

Article

Optimizing NOM Removal: Impact of Calcium Chloride

Alfredo Gonzalez-Perez ^{1,2,*}, Kristofer Hägg ^{3,4,*} and Fabrice Duteil ⁵

¹ Department of Materials Science and Applied Mathematics, Faculty of Technology and Society, Malmö University, Nordenskiöldsgatan 1, 211 19 Malmö, Sweden

² Vatteninfo Sverige AB, Kaserngatan 11, 761 46 Norrtälje, Sweden

³ Department of Building and Environmental Technology, Lund University, John Ericssons Väg 1, Box 118, 221 00 Lund, Sweden

⁴ Sweden Water Research AB, Ideon Science Park, Scheelevägen 15, 223 70 Lund, Sweden

⁵ TETRA Chemicals Europe AB, P.O. Box 901, 251 09 Helsingborg, Sweden; fduteil@tetrachemicals.com

* Correspondence: alfredo.gonzalez-perez@mau.se or alfredo@vatteninfo.com (A.G.-P.); kristofer.hagg@sydvatten.se (K.H.)

Abstract: Understanding the character of natural organic matter (NOM) and assessing its impact on water quality is paramount for managers of catchments and water utilities. For drinking-water producers, NOM affects disinfectant demand and the formation of by-products which can have adverse health effects. NOM content in raw waters also has an impact on water treatment processes by increasing required coagulant dosages, reducing the effectiveness of adsorption processes and fouling membrane systems. This study investigated the effects of calcium chloride (CaCl_2) as a co-coagulant in Al^{3+} and Fe^{3+} assisted coagulation, flocculation and sedimentation processes for NOM-removal from raw water collected from Lake Bolmen, in southern Sweden. Jar tests were conducted at Ringsjö Water Works (WW), a surface water treatment plant (WTP), to investigate the potential reduction in primary coagulants aluminum sulphate ($\text{Al}_2(\text{SO}_4)_3$) and ferric chloride (FeCl_3). This work shows that CaCl_2 can, in certain situations, reduce the need for primary coagulants, which would reduce the environmental impact and costs associated with primary coagulant consumption.

Keywords: NOM; coagulation; flocculation; Lake Bolmen; aluminum sulphate; ferric chloride; calcium chloride; raw water



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1. Introduction

Natural organic matter (NOM) is generated through the breakdown of bacteria, algae and both terrestrial and aquatic plants, resulting in a complex mixture of organic substances. In freshwater bodies used as a raw water source for drinking water production NOM needs to be removed, representing a substantial challenge due to changes in seasonal NOM composition. During drinking water production, the percentage of removal of NOM affects the efficacy of drinking water treatment processes. If chlorine is used as a disinfectant, the presence of NOM will result in the generation of potentially carcinogenic disinfection byproducts [1,2]. A comprehensive review of the different methods for the characterization of NOM in relation to drinking water production was published by Matilainen et al. [3].

Advanced technologies for the removal of NOM include enhanced coagulation, membrane filtration, and granulated activated carbon, among others [4]. The different existing technologies for water treatment, including NOM removal, have been summarized elsewhere [5,6]. A wide variety of conventional and non-conventional adsorbents have been recently reviewed by Bhatnagar and Sillanpää [7] for their potential in NOM removal from water. The authors showed that modified adsorbents, composite materials and a few nanomaterials have shown promising results for NOM removal from water. A brief overview of the state-of-the-art methods for NOM removal from supply waters was recently published by Callegari et al. [8].

Among the different options, coagulation is one of the most widely used techniques for the removal of NOM in drinking water production [9] because of its low cost and efficiency in removing both disinfection by-product precursors and turbidity [10]. The most commonly used coagulants are based on trivalent ions such as aluminum sulphate, $\text{Al}_2(\text{SO}_4)_3 \cdot x\text{H}_2\text{O}$ (with x between 14 and 21), aluminum chloride and sodium aluminate. There are also iron-based coagulants such as ferric sulphate, ferrous sulphate, ferric chloride and ferric chloride sulphate [11].

Coagulation operates by destabilizing the colloidal particles that comprise NOM and that remain in solution due to their natural negative charge. Of the four known methods to destabilize colloids—charge neutralization, double layer compression [12], particle adsorption and bridging, and enmeshment in a precipitate [13]—three of them can be addressed by using aluminum sulphate or ferric chloride, which makes these coagulants ideal choices for a wide variety of water compositions in different areas or that experience seasonal variations [14]. It is simple to adjust the coagulant dosage and pH value for different raw water compositions, making the process easy to operate. When using aluminum, the pH range for coagulation was found to be between approximately 5.2 and 6.4, while in the case of using iron as a coagulant, the pH range is between 4.2 and 5.5 [13]. The speciation and solubility are controlled by pH which is a key parameter to adjust during the coagulation process.

Different factors affect the coagulant's effectiveness, such as raw water alkalinity, pH, turbidity, and mixing conditions. Other NOM properties such as molecular weight distribution, charge density, solubility and functional groups can also affect the coagulation process [15]. The different methods used for the removal of NOM from drinking water, including coagulation, adsorption, oxidation, membrane, biological treatment processes and the combination of different treatment processes, have been reviewed elsewhere [4]. The use of co-coagulants is a popular alternative to improve the efficiency of NOM removal and reduce cost.

At the time of this study, very little was published on CaCl_2 and its use in water treatment [16] in comparison with other compounds used for similar purposes [17,18]. We know that CaCl_2 is more readily available than metal salts and does not have the same environmental impact. Studying CaCl_2 also suggests a potential economic impact for the industry if calcium can replace primary coagulants in drinking water production. The following economic benefit would result from the reduced requirement of $\text{Al}_2(\text{SO}_4)_3$: (a) CaCl_2 costs less than this primary coagulant, (b) the amount of sludge is reduced, (c) the amount of chemicals used for pH-adjustment is potentially decreased.

The reduction in primary coagulant in the sludge also eases sludge management in a broader sense. Primary coagulant often includes undesirable contaminants, such as manganese, which can increase to a significant concentration when they accumulate in the sludge. Handling the sludge with less Al (or Fe) and other contaminants is a significant benefit. However, these benefits still assume that the produced water quality is sufficient for its purpose.

In this paper, we investigate the effects of using CaCl_2 as a co-coagulant aid in the removal efficiency of NOM from raw water obtained from Lake Bolmen, the tenth largest lake in Sweden. The main coagulants used in this study were aluminum sulphate and ferric chloride, and the effect of adding calcium chloride was investigated.

2. Materials and Methods

2.1. Lake Bolmen

In this study, raw water samples were taken from Lake Bolmen, in southern Sweden. Lake Bolmen is an oligotrophic lake with high NOM and color content caused by a catchment area dominated by forest and iron-rich soil [6,19]. Water from Lake Bolmen is currently used as raw water for Ringsjö Water Works (WW), a surface water treatment plant (WTP). During this study, the treatment process included coagulation with FeCl_3 , floc-

ulation, lamella sedimentation, rapid and slow sand filtration, and disinfection through chlorination (Sydvatten AB 2016, Malmö, Sweden).

2.2. Jar Tests

The experiments were carried out using a flocculator with 6 glass beakers of 1 L each with programmable agitators (Figure 1). Before each experiment, pH tests were conducted in 1 L raw water using NaOH or H₂SO₄ and primary coagulant. The test was designed to make sure the desired pH was achieved during the jar tests.



Figure 1. Photo of the Kemira Flocculator 2000 (Helsingborg, Sweden) used during the study.

Each jar test began with the addition of CaCl₂ to 1 L of raw water, followed by the pre-determined NaOH or H₂SO₄ amount. The program was started, and a 200 g/L aluminum sulphate (ALG) solution was added with 3 s remaining of the rapid stirring (400 rpm). The program continued with 20 min slow stirring (50 rpm), during which the resulting pH was measured for each beaker. Lastly, the stirring was stopped, allowing for sedimentation (0 rpm for 30 min). When the sedimentation was complete, samples were taken using a 50 mL plastic syringe approximately 3 cm below the surface.

2.3. Coagulants and Additives

In this study, aluminum sulphate (ALG), Al₂(SO₄)₃, and FeCl₃ (PIX-311), a 40% by weight FeCl₃ solution, were used as the primary coagulants. ALG was prepared by dissolving pellets in distilled water to a concentration of 200 g ALG/L solution. The CaCl₂ used in the experiments was a 36% by weight food-grade quality solution provided by TETRA Chemicals AB. A NaOH (AkzoNobel, Angered, Sweden) and H₂SO₄ solution was used to control the resulting pH during the experiments.

2.4. Sample Analysis

Water samples were tested for pH and analyzed in a spectrophotometer (Hach DR 5000, Düsseldorf, Germany) where the water was tested for residues of NOM. For the NOM reduction, the absorption was measured at $\lambda = 254$ nm and $\lambda = 436$ nm with a 5 cm cuvette.

The pH was measured using a WTW pH-197 (Sigma-Aldrich Sweden AB, Stockholm, Sweden) pH meter, which contains a built-in thermometer that corrects for temperature. The program was started and when 3 s remained of the rapid mixing, coagulant was added. Once the rapid mixing was finished and slow mixing had started, the samples were pH-tested.

3. Results and Discussion

The raw water samples used in this work were taken from Lake Bolmen, located in southern Sweden (56°54'59.99" N, 13°39'59.99" E). The lake covers 184 km², has a maximum depth of 37 m and is the tenth largest lake in Sweden. The lake supplies part

of Skåne region with fresh water through an 80 km long tunnel (Bolmen water tunnel), followed by a 25 km long pipeline that ends in the municipality-owned drinking water production plant (Ringsjö Station) located in the municipality of Stehag. Samples were taken at the end point.

In Table 1, we show the analysis of raw water from the raw water supplied by Lake Bolmen at Ringsjö Station. The table shows the chemical composition of the raw water, including salinity and conductivity as well as its high UV-VIS absorbance.

Table 1. Raw water analysis of water from Bolmen Lake conducted on 14 September 2015. Samples were taken at Ringsjö Station.

Parameter	Results	Units
Temperature	23.2	°C
pH (T = 23.2 °C)	6.9	
Turbidity	1.1	FNU
CODMn	8.9	mg O ₂ /L
TOC	7.1	mg/L
Conductivity	7.2	mS/m
Alkalinity, HCO ₃	14	mg/L
Calcium, Ca	6.1	mg/L
Magnesium, Mg	1.4	mg/L
Total hardness (odH)	1.2	°dH
Aluminum, Al	0.031	mg/L
Iron, Fe	0.39	mg/L
Color (410)	52	mg Pt/L
The average VIS Abs 436 nm, m	2.26	
The average UV abs 254 nm, m	34.6	

Iron and aluminum-based coagulants are commonly used for different tasks in the water and wastewater treatment industry. In particular, iron and aluminum coagulants have been used for odour and pathogen reductions and the enhancement of dewaterability in sludge digesters [20]. Aluminum sulphate (Kemira™ ALG) and ferric chloride (Kemira™ PIX-311) are used as primary coagulants for NOM removal applied directly to the raw water.

3.1. Conventional Coagulation

In the first stage, both ALG and PIX-311 were used as main coagulants upon raw water collected from Lake Bolmen at the optimal pH operational range recommended for those primary coagulants. The jar experiments described above were conducted between pH 6 and 7 and again between pH 4.6 and 5.6.

In Figure 2, we show absorption values after flocculation and sedimentation for a series of samples at different pH using ALG as a primary coagulant.

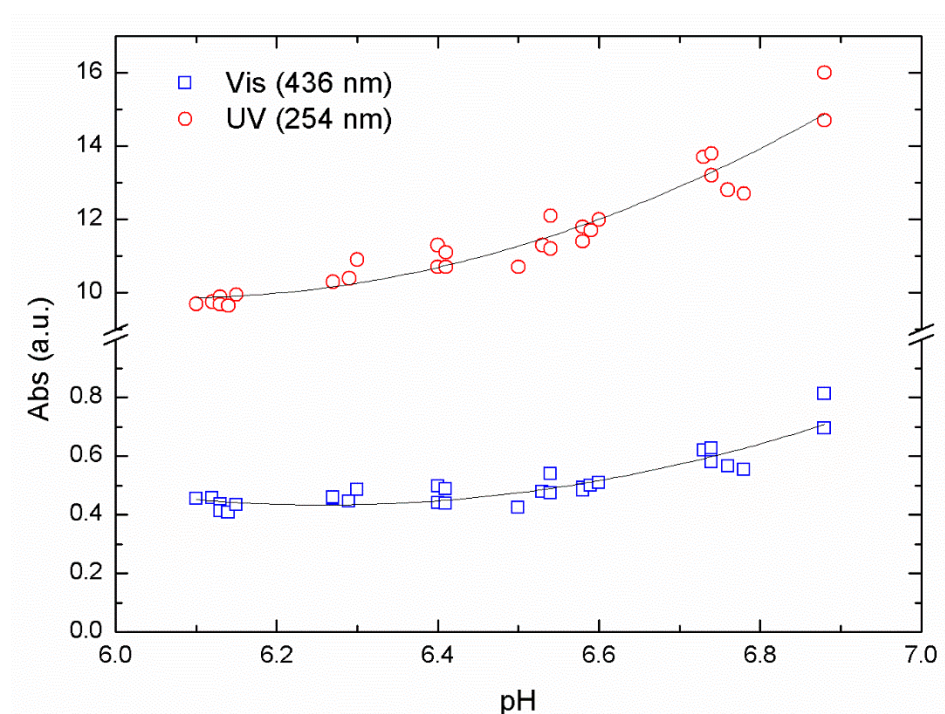


Figure 2. UV-Vis absorbance for raw water from Lake Bolmen using ALG as primary coagulant at pH values between 6 and 7.

In Figure 2, the reduction in the absorbance value passes through a minimum that represents the optimal reduction for the removal of NOM in this pH range. The optimal pH value was found between pH 6.3 and 6.1, resulting from the data collected using absorbance values of 254 nm (UV) and 436 nm (Vis) from the same samples. Using a second-order polynomial fit over our data, we obtained an R^2 of 0.83 and 0.92 for UV and Vis data respectively.

It has been reported that iron coagulants have the advantage of producing a denser floc than that produced by aluminum sulphate, resulting in improved settlement characteristics but at the expense of an approximately 40% increase in the weight of hydroxide sludge when compared to aluminum coagulants [5]. As many iron coagulants contain approximately 2–6 g of manganese per kg of iron as an impurity depending on the product specification, this might be an aspect of sludge management that needs to be taken into consideration.

Experiments with iron-based coagulant were performed with the same raw water to compare with the aluminum-based coagulant ALG used in this study. A series of jar samples were prepared using PIX-311 as a primary coagulant using raw water from Lake Bolmen, within the pH range of from 4.5 to 6. The results of absorbance in the UV and Vis ranges are shown in Figure 3.

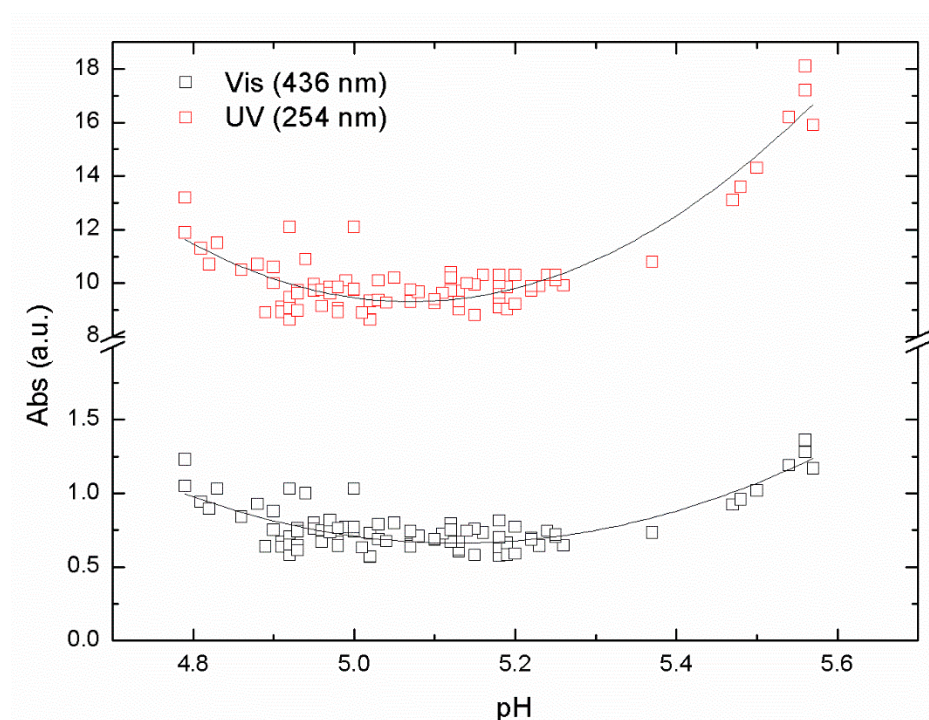


Figure 3. UV-Vis absorbance for raw water from Bolmen Lake using PIX-311 as primary coagulant at pH values between 4.5 and 6.

The results in Figure 3 again show a clear U-shape behavior with a minimum around pH 5.1 from the absorbance data at the UV and Vis wavelengths used in our study. Those pH minimum values represent the optimal pH for the removal of NOM from raw water taken from Lake Bolmen. The R^2 values from the polynomial fitting were 0.65 and 0.84 for UV and Vis data, respectively. The minimum value of the absorbance reached the point of maximum NOM removal efficiency. The pH optimal values were around 5.1 for both sets of data, UV and Vis. Interestingly, the optimal pH value for Vis was slightly higher than the optimal pH for UVA. This is likely due to differences in the organic matter, resulting in differences in requirements for charge neutralization, i.e., aromatic compounds that absorb UV light, requiring more positive charges.

3.2. Enhance Coagulation

Because conventional coagulation with aluminum sulphate or iron chloride can sometimes be unable to sufficiently reduce the presence of NOM in raw water for drinking water production, the use of additional co-coagulants or coagulant aid compounds can be an alternative enhanced coagulation treatment. The reduction in the amount of conventional coagulant needed in the NOM removal process can be of additional economic interest for water utilities if the co-coagulant is cheaper than the primary coagulant conventionally used. In the current study, we investigated the effect of CaCl_2 as a co-coagulant on NOM removal efficiency when using ALG and PIX-311 as primary coagulants.

As a first step, we sought to understand the effect of adding CaCl_2 directly as a supplement to the primary coagulant. As the FeCl_3 dosage of coagulant changes daily at the WTP (around $56.5 \text{ g PIX-311/m}^3$ or $140 \text{ mmol Fe}^{3+}/\text{m}^3$), the actual dosages of primary coagulant used during the experiments were defined as 100% dosage. For ALG dosages, 100% is equivalent to about $145 \text{ mmol Al}^{3+}/\text{m}^3$. The percentage (by mol %) of CaCl_2 added was based on the primary coagulant dosage.

In Figure 4, we show the effect of adding 40% (bottom) and 50% (top) of CaCl_2 to the primary coagulants at their optimal concentration efficiency.

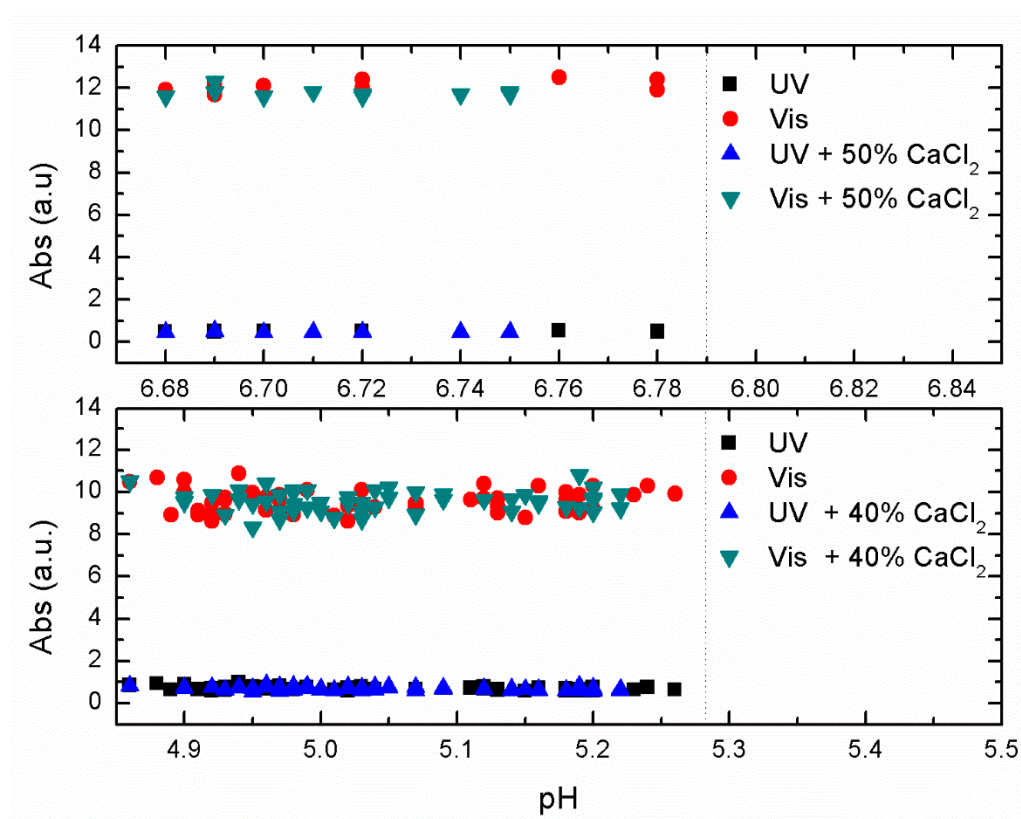


Figure 4. UV-Vis absorbance of raw water using optimal ALG concentration and adding 50% of CaCl₂ at pH values between 6.65 and 6.8 (**top**). The same treatment using optimal concentration of PIX-311 and adding 40% of CaCl₂ at pH values between 4.7 and 5.3 (**bottom**).

In Figure 4, where a limited set of jar experiments were done over a narrow range of pH around the optimal pH removal, we saw no significant improvement in removal efficiency with the addition of CaCl₂. This indicates no improvement when adding CaCl₂ to the primary coagulant under optimal conditions.

The experiments at 100% and 150% of CaCl₂ with optimal ALG concentration as well as 80% of CaCl₂ and optimal PIX-311 concentration, not shown in Figure 4, resulted in a similar lack of appreciable improvement in the removal efficiency of NOM. We can partially conclude that the direct addition of CaCl₂ as a co-coagulant to the optimal concentration of primary coagulant ALG or PIX-311 does not add any substantial or appreciable improvement in removal efficiency.

As a next step, we investigated the effect of adding CaCl₂ in combination with a lower concentration of primary coagulant to see the effect of replacing the primary coagulant with co-coagulant. In Figure 5, we show the results of optimal ALG and PIX-311 concentration compared to decreased concentration with and without the addition of CaCl₂ as a co-coagulant.

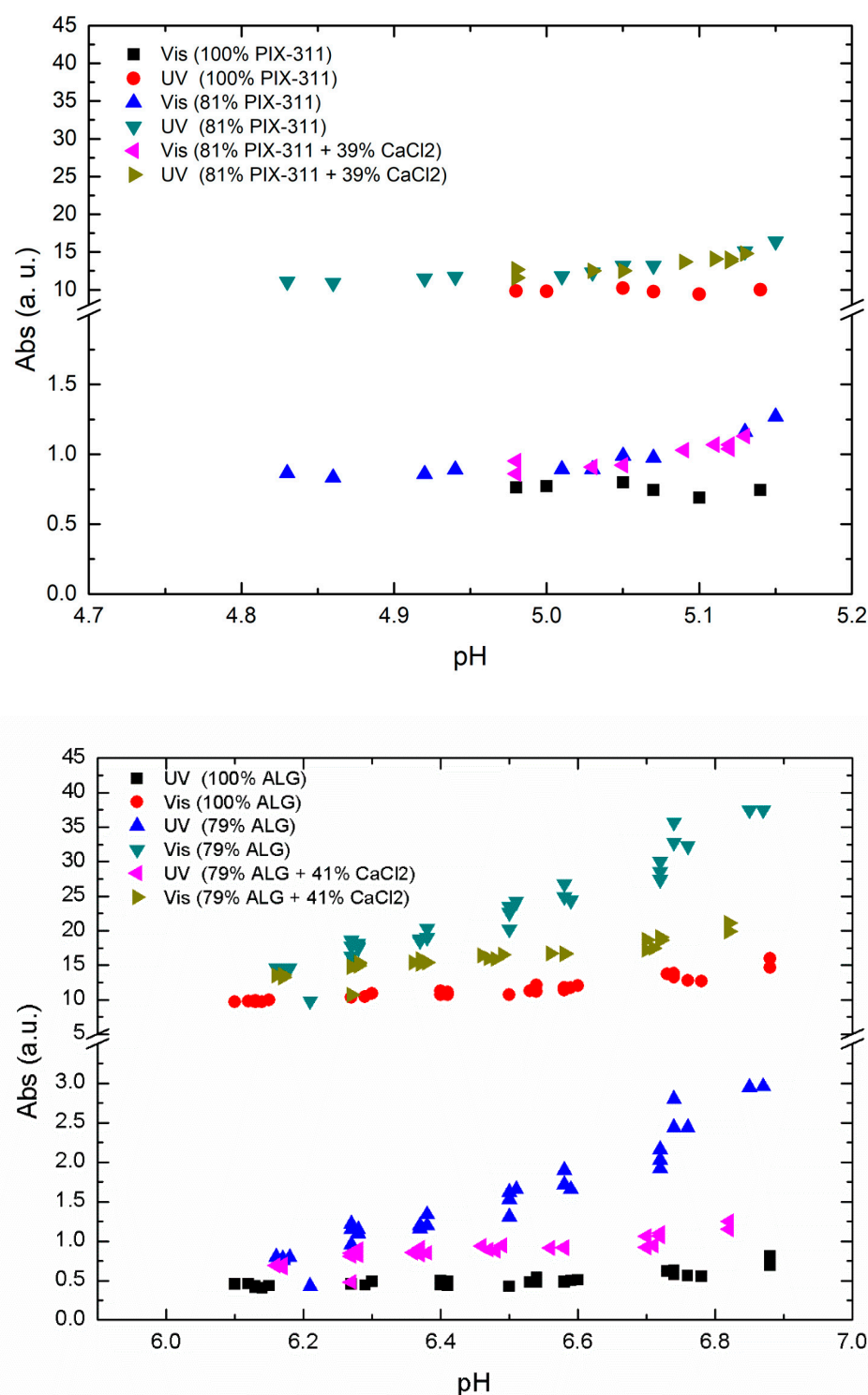


Figure 5. UV-Vis absorbance of raw water using PIX-311 (**top**) and ALG (**bottom**) vs. pH at optimal coagulant concentration, and at 81 and 79% concentration for ALG and PIX-311 respectively with and without an added 39 and 41% concentration respectively of CaCl_2 .

As shown in Figure 5 we observed a nearly zero effect from the addition of CaCl_2 in the case of PIX-311, while in the case of ALG there was an improvement in NOM removal when 79% of the optimal ALG concentration was combined with 41% of CaCl_2 . However, the efficiency presented by the optimum concentration of ALG was not matched by the combination of 71% ALG + 41% CaCl_2 . As the positive effects of calcium addition with

lower ALG dosages only occurred at a higher pH, we believe that the positive effect of adding CaCl_2 is due to calcium's ability to assist in charge neutralization. As seen in Figure 5, the positive effect of calcium chloride decreased with decreasing pH. This could mean that the positive effect of adding calcium chloride only occurs at higher pH values, and may explain why no positive effect was observed during Fe-assisted flocculation, which takes place at a lower pH. Another aspect is the natural calcium content in the raw water. In Lake Bolmen, the calcium content is very low, which could mean that the positive effects of calcium addition are reduced in other raw waters with higher calcium content. Despite the difference, the values reported by this combination could justify the use of CaCl_2 as a co-coagulant for economic purposes by reducing the overall chemical expenditures of drinking water utilities in some situations. This could be relevant in cases where the water quality demand is lower, e.g., as a pretreatment step before primary treatment. Irrespective of the primary coagulant used for natural organic matter, acidic pH values are favorable for their removal through the charge neutralization mechanism, as reported by Naceradska et al. [21].

The use of CaCl_2 in the tested water allowed for a significant reduction in the amount of added ALG coagulant. The research results presented here are representative of acidic waters. We observed that the efficiency of using CaCl_2 as a co-coagulant increased with increasing pH. These results are consistent with the results of research conducted by Kowalski and Turkiewicz [16], which showed that, in surface waters with $\text{pH} = 7.3\text{--}7.9$, the addition of calcium chloride in amounts of 10 g/m^3 and 20 g/m^3 to aluminum sulphate (ALG) allowed for an almost twofold increase in the efficiency of organic matter removal.

4. Conclusions

The use of aluminum sulphate and ferric chloride as a primary coagulant for the removal of NOM from raw water extracted from Lake Bolmen was investigated at approximately the optimal pH range. The results showed a U-shape behavior with an optimal minimum around pH 6.1 for ALG and pH 5.1 for PIX-311. A sustainable use of coagulation chemicals and co-coagulants requires the minimization (reduction) of the coagulant concentration.

The effect of adding CaCl_2 as a co-coagulant was investigated to understand the possible positive effect in the removal of NOM. The direct addition of CaCl_2 at different concentrations, with the optimal concentration of primary coagulant, resulted in no positive effect, while the addition of CaCl_2 at approximately 40% to a reduced primary coagulant concentration of approximately 70% of the optimal value resulted in no positive effect for PIX-311 and a positive reduction in NOM in the case of ALG.

We conclude that CaCl_2 could be used under certain conditions at higher pH (above 6) as a co-coagulant aid for the reduction in NOM when aluminum sulphate is used as a primary coagulant. This could result in some economic advantage by reducing the concentration of ALG needed for the coagulation process, and potentially reduce residual aluminum caused by high flocculation. Overall, the use of CaCl_2 as a co-coagulant needs to be tested in each specific case because the characteristics of the raw water will affect the NOM removal efficiency. Hence, its use in other water utilities with different characteristic raw water compositions needs to be tested to properly assess the suitability of primary coagulant concentration reduction. Our results are in agreement with those reported by Kowalski and Turkiewicz showing that the addition of calcium chloride allows for an increase in the efficiency of organic matter removal. Another potential concern is that reducing the amount of primary coagulant could lessen the need for polishing steps after primary treatment, such as increasing the pH after flocculation. In addition, allowing flocculation to occur at a higher pH has a less corrosive effect on the flocculation basins.

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References

1. Roccaro, P.; Vagliasindi, F.G.A.; Korshin, G.V. Changes in NOM Fluorescence Caused by Chlorination and their Associations with Disinfection by-Products Formation. *Environ. Sci. Technol.* **2009**, *43*, 724–729. [CrossRef] [PubMed]
2. Singer, P.C. Humic substances as precursors for potentially harmful disinfection by-products. *Water Sci. Technol.* **1999**, *40*, 25–30. [CrossRef]
3. Matilainen, A.; Gjessing, E.T.; Lahtinen, T.; Hed, L.; Bhatnagar, A.; Sillanpää, M. An overview of the methods used in the characterisation of natural organic matter (NOM) in relation to drinking water treatment. *Chemosphere* **2011**, *83*, 1431–1442. [CrossRef] [PubMed]
4. Zhang, Y.; Zhao, X.; Zhang, X.; Peng, S. A review of different drinking water treatments for natural organic matter removal. *Water Sci. Technol. Water Supply* **2015**, *15*, 442–455. [CrossRef]
5. Brandt, M.J.; Johnson, K.M.; Elphinston, A.J.; Ratnayaka, D.D. *Twort's Water Supply: Seventh Edition*; Elsevier Inc.: Amsterdam, The Netherlands, 2016; ISBN 9780081000250.
6. Ødegaard, H.; Østerhus, S.; Melin, E.; Eikebrokk, B. NOM removal technologies—Norwegian experiences. *Drink. Water Eng. Sci.* **2010**, *3*, 1–9. [CrossRef]
7. Bhatnagar, A.; Sillanpää, M. Removal of natural organic matter (NOM) and its constituents from water by adsorption—A review. *Chemosphere* **2017**, *166*, 497–510. [CrossRef] [PubMed]
8. Callegari, A.; Boguniewicz-Zablocka, J.; Capodaglio, A. Experimental Application of an Advanced Separation Process for NOM Removal from Surface Drinking Water Supply. *Separations* **2017**, *4*, 32. [CrossRef]
9. Sillanpää, M.; Ncibi, M.C.; Matilainen, A.; Vepsäläinen, M. Removal of natural organic matter in drinking water treatment by coagulation: A comprehensive review. *Chemosphere* **2018**, *190*, 54–71. [CrossRef] [PubMed]
10. Binnie, C.; Kimber, M.; Thomas, H. *Basic Water Treatment*; Royal Society of Chemistry: Cambridge, UK, 2017; ISBN 9780727763341.
11. Bratby, J. *Coagulation and Flocculation in Water and Wastewater Treatment*, 3rd ed.; IWA Publishing: London, UK, 2016; ISBN 9781780407500.
12. Crittenden, J.C.; Trussell, R.R.; Hand, D.W.; Howe, K.J.; Tchobanoglous, G. *MWH's Water Treatment Principles and Design*, 3rd ed.; John Wiley & Sons Inc.: Hoboken, NJ, USA, 2012; ISBN 9780470405390.
13. Gillberg, L.; Lindquist, A.; Hansen, B.; Karlsson, I.; Nordström Enkel, A.; Pålsson, A. *About Water Treatment*; Kemira Kemwater: Helsingborg, Sweden, 2003; ISBN 91-631-4353-4.
14. Sillanpää, M. *Natural Organic Matter in Water*, 1st ed.; Butterworth-Heinemann: Oxford, UK, 2014; ISBN 9780128015032.
15. Pernitsky, D.J.; Edzwald, J.K. Selection of alum and polyaluminum coagulants: Principles and applications. *J. Water Supply Res. Technol. AQUA* **2006**, *55*, 121–141. [CrossRef]
16. Kowalski, T.; Turkiewicz, D. Wpływ Dawkowania Soli Wapnia i Magnezu na Przebieg Koagulacji Domieszek Wód Siarczanem Glinu. Available online: http://www.os.not.pl/docs/czasopismo/1999/Kowalski_4-1999.pdf (accessed on 23 May 2021).
17. Bolto, B.A. Soluble polymers in water purification. *Prog. Polym. Sci.* **1995**, *20*, 987–1041. [CrossRef]
18. Loganathan, P.; Gradzielski, M.; Bustamante, H.; Vigneswaran, S. Progress, challenges, and opportunities in enhancing NOM flocculation using chemically modified chitosan: A review towards future development. *Environ. Sci. Water Res. Technol.* **2020**, *6*, 45–61. [CrossRef]
19. Eikebrokk, B.; Haaland, S.; Jarvis, P.; Riise, G.; Vogt, R.D.; Zahlsen, K. *NOMiNOR: Natural Organic Matter in Drinking Waters within the Nordic Region Norwegian Water*; Norwegian Water BA: Vangsvegen, Norway, 2018.
20. Akgul, D.; Abbott, T.; Eskicioglu, C. Assessing iron and aluminum-based coagulants for odour and pathogen reductions in sludge digesters and enhanced digestate dewaterability. *Sci. Total Environ.* **2017**, *598*, 881–888. [CrossRef] [PubMed]
21. Naceradska, J.; Pivokonska, L.; Pivokonsky, M. On the importance of pH value in coagulation. *J. Water Supply Res. Technol. AQUA* **2019**, *68*, 222–230. [CrossRef]