO_{1.9}$ membrane cast from THF. Figure S18: Two-dimensional SAXS patterns obtained at room temperature from the membrane in as-cast condition with a sample diameter of 2 mm of the synthesized $PI_{6.8}PS_{17.3}AmPEO_{1.9}$, $PI_{35.1}PS_{14.8}AmPEO_{1.9}$ and $PI_{26.1}PS_{67.3}AmPEO_{1.9}$. Details concerning fits to radially averaged SAXS data. Scheme S2: Procedure of sample preparation by cryo-ultramicrotomy as well as sample placement onto the grid.

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